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Enantioselective Hydrolysis of Amino Acid Esters Promoted by Bis(β -cyclodextrin) Copper Complexes

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It is challenging to create artificial catalysts that approach enzymes with regard to catalytic efficiency and selectivity. The enantioselective catalysis ranks the privileged characteristic of enzymatic transformations. Here, we report two pyridine-linked bis(β -cyclodextrin) (bisCD) copper(II) complexes that enantioselectively hydrolyse chiral esters. Hydrolytic kinetic resolution of three pairs of amino acid ester enantiomers (S_1 – S_3) at neutral pH indicated that the “back-to-back” bisCD complex CuL^1 favoured higher catalytic efficiency and more pronounced enantioselectivity than the “face-to-face” complex CuL^2 . The best enantioselectivity was observed for *N*-Boc-phenylalanine 4-nitrophenyl ester (S_2) enantiomers promoted by CuL^1 , which exhibited an enantiomer selectivity of 15.7. We observed preferential hydrolysis of *L*- S_2 by CuL^1 , even in racemic S_2 , through chiral high-performance liquid chromatography (HPLC). We demonstrated that the enantioselective hydrolysis was related to the cooperative roles of the intramolecular flanking chiral CD cavities with the coordinated copper ion, according to the results of electrospray ionization mass spectrometry (ESI-MS), inhibition experiments, rotating-frame nuclear Overhauser effect spectroscopy (ROESY), and theoretical calculations. Although the catalytic parameters lag behind the level of enzymatic transformation, this study confirms the cooperative effect of the first and second coordination spheres of artificial catalysts in enantioselectivity and provides hints that may guide future explorations of enzyme mimics.

Because of the cooperation of the catalytic centre and the second coordination sphere, a native metalloenzyme can achieve biological transformations with remarkable efficiency and enantioselectivity under mild conditions *in vivo*^{1,2}. Over the years, significant efforts have been dedicated to the biomimetic study of metalloenzymes to understand their structures and functions and design analogues to mimic the native enzymes' structures and functions³. In the past few decades, synthetic compounds with functions resembling those of, for example, hydrolases, oxidases, and reductases have been reported^{4–8}. Remarkable progress has been made in mimicking the functions of the enzymes that mediate electron transfer with respect to the catalytic efficiency^{9,10}. However, studies of hydrolase mimetics have lagged far behind, regarding to either catalytic efficiency or selectivity^{5,7,8,11}. Specifically, in biomimetic chemistry, the rational design of enantioselective hydrolase mimics remains challenging.

Modified cyclodextrins (CDs) have been exploited to construct artificial enzymes since the 1970s, as introduced by Breslow^{5,6,12–14}. Previous studies have indicated that CDs' host-guest interactions could result in increased substrate binding affinity and cooperative binding between intramolecular CDs^{15–17}. In fact, the chiral nature of the CD cavity, particularly the cooperative effects of CDs with coordinated metal ion, has been somewhat underestimated. Appropriately modified CDs can be used in enantioselective molecular binding studies^{18,19}, and mediate chemical^{20–22} and photochemical^{23–25} enantio-differentiating reactions. CDs play an important role in preorganizing the guest molecules through the hydrophobic interactions. A number of studies have reported the use of unmodified or simply modified CD monomers without metal ions for the enantioselective deacylation of chiral esters^{26–32}. However, the cooperative enantioselective catalytic effect arising from CDs' intramolecular

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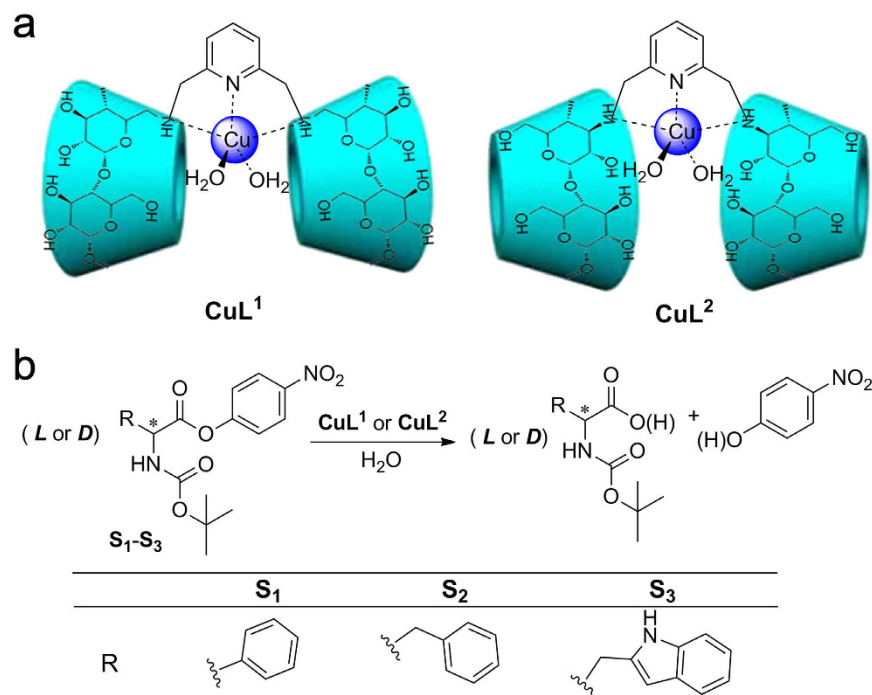


