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Operando X-ray photoelectron spectroscopy of solid electrolyte interphase formation and evolution in $\text{Li}_2\text{S-P}_2\text{S}_5$ solid-state electrolytes

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Solid-state electrolytes such as $\text{Li}_2\text{S-P}_2\text{S}_5$ compounds are promising materials that could enable Li metal anodes. However, many solid-state electrolytes are unstable against metallic lithium, and little is known about the chemical evolution of these interfaces during cycling, hindering the rational design of these materials. In this work, operando X-ray photoelectron spectroscopy and real-time in situ Auger electron spectroscopy mapping are developed to probe the formation and evolution of the Li/ $\text{Li}_2\text{S-P}_2\text{S}_5$ solid-electrolyte interphase during electrochemical cycling, and to measure individual overpotentials associated with specific interphase constituents. Results for the Li/ $\text{Li}_2\text{S-P}_2\text{S}_5$ system reveal that electrochemically driving Li^+ to the surface leads to phase decomposition into Li_2S and Li_3P . Additionally, oxygen contamination within the $\text{Li}_2\text{S-P}_2\text{S}_5$ leads initially to Li_3PO_4 phase segregation, and subsequently to Li_2O formation. The spatially non-uniform distribution of these phases, coupled with differences in their ionic conductivities, have important implications for the overall properties and performance of the solid-electrolyte interphase.

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As global energy consumption continues to increase rapidly, scalable, safe, and cost-effective strategies for energy storage have become imperative. In pursuit of this goal, numerous ‘beyond Li-ion battery’ solutions have been intensively researched over the past decade^{1, 2}. Many of these next-generation battery architectures use Li metal anodes, which enable dramatically improved theoretical energy densities (gravimetric and volumetric) compared to the current state-of-the-art. A significant challenge is that the extreme reactivity of Li metal tends to cause undesirable side reactions between the Li metal and the electrolyte³. In the case of liquid electrolyte systems this often leads to Li consumption, dendrite formation, and the potential for catastrophic failure and fires^{4–7}.

One widely studied approach for improving the safety of next-generation batteries is the use of solid-state electrolyte (SSE) materials⁸. SSEs improve safety by eliminating flammable liquids, and by providing a physical barrier to dendrite propagation. On the other hand, SSE conductivities are typically lower than liquid electrolytes and nearly all SSE materials are unstable against Li metal. This instability results in degradation of SSE/Li interfaces, creating large interfacial resistances that severely compromise battery performance⁹. Therefore, the interfacial interactions between Li and the SSE must be thoroughly investigated to enable the rational design of stable next-generation battery interfaces. Unfortunately, because of the numerous challenges associated with performing detailed chemical analyses on these types of interfaces^{9–12}, to date few experimental studies have been reported on their chemical composition and structure.

A current state-of-the-art SSE is sputtered lithium phosphorus oxynitride (LiPON). This material is known for its moderate interfacial stability^{9, 13, 14}, but relatively poor room-temperature ionic conductivity ($\sim 10^{-6}$ S/cm)¹⁵. In order to compensate for its low ionic conductivity, much research on LiPON focuses on thin-film fabrication techniques and applications. To circumvent the synthesis challenges of thin film SSEs, researchers have begun to explore higher conductivity systems, like $\text{Li}_{10}\text{GeP}_2\text{S}_{12}$ (LGPS) and $\text{Li}_2\text{S-P}_2\text{S}_5$ (LPS) where conductivities exceeding 5 mS/cm have been demonstrated^{16, 17}. Although there is great promise in these sulfur-based SSEs, these highly reactive materials have unstable interfaces with Li, leading to worse rate capabilities than LiPON, even though the initial bulk conductivities are higher^{18, 19}. Therefore, the future of sulfur-based SSEs depends on engineering more stable SSE/Li interfaces that enable both high rate capability and extended cycle life.

To optimize these highly conductive SSEs and/or design interfacial barrier materials that enable next-generation SSE battery architectures, a critical first step is understanding how the solid electrolyte interphase (SEI) forms and evolves both chemically and morphologically³. Unfortunately, since the SEI is a buried interface (and therefore not readily accessible to the majority of standard analytical techniques) these issues are difficult to elucidate experimentally. A practical challenge that constrains interfacial battery characterization experiments is the extreme reactivity of Li metal, the SSE, and even SEI phases to oxygen, moisture, and organic species. These reactivities limit the utility of typical preparation methods (e.g., focused ion-beam milling, mechanical polishing, etc.) for studying buried interfaces, because such methods can damage, smear, or otherwise fundamentally alter these highly reactive interfaces²⁰. Preparing a sample for characterization where the SSE/Li interface is representative of the chemical reactions occurring during operation is extremely challenging and resulting artifacts might limit the utility of such data.

Recently, an in situ X-ray photoelectron spectroscopy (XPS) study by Wenzel et al. helped elucidate the reaction between Li and single-crystal $\text{Li}_7\text{P}_3\text{S}_{11}$ by providing chemical state

information before, during, and after physical vapor deposition of Li metal on the surface¹¹. These measurements clearly showed that $\text{Li}_7\text{P}_3\text{S}_{11}$ reacts with Li metal to form Li_2S and Li_3P decomposition phases. However, samples in this study were not subjected to electrochemical bias during XPS measurements, and the chemical composition and evolution of the SEI under battery operating conditions was not probed. Therefore, to shed light on how the SEI evolves chemically during cycling, we provide operando measurements of an Li/LPS interface. In this experiment, an electron gun bias drives Li^+ migration, while XPS measurements monitor changes at the exposed surface. These data provide detailed compositional and chemical-state information related to SEI formation at the Li/LPS interface. Furthermore, Li^+ migration in the cell can be reversed by using ultraviolet (UV) light to extract valence electrons from the surface and create a driving force for Li^+ migration away from the surface back toward the other electrode. Our results reveal a chemically layered SEI configuration that evolves during cycling. Furthermore, the operando XPS (opXPS) measurements directly reveal the overpotentials of individual phases in the SEI, providing clues as to which SEI components limit battery performance.

Results and Discussion

Electrochemical cycling and ex situ XPS analysis. To document the composition of an electrochemically formed SEI at an LPS interface, a Cu/LPS/Li coin cell was assembled and biased to drive Li^+ toward the Cu electrode. After 1.5 h of galvanostatic charging at 0.17 mA/cm^2 (Fig. 1a) the coin cell was disassembled in a glovebox and transferred through an interconnected ultrahigh-vacuum transfer system to an XPS system. The XPS spectra are shown Fig. 1b–e. The chemical states observed in these XPS spectra are in good overall agreement with the results from Wenzel et al.¹¹. Electrochemical impedance spectroscopy (EIS) data sets acquired during the first few cycles of a sister sample, Supplementary Figure 1 (and Supplementary Note 1), demonstrate that the SEI continues to evolve during cycling, especially during the first few cycles. Therefore, understanding how these SEI components evolve during cycling and what effects they have on battery performance and stability are essential.

Virtual electrode cycling and operando XPS. XPS core-level spectra acquired from the initial LPS surface showed the presence of Li, S, P, C, and O. XPS depth profile data revealed that oxygen contamination ($\sim 3\text{--}7$ atomic %) existed throughout the LPS samples tested. Consequently, oxygen was present on sample surfaces even after sputter cleaning, while no significant carbon contamination was observed (<0.5 atomic %). Figure 2a shows the Li 1s core level from the sputter-cleaned LPS surface. After a few minutes of charging via the virtual electrode (Fig. 2b) a clear shift in the Li 1s core-level is observed. As described in more detail in the following section, this shift results from the chemical transformation of LPS into SEI phases at the analysis surface, driven by Li^+ migration. As charging continues (>5 h in Fig. 2c) metallic Li appears at the surface, as evidenced by the appearance of a low-BE feature in the spectra.

