

These determinations include the most classical labours on record, and the general agreement with the calculated numbers is surprising, and the more conspicuous in the cases in which the efforts of the experimenters to exclude error have been pushed to the utmost limits, as in Stas's syntheses and in Prof. Crookes's synthesis of thallium nitrate. Notwithstanding the difficulty in this case, because the element is the heaviest of all so far discovered, one experiment has yielded the identical calculated number, and the mean of all deviates from it only by 0.00131. Moreover the same weights recur in similar compounds; all nitrates, for instance, have a lower value than the corresponding chlorides and sulphates, and the value is the lower the greater the composition, as in the alums. The evidence is such that no doubt seems to be admissible as to the reality of a variation of the atomic weights. This conclusion is independent of any value of the atomic weights; for the discrepancies exhibited in the results of Prof. Clarke's recalculations from the same experimental data above quoted are inevitable if the variation of the atomic weights is not taken into account. In $c$ units Ag is 108.09679 if $\mathrm{H}=\mathrm{I}$, calculated from the weights of column $t$; Cl in the gaseous state is $=35^{\circ} 66$; the calculated weights correspond therefore, within the limits of experimental errors, to the atomic, but the weights are those of different states.

The difference between the weights of the gaseous and the other states is very considerable ; the weight of 3 molecules of $\mathrm{H}_{4} \mathrm{~N}_{3} \mathrm{I}$. HgI, for instance, is $=378$ in the state of gas, $354^{\circ} 734$ in $t$ units, 352.847 in units $=c$; the discrepancies are so great that they exceed by far the limits of possible errors, and as from the comparisons made it appears certain that the different values are realities, the only explanation is that the atomic weights vary. If in new experiments, in which the possibility of variation is kept in view, all discrepancies which actually exist should disappear, variation will be established beyond all doubt. It will then be in order to inquire into its cause. How the weights of the table have been obtained is, for the present, unessential; it is only necessary to add that column $v$ contains Prof. Clarke's recalculated weights, and column $u$ the same values calculated from the weights of column $t$, column $x$ giving the number of atoms represented in each instance. Column $w$ shows the corresponding weights of the gaseous state. These columns have been added for the sake of comparison.