Figure 1. Schematic representations of (a) the structures of **CuL¹** and **CuL²** and (b) the hydrolysis of **S₁**, **S₂** and **S₃** catalysed by **CuL¹** and **CuL²**.

interactions with coordinated metal ions have rarely been reported, and the cooperative mechanism also remains unclear. We reported the construction of metallohydrolase mimics and superoxide dismutase mimics with CD domains^{5,33–39}. The hydrolase activities were studied with carboxylic acid esters and phosphate esters as model, non-chiral substrates. In this study, chiral *t*-butyloxycarbonyl (**Boc**)-protected aromatic amino acid esters **S₁–S₃** were chosen as substrates (Fig. 1). These substrates, which contain amino acid moieties, are biologically relevant than the previously used substrates and thus allowed us to evaluate the interactions of chiral CD cavities with naturally occurring substances. Two bisCD copper(II) complexes were utilized as enantioselective hydrolase mimics, denoted **CuL¹** (**L¹** = 2,6-bis(6-mono-amino- β -cyclodextrin-methyl)-pyridine, the “back-to-back” complex³⁹) and **CuL²** (**L²** = 2,6-bis(3-mono-amino- β -cyclodextrin-methyl)-pyridine, the “face-to-face” complex⁴⁰). The different CD orientations of the two mimics were studied in parallel to evaluate the effect of the chiral cavities on enantioselective catalysis. Electrospray ionization mass spectrometry (ESI-MS), an inhibition assay, rotating-frame nuclear Overhauser effect spectroscopy (ROESY), and theoretical calculations were performed to gain deeper insights into the underlying mechanism.

Results

Synthesis and characterization of **CuL¹ and **CuL²**.** The two copper(II) complexes, **CuL¹** and **CuL²**, were synthesized according to our previously reported procedures (see Supplementary Figs S1 and S2 for the ESI-MS results)^{35,38,39}. Because of the relationship between coordination geometry and catalytic ability, we were particularly interested in probing the coordination geometries of the copper(II) centre by using electron paramagnetic resonance (EPR) spectroscopy, which were performed at 100 K in a frozen solution of water and dimethylsulfoxide (DMSO) (see Supplementary Fig. S3). As a result, g_{\parallel} 2.25 and A_{\parallel} 160 were obtained for **CuL¹**, and g_{\parallel} 2.26 and A_{\parallel} 165 were obtained for **CuL²**, indicating a distorted square pyramidal coordination geometry at the copper(II) centre in both cases^{41,42}. This proposed geometry was also supported by the UV-Vis spectra (Supplementary Fig. S4), which showed a broad absorption band at approximately 700 nm that was assigned to the *d-d* transitions^{42,43}.

Hydrolysis of Boc-protected amino acid esters. The hydrolysis of **S₁**, **S₂** and **S₃** at neutral pH (50-mM 2-[4-(2-hydroxyethyl)piperazin-1-yl]ethanesulfonic acid (HEPES) buffer containing 10% CH₃CN, pH 7.2) and 298 ± 0.1 K were monitored by UV-Vis spectroscopy, focusing on the absorption at 400 nm, which were proportional to the concentration of the hydrolytic product 4-nitrophenolate (**NP**). Figure 2 reveals that the concentration of **NP** ($\epsilon_{\text{obs}} = 10398 \text{ M}^{-1} \text{ cm}^{-1}$) varied as a function of time during the **CuL¹**- and **CuL²**-promoted hydrolysis of **S₁**, **S₂** and **S₃**. The initial rate constants, k_{in} (s^{-1}) of the substrate cleavage events are listed in Table 1. Accordingly, **CuL¹** exhibited a higher catalytic efficiency than **CuL²**. Moreover, the differences in the hydrolysis rates of the substrate enantiomers in the reactions with **CuL¹** were more pronounced than those in the reactions with **CuL²**. It should be noted that in the presence of **CuL¹** (50 μM), *L*-**S₂** achieved the highest hydrolysis rate ($k_{\text{in}}^L = 5.4 \times 10^{-5} \text{ s}^{-1}$) among the tested substrates (Table 1), which was 168-fold higher than the spontaneous hydrolysis rate ($k_{\text{uncat}} = 3.2 \times 10^{-7} \text{ s}^{-1}$) (2.5 μM) (Supplementary Fig. S12). In contrast, the **CuL¹**-catalysed hydrolysis of *D*-**S₂** was 10-fold slower than that of *L*-**S₂** under the same conditions (Fig. 2b). The enantioselectivity was