For discharge, as shown in Fig. 2d, exposure to UV photons photoionizes metallic Li according to:



where e_{PE}^- denotes the outgoing photoelectron. The resulting accumulation of positive charge drives Li^+ away from the free surface and toward the opposite electrode so that metallic Li is stripped and the SEI is uncovered (Fig. 2e). As the SEI is

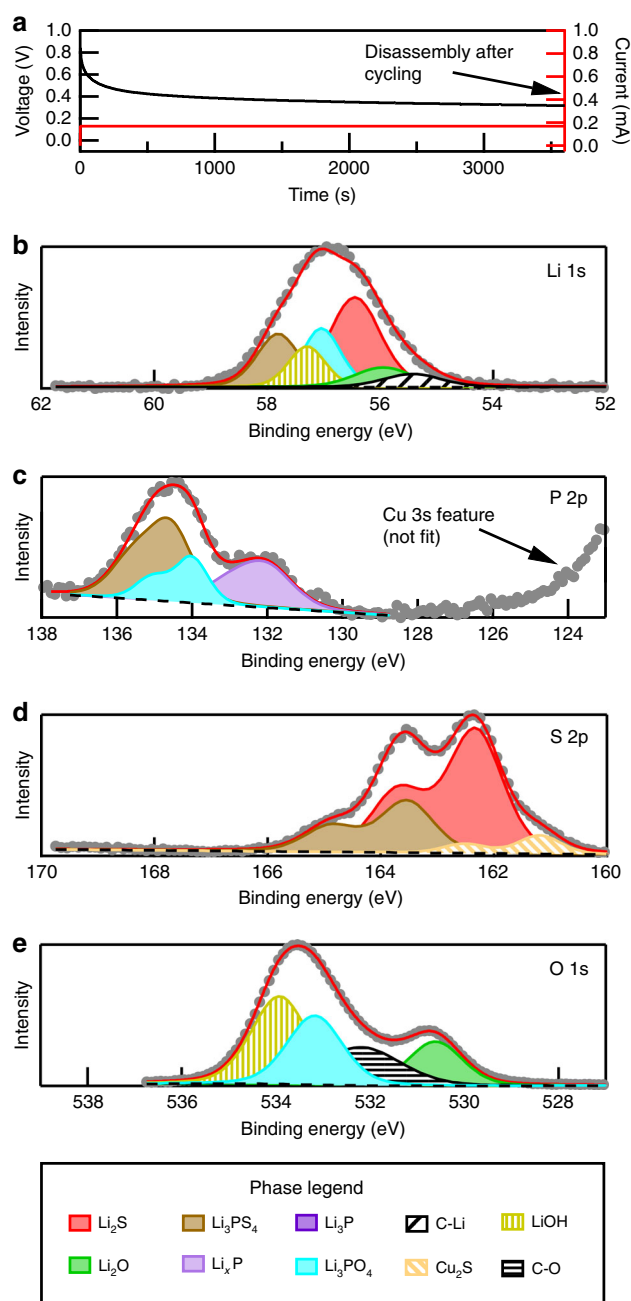


Fig. 1 Ex situ XPS spectral decomposition of Li/LPS SEI. **a** Galvanostatic voltage profile of Cu/LPS/Li coin cell during initial charging. The cell was disassembled after 1.5 h of Li deposition onto the Cu electrode. Ex situ XPS spectra are shown for the Li 1s (**b**), P 2p (**c**), S 2p (**d**), and O 1s (**e**) core levels

uncovered, spectra similar to those in Fig. 2b prior to the onset of Li^0 plating are observed. Eventually minimal further changes are observed in the Li 1s region, but it is noteworthy that the spectrum never returns to the initial state, indicating as expected that some SEI components are irreversible.

To provide chemical-state information related to SEI formation and evolution at the LPS/Li interface, the Li 1s and other core-level spectra shown in Fig. 3 have been decomposed into contributions from individual SEI phases. Supplementary Table 1 (and Supplementary Note 2) describes the fitting constraints used to obtain the information detailed below. As seen in Fig. 3a, the initial LPS surface is comprised almost entirely of documented

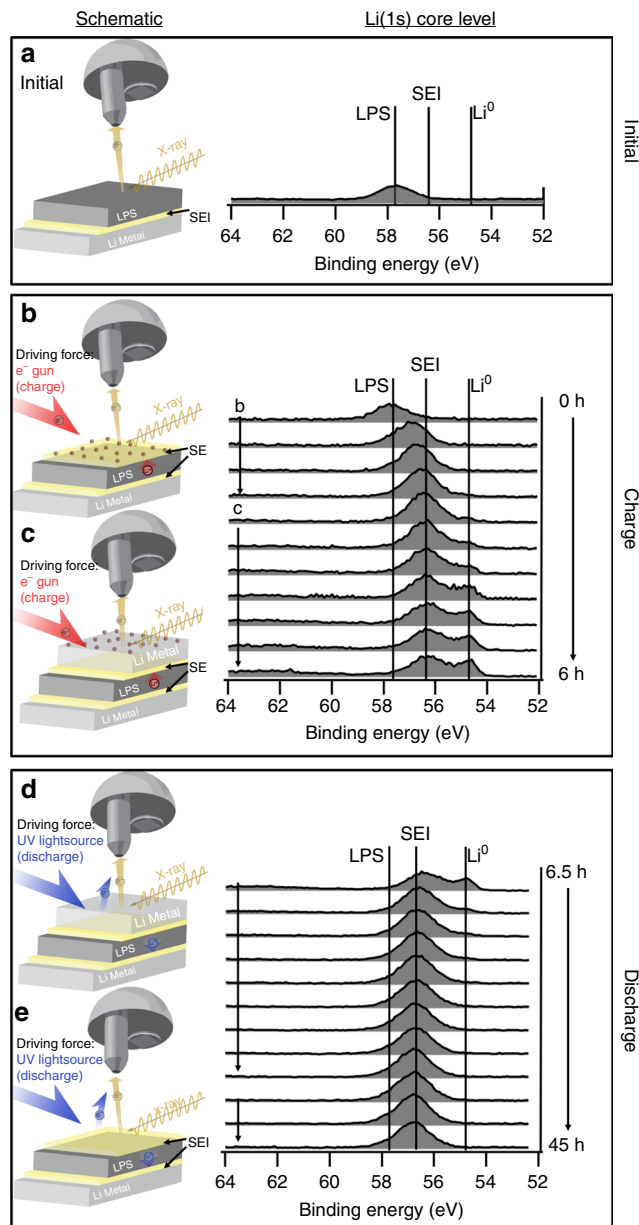


Fig. 2 opXPS schematic and Li 1s core level evolution. **a** Clean LPS surface, **b, c** SEI formation during charging, and **d, e** SEI evolution during discharging

LPS functional groups¹¹, except for a small amount of oxygen contamination. XPS quantification using tabulated elemental XPS sensitivity factors indicates that the O/(S + O) ratio is 0.14. We speculate that initially this oxygen exists primarily as isoelectronic substitutional defects on the LPS anion sublattice, O_s (brown shaded traces in Fig. 3), but cannot rule out contributions from phase-separated Li_3PO_4 (LPO) domains. As will be discussed, oxygen-containing compounds are non-negligible components of the Li/LPS SEI. Also, recent studies have suggested that substitutional oxygen defects in LPS might enhance ionic conductivity and/or improve SEI properties^{21–24}. For these reasons it is important to understand the role of oxygen in the formation and evolution of the SEI.