|  |  |  | $s$ | $t$ | $u$ | $v$ | 2v | $x$ |
| :---: | :---: | :---: | :---: | :---: | :---: | :---: | :---: | :---: |
| Li |  | $\ldots$ | 22 | 2.36559 | 7'412 | $7^{\prime} \mathrm{O} 235$ | 7.333 | 3 |
| Ca | ... | $\ldots$ | 58 | $6 \cdot 23656$ | $39^{\prime} 0824$ | 40.082 | $38 \cdot 666$ | 6 |
| Na |  | $\ldots$ | 70 | 7.52688 | 23.5842 | $23^{\circ} \mathrm{O} 51$ | $23 \cdot 333$ | 3 |
| K |  | $\ldots$ | 118 | 12.68817 | $39^{\circ} 7564$ | $39^{\circ} 109$ | 39.333 | 3 |
| Rb |  |  | 256 | 27.5269 | $86 \cdot 2424$ | 85.529 | $85 \cdot 333$ | 3 |
| Mg | ... | $\ldots$ | 36 | $3 \cdot 8537$ | 24.15 | $24^{\circ} \mathrm{OI} 4$ | 24 | 6 |
| Sr |  | $\ldots$ | 132 | 14.1303 | $88 \cdot 5498$ | 87.575 | 88 | 6 |
| Ba | $\cdots$ | $\ldots$ | 206 | 20.05183 | $138 \cdot 1915$ | $137^{\circ} 007$ | $137 \cdot 333$ | 6 |
| Pb |  | ... | 306 | $32 \cdot 7566$ | $205 \cdot 2748$ | 206.946 | 204 | 6 |
| Ag | $\ldots$ | $\ldots$ | 324 | 34.683467 | $108 \cdot 6748$ | 107.923 | 108 | 3 |
| Cs | $\ldots$ | $\ldots$ | 398 | $42 \cdot 605$ | $133 * 496$ | 132.918 | 132.666 | 3 |
| H |  | $\ldots$ | 3 | 0.31915 | I | I'0023 | I | 3 |
| N |  | $\ldots$ | 14 | I 48936 | 14 | 14.029 | 14 | 9 |
| O |  |  | 24 | 2.55319 | 16 | 16 | 16 | 6 |
| F |  | $\ldots$ | 58 | 6.04166 | 18.93 | $19^{\circ} \mathrm{O} 27$ | 19.333 | 3 |
| Cl | ... |  | 107 | 11'14583 | $34 * 9236$ | 35.451 | 35.666 | 3 |
| Br | $\ldots$ | $\ldots$ | 243 | $25 \cdot 3125$ | $79 \cdot 3125$ | 79.951 | 81 | 3 |
| I | ... | $\ldots$ | 387 | $40 \cdot 3125$ | 126.313 | 126.848 | 129 | 3 |
| B | ... | $\ldots$ | 1 I | I'14583 | 10.771 | 10.966 | r 1 | 9 |
| G | ... | $\ldots$ | 14 | I 45833 | $9^{\circ} 072$ | 9*106 | 9*333 | 6 |
| C | $\ldots$ | ... | 18 | I.875 | 11.75 | 12.0011 | 12 | 6 |
| Si |  |  | 22 | 2.29166 | $28 \cdot 722$ | $28 \cdot 26$ | 29.333 | 12 |
| Al | ... |  | 28 | 2.9166 | $27^{\circ} 416$ | 27.075 | 28 | 9 |
| P | $\ldots$ | ... | 32 | 3.3333 | 31.33 | 3I.029 | 32 | 9 |
| Ti |  |  | 42 | 4.375 | 54.833 | 49.961 | 56 | 12 |
| La |  |  | 44 | 4.5833 | 143.61 | $138 \cdot 844$ | $146 \cdot 666$ | 30 |
| S | .. |  | 48 | 5 | 31.33 | 32.058 | 32 | 6 |
| Di | ... | $\ldots$ | 50 | $5 \cdot 20833$ | $146 \cdot 875$ | 144.906 | 150 | 27 |
| Yt |  |  | 60 | $6 \cdot 25$ | $88 \cdot 125$ | 90.023 | 90 | $13^{\prime} 5$ |
| Yb |  |  | 62 | 6.45833 | 182.125 | 173.158 | 186 | 27 |
| Ce |  |  | 64 | 6.6666 | 139.26 | $140 \cdot 747$ | 142.222 | 20 |