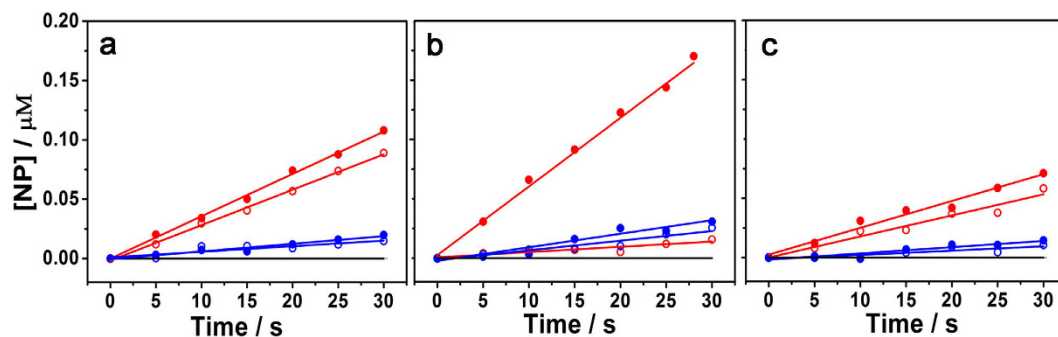


Figure 2. Initial-rate kinetics. Hydrolysis of single enantiomers of the substrates (2.5 μM) a) S_1 , b) S_2 , c) S_3 by CuL^1 (50 μM) (● for the *L*-isomer, ○ for the *D*-isomer) and CuL^2 (50 μM) (● for the *L*-isomer, ○ for the *D*-isomer) and spontaneous hydrolysis (the solid line in black). The reactions were performed in a 10% MeCN solution in HEPES buffer (pH 7.2, 50 mM) at 298 ± 0.1 K.

Catalyst	k_{in} (10^{-5} s^{-1})					
	S_1		S_2		S_3	
	<i>L</i> -	<i>D</i> -	<i>L</i> -	<i>D</i> -	<i>L</i> -	<i>D</i> -
CuL^1	3.9	2.1	5.4	5.2×10^{-1}	2.0	1.3
CuL^2	9.5×10^{-1}	7.7×10^{-1}	1.6	7.2×10^{-1}	3.4×10^{-1}	1.9×10^{-1}
L^1	6.2×10^{-1}	4.3×10^{-1}	8.5×10^{-1}	1.8×10^{-1}	2.1×10^{-1}	1.2×10^{-1}
L^2	4.4×10^{-1}	3.1×10^{-1}	4.8×10^{-1}	1.2×10^{-1}	4.3×10^{-1}	1.7×10^{-1}
Cu^{2+}	1.2×10^{-1}		6.7×10^{-2}		6.1×10^{-2}	
Buffer	5.0×10^{-2}		3.2×10^{-2}		2.8×10^{-2}	

Table 1. The initial rate constants for S_1 - S_3 (2.5 μM) promoted by different catalysts (50.0 μM) in a 10% MeCN solution in HEPES buffer (pH 7.2, 50 mM) at 298 ± 0.1 K.

reduced by changing either the substrate or the catalyst (Fig. 2a,c), while S_2 remained the most favored substrate by both CuL^1 and CuL^2 . Additionally, changing the catalyst reduced the enantioselectivity more than that induced by changing the substrate, suggesting that the CD orientation plays a very important role in the enantioselective catalysis. The results of these screening experiments indicated that the “back-to-back” bisCD complex CuL^1 favoured the enantioselective hydrolysis of the amino acid-containing substrates, which was particularly evident for S_2 . The control experiments using the metal or ligands separately resulted in substantially decreased catalytic efficiency and enantioselectivity (Table 1 and Supplementary Figs S9–12).

To fully assess the enantioselective hydrolysis of S_1 - S_3 by CuL^1 , a revised Michaelis–Menten kinetic experiment was performed in the presence of excess catalyst^{27,29,44–46}. By increasing the initial catalyst concentration from 5.0 to 125 μM , the hydrolysis rate increased, resulting in a levelled-off curve for each substrate (Fig. 3) and implying that the catalyst-substrate complex formed prior to the catalytic reaction. The kinetic parameters were deduced by fitting the data to the Michaelis–Menten equation (Table 2)^{47,48}. The best catalytic ability ($k_{\text{cat}}/k_{\text{uncat}}$) was observed for *L*- S_2 , which also showed the most pronounced enantioselectivity, with an enantiomer selectivity (*L/D*) of 15.7. Furthermore, the k_{cat} values for S_1 - S_3 ranged from 3.3 – $6.2 \times 10^{-5} \text{ s}^{-1}$ for the *L*-isomers and 1.6 – $4.2 \times 10^{-5} \text{ s}^{-1}$ for the *D*-isomers. As a reference, the Michaelis–Menten kinetic experiment was also performed to investigate S_2 catalysis by CuL^2 (see Supplementary Fig. S14) and resulted in an enantiomer selectivity of 3.3, as shown in Supplementary Table S1.