After 0.5 h of charging (see Fig. 3b) significant changes at the surface occurred as evidenced by the appearance of new chemical states in both the S 2p and P 2p core levels. These additional peaks indicate transformation of the P and S species at the surface

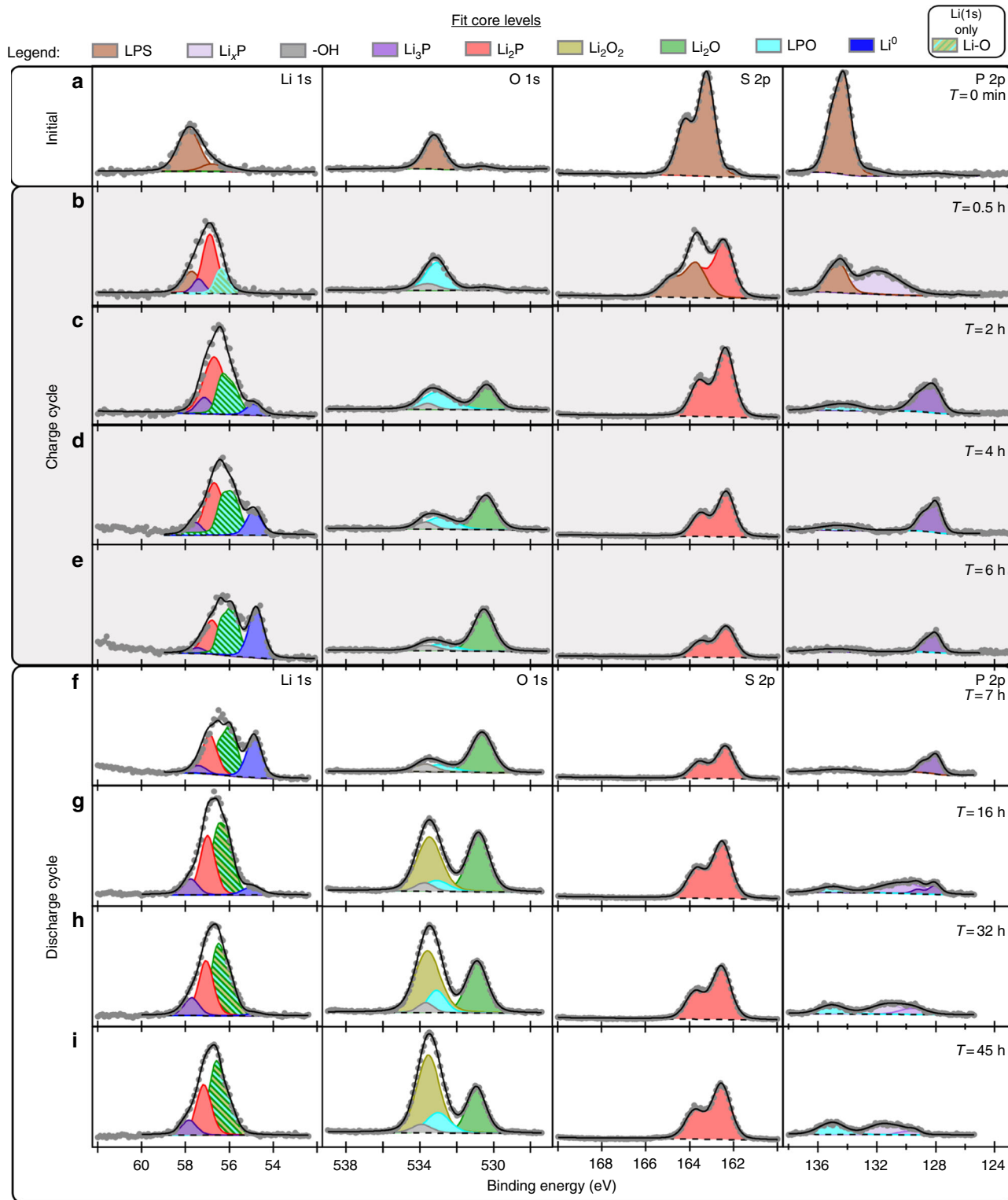


Fig. 3 opXPS spectral decomposition of Li/LPS SEI for the first charge-discharge cycle. XPS spectra showing curve-fitting results for: **a** spectra from the sputter-cleaned LPS surface; **b-e** four sets of spectra during the charging step; and **f-i** four sets of spectra during the discharge half cycle. Due to existence of multiple O-containing species that can be fit in the Li 1s spectra, a single peak representing the summation of O-containing Li 1s components is shown with crosshatched colors of the various contributing components. For both the S 2p and P 2p core levels the $2p_{3/2}$ and $2p_{1/2}$ peaks have been combined into single lineshapes for simplicity. However, it should be noted that the appropriate intensity-ratio and BE separation constraints were applied during curve fitting of these features

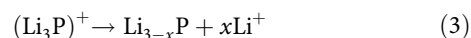
from the original LPS functional groups (i.e., Li-S-P, P-S-P, and P = S) to Li₂S (shaded red) and Li-deficient Li-P functionalities (Li_{3-x}P; shaded light purple). During this initial phase of charging, a slight shift to lower BE was also observed for the O 1s peak. This new BE position for the O 1s closely matches the baseline spectra acquired from Li₃PO₄ (LPO; shaded light blue in Fig. 3b). Therefore, we tentatively conclude that phase segregation [$4\text{Li}_3\text{PO}_x\text{S}_{4-x} \rightarrow (4-x)\text{Li}_3\text{PS}_4 + x\text{Li}_3\text{PO}_4$] begins to occur at the surface during the charging step. It is also reasonable to conclude that after 0.5 h without sputtering, a submonolayer coverage of -OH groups exist on the surface (shaded gray in Fig. 3b), as a result of trace residual H₂O vapor in the analysis chamber.

After 2 h of charging (Fig. 3c), the LPS functional groups are completely attenuated. At the same time, Li_{3-x}P transforms to Li₃P (shaded dark purple), Li₂S increases slightly, dramatic changes in the O 1s spectra are observed, and a small amount of metallic Li (Li⁰; shaded dark blue) appears. The striking changes in the O 1s spectra can be summarized by three observations: (i) a slight increase in the total intensity of the O 1s spectra; (ii) the appearance of a new low BE feature, assigned as Li₂O (shaded green); and (iii) a decrease in signal from the Li₃PO₄ (LPO) phase. The origin of the slight increase in oxygen content is somewhat ambiguous, as it is possible that Li⁰ can act as a getter material for oxygen and water residual gases in the chamber^{25, 26}. However, the relatively small changes in the O 1s spectra throughout the first 0.5 h indicate that a significant amount of oxygen originated from within the sample itself rather than from background chamber contamination, as discussed in more detail later. Most importantly these results demonstrate that oxygen is an important constituent of the Li/LPS SEI, because it is difficult or impractical to remove all oxygen from the bulk and surface of these highly reactive materials during synthesis, processing, and cell fabrication. At this point during the first charge half cycle, the net charge is similar to that for the ex situ XPS analysis illustrated in Fig. 1. To demonstrate the similarities between these spectra, the XPS spectra for the cycled coin-cell and virtual-electrode samples are plotted together in Supplementary Figure 2 (also see Supplementary Note 3). The remarkable agreement between these spectra supports the validity of the virtual-electrode operando approach as a method for probing SEI evolution during electrochemical cycling. However, the comparisons of these spectra also reveal the appearance of an additional feature in the S 2p spectra for the cycled coin-cell sample that is attributed to the presence of Cu₂S. We conclude this phase results from side reactions between the LPS and Cu electrode.

As charging continued beyond 2 h the Li⁰ feature increased in intensity as Li metal began to accumulate at the surface (Fig. 3d, e). This also led to attenuation of the Li₂S and Li₃P signals (a consequence of the limited information depth of XPS measurements, ~5–10 nm). Due to disagreements in the literature on the reported BE value for Li⁰, a sputter-cleaned Li metal reference sample is shown in the Supplementary Information as a baseline reference (Supplementary Figure 3 and Supplementary Note 4). After charging for 6 h, a 0.5 h rest period elapsed prior to initiation of the discharge cycle.

During the initial minutes of discharging the only significant spectral change was a decrease in the amount of Li metal present at the surface and a slight transition from low-BE O 1s functional groups to high-BE species (Fig. 3f). After the majority of metallic Li had been stripped from the analysis surface (16 h, Fig. 3g), more pronounced changes in the SEI were observed for all core levels. For the S 2p core level, the Li₂S peak became more intense, but the BE remained relatively unchanged. An overall increase in P 2p intensity was also observed, however the Li₃P peak decreased in intensity as Li_{3-x}P peaks appeared, indicating at least partial

redox activity of the Li-P phases in the SEI under Li-deficient conditions. Similarly, the Li₂O component of the SEI also appears redox active, as the low BE portion of the O 1s spectra was converted into higher BE species. While there are several possible mechanisms for this conversion, perhaps the most likely is a reaction between Li-P and Li₂O where Li₂O₂ (shaded yellow) and Li₃PO₄ (shaded light blue) are formed. We speculate that UV exposure photoionizes Li₃P according to Eq. (2), leading to electrochemical reduction and generation of Li⁺ according to Eq. (3). The Li-deficient Li-P phase (Li_{3-x}P) then chemically reacts with Li₂O to form Li₂O₂ and Li₃PO₄ according to Eq. (4). Both Eqs. (3) and (4) are accompanied by migration of Li⁺ out of the SEI through the LPS. (For Eq. (4), additional photoionization events according to Eq. (1) also occur.) The net reaction is shown in Eq. (5).