| Sc | $\ldots$ | $\cdots$ | 66 | $6 \cdot 875$ | $43^{\circ} 0833$ | 44.081 | 44 | 6 |
| Zr | $\cdots$ | $\ldots$ | 68 | 7.0833 | $88 \cdot 7777$ | 89.573 | $90 \cdot 666$ | 12 |
| Ga | $\cdots$ |  | 72 | 7.5 | $70^{\circ} 5$ | 68.963 | 72 | 9 |
| As |  |  | 76 | 7.9166 | 74.417 | $75^{\circ} \mathrm{O9}$ | 76 | 9 |
| V | $\cdots$ |  | 78 | $8 \cdot 125$ | 50.9166 | 51.373 | 52 | 6 |
| Cr | ... | - | 80 | $8 \cdot 3333$ | $52 \cdot 222$ | 52.129 | 53.333 | 6 |
| Mn |  | $\ldots$ | 84 | $8 \cdot 75$ | 54.833 | 54.029 | 56 | 6 |
| Fe | $\ldots$ | $\ldots$ | 86 | 8.9583 | $56 \cdot 139$ | $56 \cdot 042$ | $57 \cdot 333$ | 6 |
| Ni |  | $\ldots$ | 90 | 9.375 | $58 \cdot 75$ | 58.062 | 60 | 6 |
| Co |  |  | 91 | $9 * 4792$ | 59.403 | $59^{\circ} \mathrm{O2} 3$ | $60 \cdot 666$ | 6 |
| Sn |  |  | 92 | 9.5833 | 120.11 | I 17'968 | 122.666 | 12 |
| Cu |  |  | 96 | 10 | 62.666 | $63 \cdot 318$ | 64 | 6 |
| Nb |  |  | 98 | 10'20833 | 95.95833 | 94.027 | 98 | 9 |
| Zn | ... |  | 100 | 10.4166 | $65 \cdot 278$ | 65.054 | $66 \cdot 666$ | 6 |
| Ta | ... |  | 106 | II ${ }^{\prime} 04166$ | 184.5186 | 182.562 | 188.444 | 16 |
| Se |  | $\cdots$ | 120 | 12.5 | 78.333 | 78.978 | 80 | 6 |
| Sb | ... |  | 126 | 13.125 | 123.375 | 120.231 | 126 | 9 |
| W |  |  | 142 | 14.79166 | 185.3888 | 184.032 | 189.333 | 12 |
| Mo | $\ldots$ |  | 150 | 15.625 | 97.9166 | 95.747 | 100 | 6 |
| Cd |  |  | 170 | $17 \cdot 7083$ | I 10.972 | I 12.092 | 113.333 | 6 |
| In |  |  | 176 | $18 \cdot 3333$ | 114.888 | 113.659 | 117.333 | 6 |
| Th |  | . | 178 | 18.54166 | $232 \cdot 389$ | 233.951 | $237 \cdot 333$ | 12 |
| U |  | .. | 184 | 19'1666 | $240 \cdot 222$ | $239^{\circ} \mathrm{O} 3$ | $245 \cdot 333$ | 12 |
| Te | ... | ... | 196 | 20*4166 | 127.945 | 128.254 | 130.666 | 6 |
| Au | $\ldots$ |  | 204 | 21'25 | 199*75 | 196.606 | 204 | 9 |
| Bi |  | $\cdots$ | 216 | 22.5 | 211.5 | 208.001 | 216 | 9 |
| Ir | $\ldots$ | .. | 300 | 31.25 | 195.833 | 193.094 | 200 | 6 |
| Pt | ... | .. | 304 | $31 \cdot 6666$ | 198.444 | 194.867 | $202 \cdot 666$ | 6 |
| Hg | * | ... | 306 | 31.875 | 199.75 | 2001771 | 204 | 6 |
| Os | ... |  | 308 | 32.0833 | 201.056 | 198.951 | $205 \cdot 333$ | 6 |
| Ru |  |  | 318 | $33 \cdot 125$ | $103 \cdot 7916$ | 104.457 | 106 | 3 |
| Rh |  |  | 320 | 33.3333 | 104.444 | 104.285 | $106 \cdot 666$ | 3 |
| Pd |  |  | 326 | 33'95833 | 106.403 | 105.981 | 108•666 | 3 |
| Tl |  |  | 618 | 64.375 | 201 708 | 204'183 | 206 | 3 |