To further confirm that the *L*-isomer was preferentially hydrolysed over the *D*-isomer, a chiral high-performance liquid chromatography (HPLC) analysis were performed with racemic S_2 . As shown in Supplementary Fig. S16, after different ratios of racemic S_2 (10.0 μM) were consumed by CuL^1 (100 μM) at 298 K, the reaction aliquots were subjected to ethyl acetate extraction to transfer the unreacted S_2 into the organic phase⁴⁴, followed by chiral HPLC analysis with CHIRALPAK[®] IC. When 35%, 40%, and 50% of S_2 was consumed, the remaining *L:D* ratios were 28:72, 23:77 and 20:80, respectively (Table 3). Based on these conversions and the remaining substrate enantiomeric ratios (e.r.), the consumption ratios of the *L*- S_2 and *D*- S_2 enantiomers were calculated to be 94:6, 92:8, and 82:18, thus indicating that the corresponding hydrolysed product (Boc-Phe-OH) were present at e.r. of 94:6, 92:8 and 82:18 (*L/D*) (Table 3). These results demonstrated that CuL^1 possessed high enantioselectivity for the hydrolysis of racemic S_2 and showed a preference for the *L*-isomers.

Catalytic inhibition assay. To gain more insights into the reaction mechanism, di(*p*-*tert*-butylbenzyl) amine (DBBA) was applied as a CD inhibitor to investigate the role of the CD cavity^{34,36}. In our previous study, we have excluded the possibility that DBBA exerts its inhibitory effect through coordination with the metal ions³⁶. The strong binding between DBBA and the CD cavity was demonstrated by ROESY. The nuclear Overhauser effect (NOE) cross-peaks could be observed as the protons are closer than 0.4 nm in space. Therefore, the

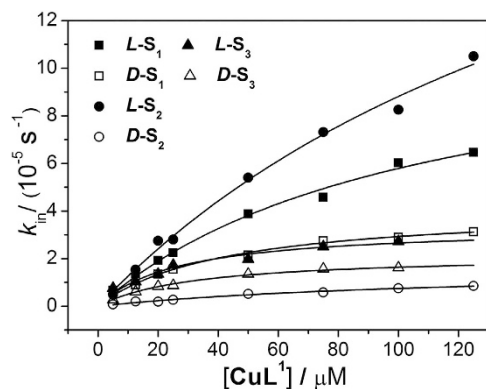


Figure 3. Saturation kinetics. Michaelis–Menten kinetics for the hydrolysis of the substrates (2.5 μM) in the presence of CuL^1 (5.0–25 μM) in a 10% MeCN solution in HEPES buffer (pH 7.2, 50 mM) at 298 ± 0.1 K.

		k_{cat} (10^{-5} s^{-1})	K_{m} (μM)	$k_{\text{cat}}/K_{\text{m}}$ ($\text{M}^{-1} \text{ s}^{-1}$)	k_{uncat} (10^{-7} s^{-1})	$k_{\text{cat}}/k_{\text{uncat}}$	$k_{\text{cat}}^L/k_{\text{cat}}^D$
S_1	L-	12.0 ± 1.0	107.0 ± 16.0	1.1	5.0	2.5×10^2	2.9
	D-	4.2 ± 0.2	42.5 ± 4.3	1.0		8.4×10	
S_2	L-	26.2 ± 4.2	197.0 ± 46.0	1.3	3.2	8.1×10^2	15.7
	D-	1.7 ± 0.3	124.0 ± 32.0	0.1		5.2×10	
S_3	L-	3.3 ± 0.3	28.5 ± 5.6	1.2	2.8	1.2×10^2	1.5
	D-	2.2 ± 0.1	34.6 ± 3.4	0.6		7.9×10	

Table 2. Kinetic parameters for the hydrolysis of S_1 – S_3 (2.5 μM) in the presence of CuL^1 (5.0–125 μM) in a 10% MeCN solution in HEPES buffer (50 mM, pH = 7.2) at 298 ± 0.1 K.