Upon continued discharge (~32 h; Fig. 3h), the Li₂O₂ component remains relatively constant while the Li₃PO₄ phase grows. This coincides with a complete transformation of the Li₃P into Li_{3-x}P, consistent with the proposed mechanism. The decrease in Li₃P also demonstrates that Li₃P is a redox-active component of the SEI that can be delithiated during discharge under Li-deficient conditions. By comparison, the Li₂S peak appears almost completely unchanged from Fig. 3g, again indicating it is a stable component of the SEI. By 45 h of discharge (Fig. 3i) the Li⁰ is completely gone and no other significant changes are observed.

Evolution of SEI phase composition during cycling. In addition to the observed chemical-state changes, the order of appearance and attenuation during charging (or intensification during discharging) of SEI phases reveals an apparent layered configuration of the SEI. We also note that the underlying LPS composition and/or SEI formation mechanisms are likely laterally heterogeneous. Consequently, spatial heterogeneities probably also exist within the SEI. Therefore, when referring to a ‘layered configuration’ we do not intend to convey that the SEI is uniform and homogeneous. Rather, this terminology indicates that certain phases are more likely to form after (and on top of) other phases are already present at the interface.

As illustrated in Fig. 4, changes in the percentage phase compositions during the charge half cycle indicate that the general trend in the SEI layered structure (top to bottom) is Li⁰/Li-O/Li-P/Li₂S/LPS. Before charging, ~5% oxygen was observed in the LPS sample. After a short period of charging the amount of oxygen present in the sample remained nearly constant, whereas the Li_{3-x}P and Li₂S phase contents increased dramatically, from 0 → 12 atm% and 1 → 50 atm%, respectively. This indicates that initially the LPS surface reacts with the incoming flux of Li⁺ at the surface and decomposes into Li_{3-x}P and Li₂S. From Fig. 4, Li₂S formation appears to stop first, as the signal begins to attenuate after approximately 2 h. The P containing species (Li₃P + Li_{3-x}P) exhibit a slight increase in total concentration until around ~3.5 h when the Li₃P phase begins to attenuate. Subsequently, the Li₂O signal becomes dominant, and the other

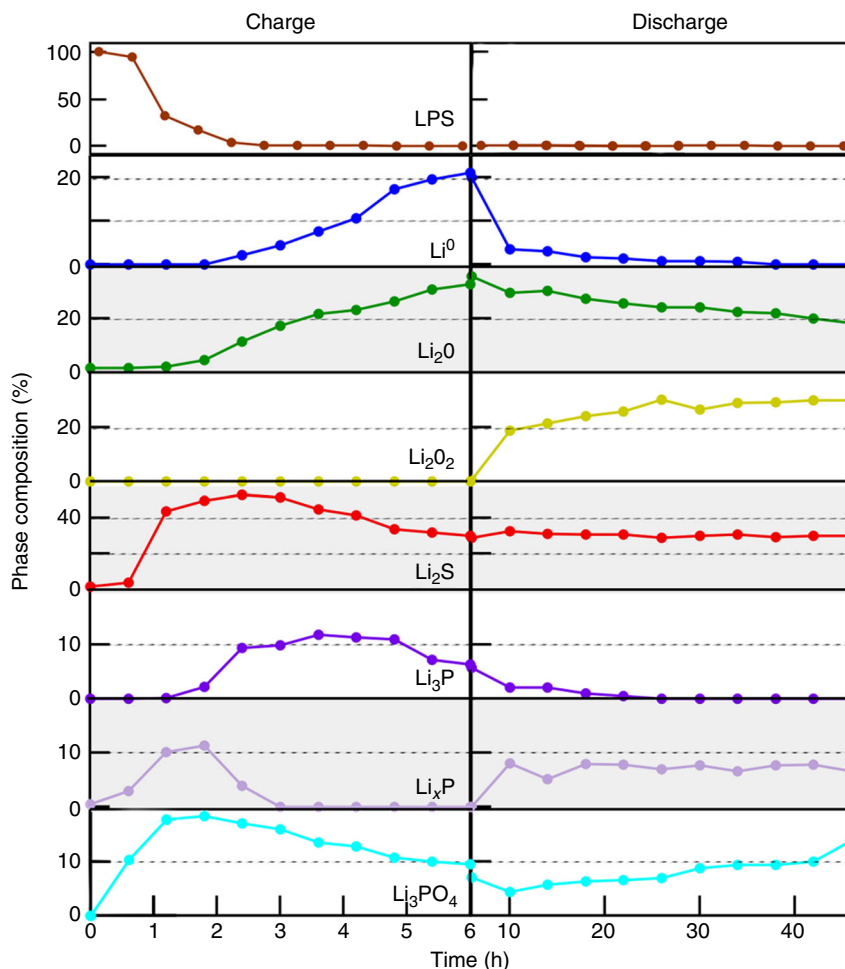


Fig. 4 Evolution of SEI phase composition. Phase composition as a function of time for both the charge and the discharge half-cycles. Data from this graph clearly show the redox activity of the Li_2O and Li_3P components of the SEI during the discharge half-cycle

signals attenuate, indicating that Li_2O forms on top of the $\text{Li}_2\text{S}/\text{Li}_3\text{P}$ layers. Eventually Li^0 appears at the surface, and near the end of the charge step the increase in Li_2O levels off as Li metal begins to attenuate all signals from the SEI (Supplementary Figure 4 and Supplementary Note 5).

After the 0.5 h rest period, the virtual electrode polarity was reversed, and within 4 h (10 h total), significant changes in phase composition were observed. This effect was not seen in control experiments in the absence of UV light bias (Supplementary Figures 5 and 6 and Supplementary Note 6 and 7). During the discharge, as Li^0 was removed from the surface, the total peak intensities (Fig. 3) of O, S, and P species on the surface increase, consistent with an uncovering of the SEI. While all elements in the SEI increase in concentration, both the Li_2O and Li_3P phases decrease in concentration, as other Li-deficient phases appear, namely $\text{Li}_2\text{O}_2/\text{Li}_3\text{PO}_4$ and Li_{3-x}P , respectively. This highlights that some phases within the SEI are redox active, and more specifically the appearance of Li-deficient phases indicates that Li can be removed from some SEI phases (e.g., Li_3P and Li_2O) during the discharge process. In contrast, after 4 h of discharging the Li_2S phase reaches a maximum concentration, and no BE shifts occur, nor do additional S-containing phases appear. For the remainder of the discharge half cycle, this phase remains stable, indicating that Li_2S is not delithiated, at least under the modest potentials applied during discharge in these experiments.

After ~20 h, no Li_3P remained on the surface and the amount of Li_2O had been substantially reduced. However, the amounts of

P, S, and O on the surface did not decrease compared to times after ~4 h of discharge (10 h total). The absence of substantial changes in the near-surface distribution of anion species within the SEI indicates that while some SEI components form irreversibly (e.g., Li_2S and Li_3PO_4), others remain redox active and can be partially delithiated (e.g., $2\text{Li}_2\text{O} \leftrightarrow \text{Li}_2\text{O}_2 + 2\text{Li}^+ + 2e^-$). It should be noted that even after no further changes in the SEI were observed, Li^+ must have continued to flow away from the surface and through the LPS since the sample did not exhibit extreme BE shifts due to surface charging. Therefore, we speculate that this current flow is associated with ongoing compositional and chemical-state changes deeper within the SEI, outside of the XPS information depth.

