

Fig. 4. Ca and Si X-ray intensities for the grain in the left of Fig. 1, showing anhydrous 3CaO, SiO_2 , $\text{Ca}(\text{OH})_2$ and engulfed silicate particles.

This work was supported by the US National Science Foundation. We thank Dr E. W. White for assistance and advice and Mr W. Zeigler for instrumental help.

> MICHAEL W. GRUTZECK DELLA M. ROY

Department of Geochemistry and Mineralogy and Materials Research Laboratory, Pennsylvania State University, University Park, Pennsylvania 16802.

Received February 27; revised May 23, 1969.

- ¹ Roy, D. M., and Grutzeck, M. W., Fifth Intern. Symp. Chemistry of Cement, Tokyo (October 1968).
- White, E. W., Denny, P. J., and Irving, S., in *The Electron Microprobe* (edit. by McKinley, Heinrich, and Wittry), 791 (John Wiley and Sons, 1966).
- Kell, K., Fortschritte der Mineralogie, 44, 4 (1967).
 White, E. W., and Rustom, R., Proc. Symp. IITRI, 91 (1968).
- ⁵ Grutzeck, M. W., thesis, The Pennsylvania State University (1968).
- ⁶ Kantro, D. L., Weise, C. H., and Brunauer, S., in Symp. on Structure of Portland Cement Paste and Concrete, Special Report 90, Highway Res. Board, Washington, DC (1966).
- ⁷ Kondo, R., in Fifth Intern. Symp. on the Chemistry of Cement, Tokyo (October 1968).

Infrared Absorption Spectra of α and **β-Calcium Sulphate Hemihydrates**

In their recent communication concerning the application of infrared spectroscopy to the calcium sulphatewater system, Bensted and Prakash¹ stated that they obtained no spectral differences between the a and B forms of the hemihydrate. I would like to report that I have found differences in the spectral range 1,200 to 1,000 cm⁻¹ associated with the vibrations of the tetrahedral SO₄ ion as shown in Table 1.

> Table 1. INFRARED ABSORPTION (CM-1) a-CaSO4. H2O β-CaSO₄ . ½ H₂O 1,157 1,119 1,098 1,155

The samples of hemihydrate, from BPB Industries (Research and Development) Ltd, were characterized by differential thermal analysis2,3 and their infrared absorption spectra were recorded on Hilger and Watts Infrascan and Unicam SP 100 spectrophotometers using the potassium bromide disk technique4. In the case of the Infrascan the wavenumber scale was calibrated by means of a polystyrene film (supplied by Perkin Elmer Ltd).

In addition to minor positional differences, the absorption maxima were much sharper in the 3-hemihydrate than in the α form. These results indicate that the same structure is present in both forms of the hemihydrate but that the lattice of the β-form is less strained than in the a form⁵. This result is in accord with the conclusions drawn from differential thermal analysis^{2,3}.

D. A. BARTRAM

Brixton School of Building, London SW4.

Received February 18; revised May 30, 1969.

- ¹ Bensted, J., and Prakash, S., Nature, 219, 60 (1968). Holdridge, D. A., Trans. Brit. Ceram. Soc., 64, 211 (1965).
- Holdridge, D. A., and Walker, E. G., Trans. Brit. Ceram. Soc., 66, 485 (1967).
- ASTM E 168 64T General Techniques of Infrared Quantitative Analysis,
- Duyckaerts, G., The Analyst, 84, 201 (1959).

Morphology of ABA Block Polymers

MICROPHASE separation is believed to occur in many block copolymers. In interpreting the mechanical properties of block copolymers of the type $A_x B_y A_x$, where A_x and B_y are blocks having x repeat units of type A and y repeat units of type B respectively, it has been suggested that when y is much greater than x the A_x blocks aggregate to form spherical domains in a matrix composed of B_y blocks¹. Partial support for this model has recently been obtained from electron microscopy studies carried out on copolymers in which the outer blocks