Entry	Conversion (%)	Remaining e.r. of S_2 (L:D) ^[a]	e.r. of Boc-Phe-OH (L:D) ^[b]
1	35	28:72	94:6
2	40	23:77	92:8
3	50	20:80	82:18

Table 3. Chiral HPLC analysis of the conversion of various amounts of racemic S. [a] The enantiomeric ratio (e.r.) values of remaining substrates were determined by HPLC analysis with a chiral stationary phase. [b] e.r. values of products were calculated from the result of HPLC analysis.

binding substituent group within the β -CD cavity could be estimated according to the relative intensity of the cross-peaks⁴⁹. The interactions between the aryl protons of **DBBA** and the protons in the CD of L^1 were observed (Supplementary Fig. S17). As expected, the initial rate of the CuL^1 -catalysed hydrolysis of L-S_2 was dramatically decreased by more than 13-fold in the presence of **DBBA**, suggesting that CuL^1 was significantly inhibited (Fig. 4a,b). Although the hydrolysis rate for D-S_2 was low, a 1.3-fold decrease was still observed in the presence of **DBBA**. The kinetic parameters are displayed in Supplementary Table S2. Moreover, 2D ROESY nuclear magnetic resonance (NMR) experiments of L^1 with *L*- or *D*-Boc-Phe-OH, instead of the CuL^1 and S_2 isomers, respectively, were performed (Supplementary Fig. S18). The ROESY spectrum of L^1 with Boc-Phe-OH displayed NOE cross-peak signals between **H-3,5,6** of β -CD and the protons of the **Phenyl** and **Boc** groups of the Boc-Phe-OH enantiomers, indicating that the two groups could be self-included into the CD cavities of L^1 from the primary side (Note: the binding constants (K_a) by β -CD for benzene⁵⁰ and *t*-BuOH^{51,52} are 120 M^{-1} and 48 M^{-1} , respectively, at 298 K in H_2O). These observations demonstrated that the hydrophobic CD cavities were indeed involved in the catalysis of both isomers.

ESI-MS analysis of the mixture of CuL^1 with each S_2 enantiomer were performed (see Supplementary Figs S19 and S20). The catalyst–substrate (1:1) complexes were detected by positive-ion ESI-MS, and no catalyst–product complexes were observed, thus indicating that the product binds the catalyst more weakly than the substrate. Considering the **NP** group could easily be bound in the β -CD cavity^{38,53}, it leads to the proof that the substrate should be bound by the two cooperative hydrophobic cavities. Moreover, negative-ion ESI-MS confirmed that the hydrolysis products were *N*-Boc-phenylalanine (Boc-Phe-OH) and **NP**, and no Boc-deprotected product was observed. These observations suggested that there was enhanced affinity between the catalysts and substrates and that **NP** was cleaved from the substrates following the hydrolysis.