Overpotential measurements across individual SEI components. In addition to the compositional and chemical-state changes described above, BE shifts that result from sample biasing can be related to cell polarization (described in the experimental section). By combining data from these cell polarization measurements with the proposed layered configuration of the SEI, the overpotential (η) associated with each SEI phase can be estimated. This is calculated according to Eq. (6) by subtracting the cell polarization measured at one layer, ϕ_A , from the cell polarization observed for an adjacent layer, ϕ_B , (where layer B is on top of layer A) according to:

$$\eta_B = \phi_B - \phi_A \quad (6)$$

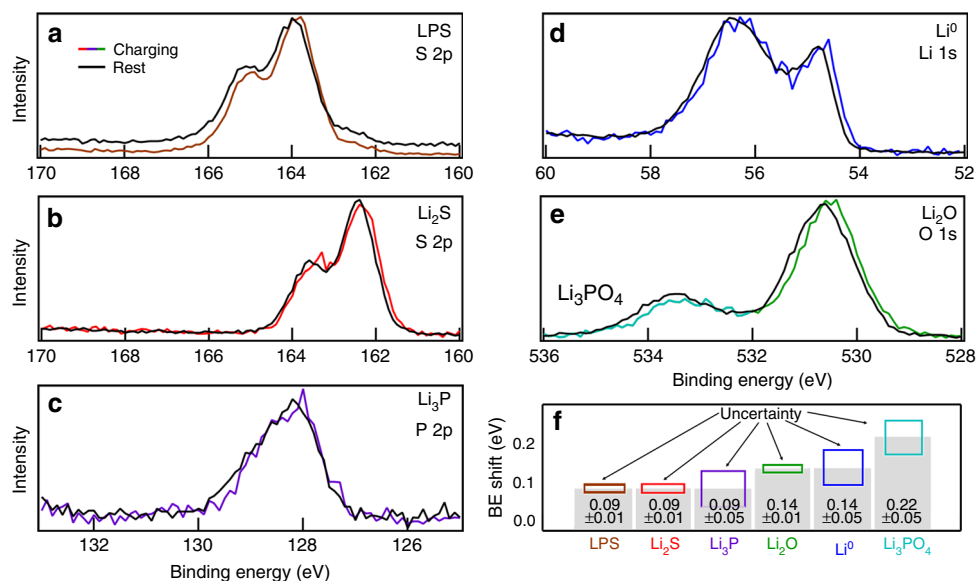


Fig. 5 opXPS determination of chemically resolved cell polarization/overpotential. XPS spectra showing chemically resolved BE shifts caused by applied current bias **a** before charging (LPS/Li_{foil}); and **b–e** after the charging half-cycle (Li/LPS/Li_{foil}). **f** Summarizes the current-bias-induced BE shifts extracted by peak fitting for each of the observed SEI phases. Individual overpotential losses associated with each phase can be determined from these shifts

Figure 5 shows XPS spectra from a sample with and without bias. These data were acquired at two different points during charging: before cycling (to measure the cell polarization of the initial LPS/Li_{foil} sample; Fig. 5a), and after the charge step (to measure cell polarization values across various SEI phases; Fig. 5b–e). From these spectra, BE shifts have been extracted via curve fitting and are summarized in Fig. 5f. This analysis shows that the net cell polarization across the LPS/Li_{foil} stack is 90 ± 10 mV (brown) at 0.17 mA/cm². Assuming that negligible impedance changes occur across this portion of the cell during the first charge step, overpotentials across individual SEI phases can be extracted. Comparing the voltage drops of both the Li₂S (red; 90 ± 10 mV) and the Li₃P (purple; 90 ± 50 mV) to that of the initial LPS/Li_{foil} reveals there are minimal contributions to cell overpotential from these SEI components. This indicates that these phases are either extremely thin or have good ionic conductivity. By comparison, the cell polarization associated with Li₂O (green) is much higher, 140 ± 10 mV, indicating that a significant overpotential drop occurs across this layer ($\eta_{\text{Li}_2\text{O}} = 50$ mV). For Li⁰, the same cell polarization as Li₂O is observed (blue; 140 ± 50 mV). This is consistent with Li⁰ plating on top of Li₂O, supporting the conclusion that oxygen observed in SEI species originates within the LPS SSE. Perhaps most importantly, the BE shift for the peaks associated with Li₃PO₄ show the largest total cell polarization 220 ± 50 mV. This suggests that Li₃PO₄ either forms on top of Li⁰ (unlikely), or that Li₃PO₄ exhibits such poor ionic conductivity that it acts as a blocking layer for Li⁺ transport. It is also noteworthy that local cell polarizations determined for Li₂O and Li⁰ (140 mV) closely match the value determined from a cycled symmetric coin cell (Supplementary Figure 7 and Supplementary Note 8), again indicating that the Li₂O layer is in the primary conductive pathway through the SEI. The overpotentials for each phase are summarized at the end of the next section.

Observation of SEI non-uniformity via in situ AES mapping.

To provide further evidence that Li₃PO₄ inhibits Li⁺ migration, the virtual electrode concept was used in combination with AES mapping to provide two-dimensional compositional information related to SEI formation. These results, shown in Fig. 6a–d (and Supplementary Movie 1), reveal the Li/LPS interface evolves

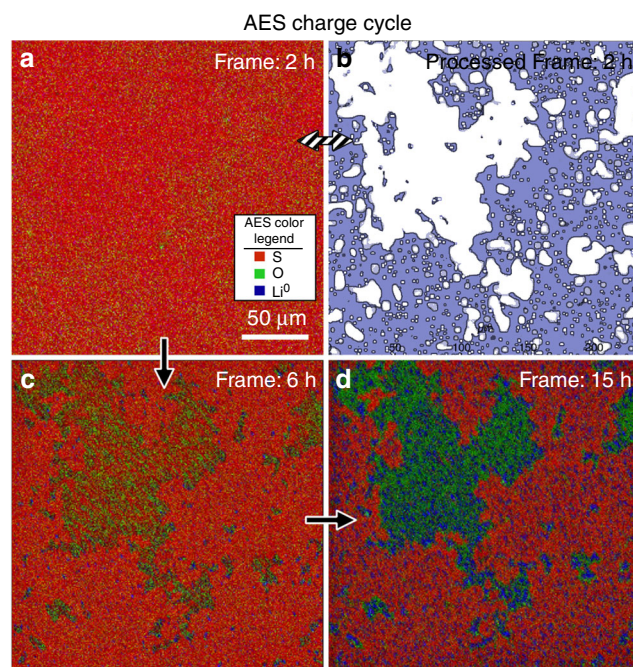


Fig. 6 In situ AES mapping of SEI heterogeneity. **a** shows an AES mapping image of the LPS SEI after 2 h of charging. **b** is a threshold-filter processed version of the same 2 h AES map, where white areas highlight little or no oxygen and blue areas represent higher oxygen concentrations, likely in the form of Li₃PO₄. **c**, **d** show the SEI after 6 and 15 h of charging, respectively

heterogeneously during the charge cycle. Careful inspection of Fig. 6a, representing the initial surface, reveals a heterogeneous distribution of oxygen. To highlight these subtle differences, the AES map in Fig. 6a was processed via a series of threshold filters, to show areas of higher vs. lower oxygen content (light blue vs. white, respectively) in Fig. 6b. Comparison of this processed image to the final AES map in Fig. 6d reveals a clear correlation between the SEI where Li₂O (green) and Li⁰ (blue) appear and the white areas in Fig. 6b. This observation is consistent with the