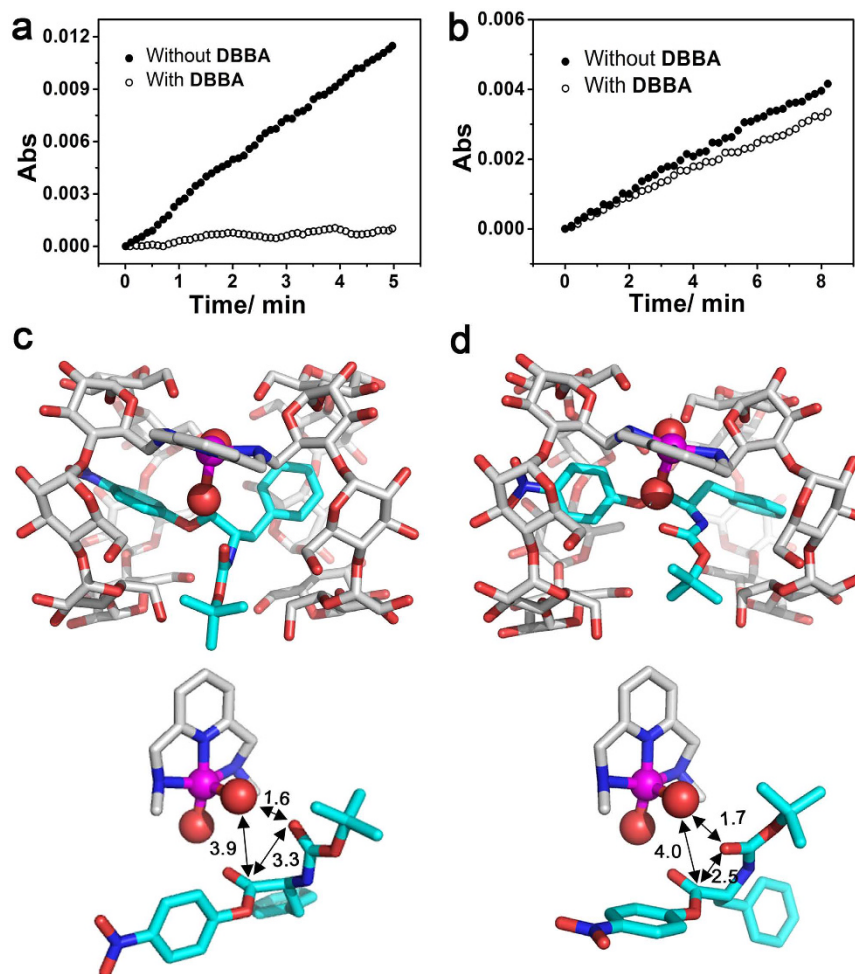


Figure 4. Inhibition reaction and theoretical calculation. Hydrolysis of (a) *L*- S_2 , and (b) *D*- S_2 by CuL^1 in the presence or absence of DBBA in HEPES buffer (pH 7.2, 50 mM) containing 10% MeCN at (298 ± 0.1) K, $[CuL^1] = [S_2] = [DBBA] = 10.0 \mu M$. Optimized structures of the (c) *L*- S_2 - CuL^1 and (d) *D*- S_2 - CuL^1 complexes at the ONIOM(B3LYP/lanl2dz:UFF)/IEFPCM level of theory. Color code: O, red; N, blue; Cu, magenta; C, light grey (catalysts) and cyan (substrates). Hydrogen was omitted. The intramolecular CDs were also omitted in the bottom images for clarity. The metal ions and coordinated water are shown as spheres. All labelled distances are shown in angstroms (\AA).

Based on the above results, we speculate that NP ($K_a = 1503 M^{-1}$ (β -CD, 298 K, H_2O))⁵⁴ and the Phenyl groups of S_2 were encapsulated in the cavities of CuL^1 during the reaction. The enantioselectivity originated from the different geometries of the substrates that were regulated by the closely linked CD chiral cavities. The mechanism was further investigated through theoretical calculations. The structures of the *L*- S_2 - CuL^1 and *D*- S_2 - CuL^1 complexes were optimized at the ONIOM(B3LYP/lanl2dz:UFF)/IEFPCM level of theory (Fig. 4c,d). In both cases, the two aromatic side arms were encapsulated within the CD cavities, a result in good agreement with the results of the inhibition assay. The two isomers with different conformations were regulated by the CD chiral cavities. A major difference in the spatial distance between the ester carbonyl carbon and the flanking carbonyl oxygen was observed. In the case of *L*- S_2 , the distance was 3.3 \AA (Fig. 4c), and the ester carbonyl carbon was exposed to the metal-coordinated nucleophile. In the case of *D*- S_2 , the flanking carbonyl oxygen was oriented toward the ester moiety, with a shorter distance of 2.5 \AA (Fig. 4d), which probably hindered the nucleophilic attack from the metal-coordinated nucleophile because of possible negative charge repulsion. This difference should lead to the enantioselective hydrolysis of *L*- S_2 over *D*- S_2 by CuL^1 .

Discussion

The initial-rate kinetic study showed that the “back-to-back” bisCD complex CuL^1 exhibited higher catalytic efficiency and more pronounced enantioselectivity than the “face-to-face” analogue CuL^2 . These results indicated that the closely linked CD cavities on the primary face exhibit a much stronger cooperative effect with the coordinated copper ion in terms of differentiating the chiral substrates than those linked on the secondary face. Moreover, in most cases, the enantioselectivity was significantly decreased when only L^1 or L^2 was used without the metal. S_3 enantiomers were an exception for L^2 , because containing the bulk R group, S_3 would have an

unsuitable location on L^2 induced by the Cu^{2+} coordination with the bridge, as observed for the metal-inhibited enzymes³⁹. Besides, in the presence of either CuL^1 or CuL^2 , the initial hydrolysis rates of substrate L -isomers exceeded those of D -isomers in all cases, indicating that the chiral CD cavities in our models had an overall preference for the L -isomers of the amino acid esters. EtOH was tested as another cosolvent and exhibited an enantioselectivity similar to those of the reactions performed in MeCN. This finding indicated that a cosolvent effect of MeCN could be excluded (see Supplementary Fig. S13 and Table S4).

The k_{cat}/k_{uncat} value of 811 for $L-S_2$ was in agreement with the largest values obtained for previously reported CD-based or other simple artificial mimics with similar substrates^{26–28,55,56}, although some protein-based mimics produced higher values^{44,57–59}. The most pronounced enantioselectivity was observed for the catalysis of S_2 by CuL^1 , with an enantiomer selectivity of 15.7. This value is comparable to that of the reported natural protein-based model⁴⁴. Although some reports have obtained better enantioselectivity with protein-modified catalysts, the complicated structure has made preparing the catalyst and performing mechanistic studies challenging^{29,30,57–59}. The chiral HPLC analysis of S_2 hydrolysis by CuL^1 confirmed that the L -isomer was preferentially hydrolysed over $D-S_2$, with an e.r. of 94 : 6 (L/D) when 35% of the racemic S_2 was consumed. Finally, the S_2 enantiomers were determined to be the optimal substrates for CuL^1 in our study, and the removal of the methylene on the **Phenyl** group or replacement with an **Indolyl** group diminished the enantioselectivity. CuL^2 also exhibited the highest enantioselectivity for S_2 compared with the other two pairs of enantiomers, although the cooperative effect was weaker than that of CuL^1 . Together, our results suggest that the substrate structure significantly affected the enantioselectivity, and S_2 was identified as the best substrate among those tested. The findings indicated that the catalysts recognize specific substrates.