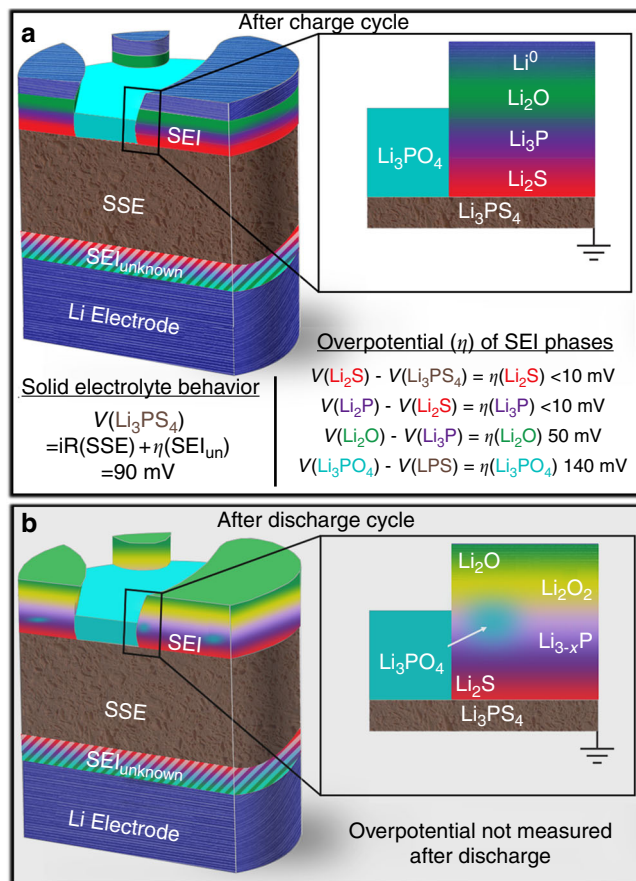


Fig. 7 Summary of SEI composition and morphology. A schematic is shown of the three-dimensional SEI structure after the charge (**a**) and discharge (**b**) step. A summary of the individual SEI overpotentials measured after the charging step is shown in (**a**)

hypothesis that oxygen contamination in LPS initially leads to Li_3PO_4 phase segregation, but also shows that oxygen content is initially non-uniform, and consequently Li_3PO_4 forms predominantly in oxygen-rich areas. Subsequently the Li_3PO_4 inhibits Li^+ transport in those areas, so that other SEI phases such as Li_2O form in complementary regions of the surface. This underscores the importance of controlling and understanding oxygen contamination and defects in LPS SSE's. Furthermore, the strong correlation between oxygen-rich areas of the SEI after charging (presumed to be Li_2O) and the appearance of Li^0 in Fig. 6 (and Supplementary Movie 1) provides further evidence for both the lateral and layered structure of the SEI.

To illustrate these conclusions graphically, Fig. 7 shows the proposed layered configuration after both charging and discharging. The phrase 'layered configuration' is not intended to convey a homogeneous structure across the entire interface, rather this term is used to describe the propensity of certain phases to form before others. Therefore, the schematic shown in Fig. 7 is the most likely structure of the SEI, although other configurations might exist at various points across the surface. Figure 7a also demonstrates the measured overpotentials across the individual SEI phases. Based on statistical analysis of local cell polarization values extracted from BE shifts (see Supplementary Figure 8, Supplementary Table 2 and Supplementary Notes 9 and 10), we can say with 99.8% certainty that Li_3P has a lower local cell polarization than Li_2O . Similarly, it can be stated with 99.6% certainty that Li_3PO_4 has a higher local cell polarization than Li^0 . The AES mapping results provide additional support for these

conclusions. Unfortunately, it is not possible to say with certainty whether Li_3PO_4 is above or adjacent to the Li_2S , Li_3P , or Li_2O phases. However, based on the combination of XPS and AES data it seems most likely that Li_3PO_4 initially forms directly on the LPS.

In this work we have developed and implemented the virtual electrode approach for cycling solid-electrolyte-based batteries while monitoring interface composition and properties with XPS or AES. These techniques enable operando and in situ measurements of SEI formation and electrochemical evolution during charge-discharge cycles. A primary finding from this study is that the evolution of phase compositions within the XPS information depth provides evidence for a layered configuration within the Li/LPS SEI. Specifically, the opXPS measurements reveal that Li_2S and Li-P phases form first, followed by Li-O phases, and finally Li^0 . OpXPS results also reveal that Li_3P and Li_2O are redox-active SEI phases, as indicated by concentration and chemical-state changes in the P 2p and O 1s core levels during discharge, consistent with transformation of these phases into Li-deficient compositions (i.e. $\text{Li}_2\text{O} \rightarrow \text{Li}_2\text{O}_2$ and $\text{Li}_3\text{P} \rightarrow \text{Li}_{3-x}\text{P}$). On the other hand, Li_2S is a stable SEI phase that does not evolve further during cycling, at least for the modest electrochemical conditions tested herein.

Our results also demonstrate that controlled surface charging using the virtual electrode approach enables overpotential losses across individual SEI phases to be directly quantified. Relative BE shifts in XPS spectra during biased and non-biased conditions allow for individual measurements of overpotentials, which can be assigned to the specific SEI phases. In the case of LPS solid electrolytes, results from this analysis show that the Li_2S and Li-P SEI components cause negligible polarization losses. On the other hand, the oxygen-containing components exhibit much larger polarization losses. OpXPS results coupled with in situ AES measurements reveal that Li_3PO_4 inhibits Li^+ migration. The Li_3PO_4 SEI phase forms initially through phase segregation of oxygen-contaminated LPS ($\text{Li}_3\text{PO}_x\text{S}_{4-x}$), and subsequently grows during discharge as a consequence of reactions between the delithiated Li_{3-x}P and Li_2O_2 phases. Of all SEI phases observed, Li_3PO_4 exhibits the largest overpotential, indicating that Li_3PO_4 hinders Li^+ migration and therefore is likely detrimental to battery performance, cycling stability and the uniform deposition of Li. To the best of our knowledge, these opXPS measurements represent the first chemically resolved overpotentials that have been measured across individual SEI phases.

Methods

LPS synthesis. The LPS electrolyte was prepared with a molar ratio of 77.5% Li_2S (99.9% Alfa-Aesar) and 22.5% P_2S_5 (99% Sigma-Aldrich). All material handling and electrode fabrication occurred in an argon environment. The material was planetary milled for 20 h in 500 mL stainless steel jars. LPS samples for opXPS experiments were fabricated by compressing 200 mg of powder at 75 MPa into 13-mm diameter pellets approximately 1.2 mm thick. This step was followed by application of a lithium metal-foil electrode on a single side at 15 MPa. Samples used during galvanostatic cycling and AC impedance testing consisted of one lithium electrode and one copper electrode pressed onto the LPS. In addition, one galvanostatic cycling data set shown in the SI was acquired from a symmetric-cell-configuration sample consisting of lithium metal electrodes on both sides of the LPS.

XPS analysis. OpXPS measurements were performed in a Physical Electronics 5600 photoelectron spectrometer under ultrahigh vacuum (UHV) conditions, with typical pressures below 5×10^{-10} Torr. The Al K α X-ray monochromator was operated with an anode power of 350 W. The sample surface normal was oriented at 45° to both the X-ray source and photoelectron spectrometer. The spectrometer pass energy was set to 29.35 eV. Binding-energy (BE) calibrations on sputter-cleaned metal foils followed values reported by Seah for Au 4f $_{7/2}$, Ag 3d $_{5/2}$, and Cu 2p $_{3/2}$ core-levels²⁷. Curve fitting and data processing was performed using Igor Pro with a custom program adapted from Schmid et al.²⁸.

Auger Electron Spectroscopy (AES) analysis. Real-time in situ AES measurements were performed using a Physical Electronics 670 system, under beam energy of 5 kV, with 20-nA beam current. Typical pressures ranged from 3×10^{-10} to 1×10^{-9} Torr.

Electrochemical operando XPS. Cohen and coworkers have previously used electron-gun induced, controlled surface charging to develop a non-destructive XPS depth-profiling method that was applied to studies of metal-organic coordinated multilayers²⁹. Subsequently, they used a similar approach to probe alkali metal deintercalation from inorganic fullerene-like materials³⁰. In the present work, we use a similar approach, where an applied current bias enables electrochemical cycling in a battery device structure by alternating between an electron source and a photon source. We refer to this method of supplying charge to the surface as the ‘virtual electrode’.

Charge half-cycle. In the ‘virtual-electrode’ approach developed in this work, a Physical Electronics model 04-090 electron flood gun was used to supply electrons to the surface during the charge half-cycle. In a typical XPS application this electron gun provides a flux of low-energy electrons that neutralize surface charge on insulating samples. Generally, higher electron kinetic energies are associated with overall higher beam currents. Experiments were performed to characterize the relative tradeoffs between electron kinetic energy vs. net electron flux at the surface. Based on these measurements, the kinetic energy was set to 11.5 eV (the maximum for this electron gun), which produced an incident current density, (measured with a Faraday cup, of ~ 0.17 mA/cm²). In virtual electrode opXPS experiments, the electron flux to the surface results in a net negative surface charge that induces an opposing Li⁺ ion migration flux. Equilibrium conditions are established as soon as the built-up negative surface charge provides a sufficient driving force for migration of Li⁺ through the LPS/Li_{foil} sample, such that the two fluxes exactly balance. The net negative charge at the surface manifests as XPS BE shifts, ΔBE , equivalent to the voltage bias across the sample: $V_{\text{bias}} = \Delta BE/q$, where q is the electronic charge. In the present study, the timescale for equilibration of electronic and ionic fluxes (as revealed by observed BE shifts) is short, on the order of <1 s. Importantly, the conclusion that net electronic and ionic current densities are equal in magnitude depends on the assumption that the LPS electronic conductivity is negligible. It should also be noted that the net electron current density at the electrode during opXPS could be lower than the incident electron flux, due to the effects of elastic and inelastic backscattering and surface charging effects. For example, the net flux into a metallic Li foil sample using the same flood gun settings produced a net electron current ~ 0.12 mA/cm², and we expect that the ratio of incident electron flux to net current density is material dependent to some extent. OpXPS control experiments on LPS samples (Supplementary Figure 9 and Supplementary Note 11) using a lower electron kinetic energy (2.2 eV) did not show any significant differences aside from slower rates of Li⁺ migration associated with lower incident electron fluxes.

Discharge half-cycle. To implement the virtual electrode concept during the discharge half-cycle, net positive surface charge was generated through photoemission using UV photons ($\lambda = 405$ nm, $h\nu = 3.06$ eV) supplied by an LED (Sensor Electronic Technology, Inc.). This positive surface charge provided the driving force for Li⁺ migration away from the free surface and into the LPS sample. The LED was mounted outside the UHV chamber and light was focused through a viewport onto the sample surface. Due to the very low kinetic energies of photoelectrons created by the UV LED, a negative voltage bias (~ 45 V) was applied between ground and the Li metal-foil electrode. This arrangement biases the entire LPS/Li_{foil} sample relative to ground potential in the vacuum chamber, which accelerates low-energy photoelectrons away from the sample surface, thereby maximizing ionic current density. With UV illumination and applied voltage bias, net photoemission currents ~ 150 nA were measured, indicating that a sufficient electrochemical driving force is applied by this method to drive Li⁺ ion migration away from the surface and through the LPS. Based on calibration measurements using an area-defining aperture, the estimated ionic current density during discharge is ~ 0.01 – 0.02 mA/cm².

Battery test structure and virtual electrode. Figure 8 shows a schematic of the opXPS experiment. The sample consists of LPS mechanically pressed onto a Li metal-foil electrode. Based on measurements by Wenzel et al.¹¹, when the Li metal foil is brought into contact with LPS an SEI forms creating a LPS/SEI/Li_{foil} structure. (In this paper we use LPS/Li_{foil} and LPS/SEI/Li_{foil} interchangeably, recognizing that an SEI always present at the rear LPS/Li_{foil} interface.) During measurements the Li_{foil} electrode is in electrical contact with the grounded or voltage-biased XPS sample stage. In this configuration, the bare LPS surface is initially exposed to only the X-ray analysis beam, in order to determine initial composition and chemical states. To drive Li⁺ through the LPS toward the surface (Fig. 8a), the virtual electrode is applied by exposing the analysis surface to a flux of low-energy electrons. After charging a symmetric cell configuration is formed (i.e., Li/SEI/LPS/SEI/Li_{foil}).

To investigate the dynamic nature of the SEI during cycling, the virtual electrode concept was also used to supply net positive charge to the exposed

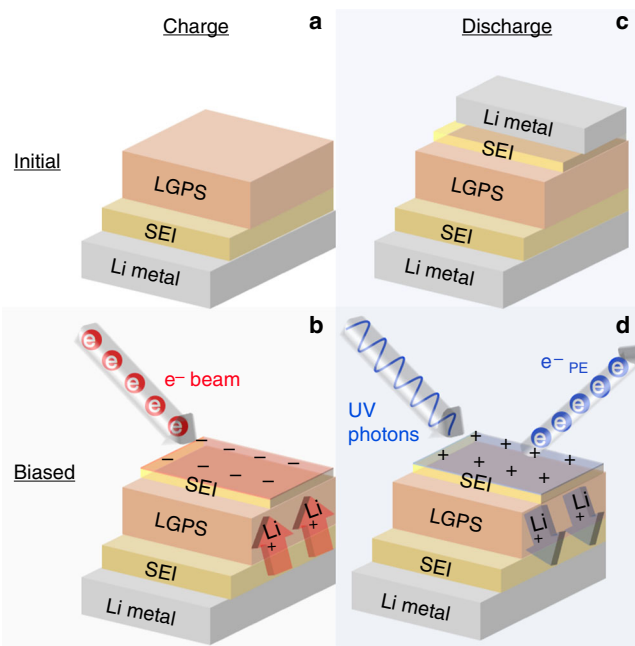


Fig. 8 Schematic of opXPS virtual electrode method. Schematic showing **a** the initial structure of the LPS/Li_{foil}, **b** the structure part of the way through the charge step, **c** the structure after the charge cycle is complete before right discharge and **d** the Li/LPS/Li_{foil} during the discharge step

analysis surface. This was accomplished by applying UV photons to the top surface (Fig. 8d), which eject valence electrons from phases with ionization potentials less than the photon energy. For the materials in the present study, this creates an equilibrium accumulation of excess Li⁺ ions near the surface, which in turn drives Li⁺ migration away from the analysis surface, through the LPS, back to the Li metal-foil electrode. This process is referred to as the discharge half cycle. After discharge the cell configuration is SEI/LPS/SEI/Li_{foil}.

It is important to note that the magnitude of the current density associated with the X-ray photoemission process used to generate XPS spectra is typically much smaller than those associated with either the electron flood gun or the UV LED light source, which together comprise the so-called virtual electrode. Nevertheless, in control experiments where high X-ray fluxes were supplied via a standard (non-monochromatic) Al anode X-ray source, we observed effects of Li⁺ migration caused by the photoelectron flux associated with the XPS measurement. Since X-ray photoemission extracts core-level electrons as well as valence electrons from all atomic species in a sample, it is possible that X-ray exposure can drive non-equilibrium electrochemical processes, which might more commonly be referred to as beam damage. Therefore, steps were taken to minimize net X-ray exposure during the opXPS measurements to levels well below beam damage thresholds determined in control experiments.

Operando XPS cell polarization/overpotential measurements. In this study local cell polarization, ϕ , refers to the net potential difference between the Li_{foil} electrode and an arbitrary point within the cell. Varying conventions exist for the definition of overpotential. To avoid confusion, herein local overpotential refers to the difference between local cell polarization and the sum of the equilibrium (thermodynamic) reduction/oxidation potentials combined with voltage losses associated with mass transport through the electrolyte. Therefore, this definition of local overpotential only includes voltage loss contributions associated with activation and transport limitations within the SEI. Information relating to both local cell polarizations and overpotentials can be extracted from opXPS measurements by comparing XPS core-level BE positions as a function of bias conditions. During charge and discharge half-cycles, current applied by the virtual electrode causes the cell to self-bias, creating a potential drop across the cell. In the case of an SEI with a layered structure, core levels associated with a particular SEI phase shift by the local cell polarization measured between ground potential and that phase. Therefore, for the initial LPS/SEI/Li_{foil} sample, the change in core-level BEs between the biased and unbiased condition is equal to the overpotential across the buried SEI added to the internal resistance voltage drop across the electrolyte. Similarly, by observing the biased vs. unbiased core-level BE shifts at the end of the charge half-cycle, it is possible to determine local overpotentials associated with individual phases within the SEI that has formed on the LPS surface.