ESI-MS proved that the catalyst formed an intermediate complex with the substrates. The 2D ROESY NMR and inhibition assay demonstrated the vital role of CDs in the catalytic pathway. Finally, the optimized structures of the catalyst-substrate complexes provided initial evidence that the enantiomers formed different geometries according to the closely linked CD chiral cavities.

In conclusion, we presented a hydrolase mimetic study focusing on chiral substrates, in which two bisCD-based copper(II) complexes were developed for the hydrolysis of Boc-protected amino acid esters under physiological conditions (pH = 7.2). Accordingly, the adjacent chiral CD cavities had an overall preference for binding L -isomers. In addition, the “back-to-back” bisCD complex CuL^1 exhibited much better catalytic efficiency and enantioselectivity towards the chiral amino acid esters than the “face-to-face” analogue CuL^2 . Mechanistic studies showed that the two closely linked chiral CD cavities played a vital role in mediating the enantioselective hydrolysis by regulating the different isomer geometries during the reaction, in which the hydrophobic groups of the substrate were embedded in the two intramolecular CD cavities. These findings indicate a cooperative effect of the first and second coordination spheres of artificial catalysts on the enantioselectivity and also provide hints that may guide future exploration of enzyme mimics.

Methods

Materials. Reagent-grade β -CD was recrystallized twice from H_2O and dried in vacuo for 12 h at 373 K. All of the amino acid esters or precursor enantiomers, N -Boc-phenylalanine 4-nitrophenyl ester (S_2), N -Boc-phenylglycine (Boc-Phg-OH), and N -Boc-Tryptophan (Boc-Trp-OH) were purchased from GL Biochem (Shanghai) Ltd. Dicyclohexylcarbodiimide (**DCC**) was purchased from Aladdin. Di(*p*-tert-butylbenzyl) amine (**DBBA**) was synthesized with previously reported methods³⁴. DMF was superdry grade and stored over a molecular sieve. Common organic reagents were reagent grade and redistilled before use. Milli-Q water was used in all physical measurements. CuL^1 and CuL^2 were synthesized as described in previous studies^{35,38,39}. All compounds were confirmed by elemental analyses, ESI-MS, and 1H -NMR spectra.

General methods. The 1H NMR spectra were recorded on Mercury plus 300 spectrometers. The 2D NMR spectra were recorded on a Bruker AvanceIII 600 spectrometer. The elemental contents were analysed with a Perkin-Elmer 240 elemental analyser. The ESI-MS spectra were collected on a Thermo LCQ-DECA-XP spectrometer. The UV/Vis spectra were monitored with a Varian Cary 100 UV/Vis spectrophotometer equipped with a temperature controller (± 0.1 K). The HPLC analyses were performed on an Agilent 1200 HPLC with CHIRALPAK[®] IC 250 \times 4.6 mm column. The EPR spectra were recorded on a Bruker A300-10-12 spectrometer.

Synthesis of the amino acid esters. The amino acid esters were prepared according to a previously described method⁶⁰, with some modifications. Boc- L -Phg-OH (3.5 g, 13.9 mmol) and 4-nitrophenol (1.9 g, 13.9 mmol) were dissolved in dry DMF (15 mL). Aliquots of the mixture were added to a stirred and cooled ($-10^\circ C$) suspension of **DCC** (2.9 g) in dry DMF (50 mL). The reaction mixture was magnetically stirred at $-10^\circ C$ for approximately 2 h and then allowed to stand at room temperature for 5 h. The reaction mixture was filtered, and the filtrate was evaporated to dryness. The crystallization of the residue from 1:12 (*v/v*) ethyl acetate/petroleum ether produced pure Boc- L -Phg-ONp ($L-S_1$) with a 78% yield. 1H -NMR (300 MHz, $CDCl_3$): 8.22 (1H, d, J 9.2, NP-H), 7.42 (2H, dd, J 9.3, 3.6, Ph-H), 7.20 (1H, d, J 9.0, NP-H), 5.52 (1H, d, J 6.9, C-NH), 5.43 (1H, d, J 7.0, CH-N), 1.47 (4H, s, CH_3); elemental analysis (calcd., observed for $C_{19}H_{20}N_2O_6$): C (61.28, 61.51), H (5.41, 5.44), N (7.52, 7.52).

$D-S_1$ was prepared through the same procedure. 1H -NMR (300 MHz, $CDCl_3$): 8.22 (1H, d, J 9.2, NP-H), 7.42 (2H, dd, J 9.3, 3.6, Ph-H), 7.20 (1H, d, J 9.0, NP-H), 5.52 (1H, d, J 6.9, C-NH), 5.43 (1H, d, J 7.0, CH-N), 1.47 (4H, s, CH_3); elemental analysis (calcd., observed for $C_{19}H_{20}N_2O_6$): C (61.28, 60.94), H (5.41, 5.48), N (7.52, 7.52).

Boc- L -Trp-ONp ($L-S_3$) and Boc- D -Trp-ONp ($D-S_3$) were synthesized as described above. $L-S_3$: 1H -NMR (300 MHz, $CDCl_3$) 8.17 (3H, d, J 9.1, NP-H, Indole-H), 7.58 (1H, d, J 7.8, Indole-H), 7.40 (1H, d, J 8.1, Indole-H), 7.26–7.20 (3H, m, Indole-H, $CHCl_3$), 7.17–7.08 (2H, m, NP-H), 6.98 (2H, d, J 9.1, Indole-H), 5.14 (1H, d, J 7.5, C-NH), 4.85 (1H, dd, J 13.5, 5.9, CH-N), 3.42 (2H, qd, J 14.3, 6.4, CH_2), 1.45 (9H, s, CH_3); elemental analysis (calcd., observed for $C_{22}H_{23}N_3O_6$): C (62.11, 62.34), H (5.45, 5.50), N (9.88, 9.77).

D-S₃: ¹H-NMR (300 MHz, CDCl₃) 8.18 (3 H, d, J 9.0, NP-H, Indole-H), 7.60 (1 H, d, J 7.9, Indole-H), 7.40 (1 H, d, J 13.4, Indole-H), 7.29–7.21 (3 H, m, Indole-H, CHCl₃), 7.13 (2 H, t, J 7.3, NP-H), 6.99 (2 H, d, J 9.0, Indole-H), 5.15 (1 H, d, J 7.7, C-NH), 4.86 (1 H, d, J 7.1, CH-N), 3.59–3.29 (2 H, m, CH₂), 1.46 (9 H, s, CH₃); elemental analysis (calcd., observed for C₂₂H₂₃N₃O₆): C (62.11, 62.11), H (5.45, 5.51), N (9.88, 9.81).

Kinetics of amino acid ester hydrolysis. The hydrolysis rates of the amino acid esters in the presence of CuL¹ and CuL² were measured with an initial slope method by monitoring the increase in the 400-nm absorption of the released NP. At this wavelength, the absorbance of the ester substrate was negligible. The reaction solution was maintained at 298 ± 0.1 K. HEPES buffer (pH = 7.2, 50 mM) was used, and the ionic strength was adjusted to 0.10 M with NaClO₄. The amino acid esters were prepared as solutions in CH₃CN, and the buffer and catalyst solutions were freshly prepared in water.

In a typical experiment, a small amount of aqueous catalyst solution was spread evenly in a 600-μL cuvette containing 10% (v/v) CH₃CN buffer solution, and the reactions were always initiated by injecting a small amount of substrate CH₃CN solution, followed by complete mixing. All of the solutions were equilibrated to 298 ± 0.1 K. The initial first-order rate constants (*k*_{in} (s⁻¹)) for the substrate cleavage were obtained directly from a plot of the NP concentration versus time by using the initial rate method. The errors in the *k*_{in} values were less than 5%. The ε value for NP at pH 7.2 was 10,398. The inhibition study was performed by following a similar procedure, except for the pre-equilibrium incubation after mixing DBBA with CuL¹.

Kinetic resolution of racemic S₂. The hydrolysis of racemic S₂ was performed using a method similar to that described in the previous section. Racemic S₂ (10.0 μM) was mixed with CuL¹ (100 μM) under the same conditions as those used in the kinetic studies. The reaction process was monitored with a UV/Vis spectrometer. Then, upon mixing of the reaction mixture with ethyl acetate, the products (Boc-Phe-OH and 4-nitrophenol) and remaining S₂ were completely extracted into the organic phase⁴⁴. The organic phase was evaporated to dryness under reduced pressure, and the resulting residue was dissolved in ethanol for HPLC analysis. The chiral HPLC analyses were performed at room temperature with a CHIRALPAK[®] IC 250 × 4.6 mm column and monitored by a UV-Vis detector at 254 nm; the elution solvent was 40/30/30 H₂O (0.1% formic acid)/CH₃CN/EtOH, and the flow rate was 1.0 mL/min.

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Author Contributions

S.-S.X., M.Z., J.W., Z.-W.M. and L.-N.J. conceived and designed the experiments. S.-S.X., Z.-F.K., B.-C.C., H.S. and Z.-K.C. performed the experiments. S.-S.X., J.W. and Z.-W.M. analyzed the data. S.-S.X., M.Z., J.W. and Z.-W.M. wrote the paper. All authors discussed the results and commented on the manuscript.

Additional Information

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