Real-time in situ AES mapping measurements. The AES electron gun was used to charge the sample surface via the virtual electrode approach, and concurrently to

obtain sequential AES compositional images. The electron gun accelerating voltage and beam current were set to 5 kV and 20 nA, respectively. The ~25 nm diameter electron beam probe thereby resulted in quite large local instantaneous beam currents ($>1000 \text{ A cm}^{-2}$). On the other hand, rastering the beam over a $200 \mu\text{m} \times 250 \mu\text{m}$ ($5 \times 10^{-4} \text{ cm}^2$) area reduced the time-averaged current density to 0.04 mA cm^{-2} . Nevertheless, because charging with a rastered electron beam differs significantly from current flow in standard electrochemical cells, we describe the AES mapping measurements as 'real-time in situ' rather than 'operando'. Also, it is worth noting that in spite of the relatively high electron kinetic energies used in AES measurements, no significant beam-damage artifacts were observed.

Data availability. All relevant data are available from the corresponding authors upon request.

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References

- Crabtree, G. The energy-storage revolution. *Nat. Outlook Energy Storage* **526**, S92 (2015).
- Choi, J. W. & Aurbach, D. Promise and reality of post-lithium-ion batteries with high energy densities. *Nat. Rev. Mater.* **1**, 16013 (2016).
- Wood, K. N., Noked, M. & Dasgupta, N. P. Lithium metal anodes: toward an improved understanding of coupled morphological, electrochemical, and mechanical behavior. *ACS Energy Lett.* **2**, 664–672 (2017).
- Wood, K. N. et al. Dendrites and pits: untangling the complex behavior of lithium metal anodes through operando video microscopy. *ACS Cent. Sci.* **2**, 790–801 (2016).
- Chen, K.-H. et al. Dead lithium: mass transport effects on voltage, capacity, and failure of lithium metal anodes. *J. Mater. Chem. A* **5**, 11671–11681 (2017).
- Aurbach, D. Review of selected electrode-solution interactions which determine the performance of Li and Li ion batteries. *J. Power Sources* **89**, 206–218 (2000).
- Aurbach, D., Zinigrad, E., Teller, H. & Dan, P. Factors which limit the cycle life of rechargeable lithium (metal) batteries. *J. Electrochem. Soc.* **147**, 1274–1279 (2000).
- Fergus, J. W. Ceramic and polymeric solid electrolytes for lithium-ion batteries. *J. Power Sources* **195**, 4554–4569 (2010).
- Zhu, Y., He, X. & Mo, Y. Origin of outstanding stability in the lithium solid electrolyte materials: insights from thermodynamic analyses based on first-principles calculations. *ACS Appl. Mater. Interfaces* **7**, 23685–23693 (2015).
- Richards, W. D., Miara, L. J., Wang, Y., Kim, J. C. & Ceder, G. Interface stability in solid-state batteries. *Chem. Mater.* **28**, 266–273 (2016).
- Wenzel, S. et al. Interphase formation and degradation of charge transfer kinetics between a lithium metal anode and highly crystalline Li₇P₃S₁₁ solid electrolyte. *Solid State Ion.* **286**, 24–33 (2016).
- Wenzel, S., Leichtweiss, T., Krüger, D., Sann, J. & Janek, J. Interphase formation on lithium solid electrolytes - an in situ approach to study interfacial reactions by photoelectron spectroscopy. *Solid State Ion.* **278**, 98–105 (2015).
- Schwöbel, A., Hausbrand, R. & Jaegermann, W. Interface reactions between LiPON and lithium studied by in-situ X-ray photoemission. *Solid State Ion.* **273**, 51–54 (2015).
- Sicolo, S., Fingerle, M., Hausbrand, R. & Albe, K. Interfacial instability of amorphous LiPON against lithium: a combined Density Functional Theory and spectroscopic study. *J. Power Sources* **354**, 124–133 (2017).
- Dudney, N. J. Addition of a thin-film inorganic solid electrolyte (Lipon) as a protective film in lithium batteries with a liquid electrolyte. *J. Power Sources* **89**, 176–179 (2000).
- Mizuno, F., Hayashi, A., Tadanaga, K. & Tatsumisago, M. New, highly ion-conductive crystals precipitated from Li₂S-P₂S₅ glasses. *Adv. Mater.* **17**, 918–921 (2005).
- Mizuno, F., Hayashi, A., Tadanaga, K. & Tatsumisago, M. High lithium ion conducting glass-ceramics in the system Li₂S-P₂S₅. *Solid State Ion.* **177**, 2721–2725 (2006).
- Takada, K. Progress and prospective of solid-state lithium batteries. *Acta Mater.* **61**, 759–770 (2013).
- Takada, K. et al. Interfacial phenomena in solid-state lithium battery with sulfide solid electrolyte. *Solid State Ion.* **225**, 594–597 (2012).
- Ma, C. et al. Interfacial stability of Li metal–solid electrolyte elucidated via in situ electron microscopy. *Nano Lett.* **16**, 7030–7036 (2016).
- Gobet, M., Greenbaum, S., Sahu, G. & Liang, C. Structural evolution and Li dynamics in nanophase Li₃PS₄ by solid-state and pulsed-field gradient NMR. *Chem. Mater.* **26**, 3558–3564 (2014).
- Lepley, N. D. & Holzwarth, N. A. W. Modeling interfaces between solids: application to Li battery materials. *Phys. Rev. B - Condens. Matter Mater. Phys.* **92**, 1–15 (2015).
- Wang, X., Xiao, R., Li, H. & Chen, L. Oxygen-driven transition from two-dimensional to three-dimensional transport behaviour in beta-Li₃PS₄ electrolyte. *Phys. Chem. Chem. Phys.* **18**, 21269–21277 (2016).
- Tao, Y. et al. Lithium superionic conducting oxysulfide solid electrolyte with excellent stability against lithium metal for all-solid-state cells. *J. Electrochem. Soc.* **163**, A96–A101 (2016).
- Zavadil, K. R. & Armstrong, N. R. Surface chemistries of lithium: detailed characterization of the reactions of H₂S, SO₂, and SO₂Cl₂ using XPS and EELS. *Surf. Sci.* **230**, 61–73 (1990).
- Wang, K., Ross, P. N. Jr., Kong, F. & McLarnon, F. The reaction of clean Li surfaces with small molecules in ultrahigh vacuum. I. Dioxigen. *J. Electrochem. Soc.* **143**, 422–428 (1996).
- Seah, M. P., Gilmore, I. S. & Beamson, G. XPS: binding energy calibration of electron spectrometers 5--Re-evaluation of the reference energies. *Surf. Interface Anal.* **26**, 642–649 (1998).
- Schmid, M., Steinrück, H. P. & Gottfried, J. M. A new asymmetric Pseudo-Voigt function for more efficient fitting of XPS lines. *Surf. Interface Anal.* **46**, 505–511 (2014).
- Doron-Mor, I. et al. Controlled surface charging as a depth-profiling probe for mesoscopic layers. *Nature* **406**, 382–385 (2000).
- Feldman, Y., Zak, A., Tenne, R. & Cohen, H. Evidences for dry deintercalation in layered compounds upon controlled surface charging in x-ray photoelectron spectroscopy. *J. Vac. Sci. Technol. A* **21**, 1752–1757 (2003).

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Author contributions

G.T. supervised the project and conceived the virtual electrode approach for XPS and AES measurements. G.T., K.X.S., C.B., and S.S. conceived and planned opXPS measurements on solid electrolyte materials. K.N.W. and G.T. further developed the opXPS technique for battery analysis, performed the experiments and analyzed the data. S.E.H. and S.H.L. fabricated the samples and performed the ex situ electrochemical measurements. K.N.W. and G.T. wrote the initial draft of the manuscript. All other co-authors discussed the results and contributed to the final manuscript preparation.

Additional information

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