San Francisco, California, July 24
E. Vogel

## UNIVERSITY AND EDUCATIONAL INTELLIGENCE

Cambridge.-The following gentlemen were on Monday, November 3, elected to Fellowships at St. John's College :C. M. Stewart, M.A., First Class in Natural Sciences Tripos of 1879, author of several papers on chemical subjects, and Master
at the Newcastle School, Staffordshire ; J. Brill, B.A., Fourth Wranger in 1882, Assistant Professor of Mathematics in University College, Aberystwith ; W. F. R. Weldon, B.A., First Class in the Natural Sciences Tripos of 1881, author of a number of papers in Zoology and Comparative Anatomy, formerly Demonstrator to the Professor of Zoology and in the Morphological Laboratory ; A. R. Johnson, B.A., Sixth Wrangler and First Division in the Mathematical Tripos of 1882-83 (new regulations), author of papers in the Messenger of Mathematics, \&c.; G. F. Stout, B.A., First Class in the Chemical Tripos of 1881-82 (new regulations), and First Class (with distinction in Metaphysics) in the Moral Sciences Tripos of 1883; G. B. Mathews, B.A., Scnior Wrangler in 1884, Profcssor of Mathematics in the University College of North Wales, Bangor. It is worth noting that Pure and Applied Mathematics, Chemistry, and Biology have been markedly recognised by this election.
Dr. Donald MacAlister has been appointed University Lecturer in Medicine, and Dr. Bushell Annington University Lecturer in Medical Jurisprudence.
Mr. Walter Heape has been approved by the Board for Biology and Geology as Demonstrator in Animal Morphology, on the nomination of the Lecturer in that subject, Mr. Sedgwick.

Prof. Sidgwick, Prof. Adamson (Owens College), and Messrs. James Ward and J. S. Nicholson are appointed Examiners for the Moral Sciences Tripos.

Mr. A. R. Forsyth of Trinity College is appointed Fxaminer in the Mathematical Tripos ('Third Part) in January next, in the place of the late Mr. R. C. Rowe.

In reference to our note a fortnight ago (vol. xxx. p. 649), we should state that, at Trinity College, Major Scholarships of the value of $80 \%$. a year, which may be raised to $100 \%$. subsequently, are open for competition in Natural Sciences as well as in Classics and Mathematics to persons not yet in residence, with the usual restriction as to age.
Sheffield.-Another step has been taken in the formation of the new Engineering School at Firth College, Sheffield, in the appointment of Mr. W. H. Greenwood to be Professor of Metallurgy and Mechanical Engineering, and Mr. Ripper to be Assistant Professor of Engineering. It may be in the memory of our readers that the City and Guilds of London Institute made a grart about eighteen months ago of 300 l . a year to the Firth College in aid of the establishment of a Chair of Engincering. Since then additional subscriptions have been promised for five years to the amount of $550 l$., together with a capital sum of over 10,000l. A site for laboratories and shops has been obtained, and these will be proceeded with as soon as possible. It is hoped that the special advantages of Sheffield will make it the central school of metallurgy, especially for iron and steel, in the kingdom, and the Committee intend to spare no efforts in rendering it a complete and effective one.

## SCIENTIFIC SERIALS

The American Yournal of Science, Scptember.-On the amount of the atmospheric absorption, by S. P. Langiey. From numerous observations taken at sea-level or at an altitude of nearly 15,000 feet, the author is led to infer that the mean absorption of light as well as of heat by our atmosphere is probably at least double the usual estimate of about 20 per cent. He also believes that fine dust particles, both near the surface and at a great altitude, play a more important part in this absorption, both general and selective, than has been hitherto supposed.--A study of tornadoes, by IIenry A. Hazen. In this paper the author examines some of the ordinary theories that are advanced for explaining the origin and development of these destructive phenomena. After showing some of the seeming difficulties involved in these theorics, he proceeds to point out a few of the characteristics of the outbursts, with a view to opening up fresh lines of investigation, upon which a further advance may be made towards a true knowledge of the forces underlying them. He is inclined to think that J. Allan Broun's theory, attributing tornadoes to the direct influence of the sun's electricity upon the moisture of the air, or possibly to the indirect effect from the sun's heat, is more satisfactory than the numcrous theories of friction, evaporation, condensation, sudden changes of temperature, and the like.-On the absorption of radiant heat by carbon dioxide, by J. E. Keeler. The author considers it probable that to the action of $\mathrm{CO}_{2}$ in the atmosphere is due one or more of the
great gaps in the invisible part of the solar spectrum which the discoveries of Prof. Langley show to be much more extensive than had hitherto been supposed. IIe further regards it as certain that some other agent than this gas contributes essentially to the total absorptive power of the atmosphere, so that a method of analysis based on this power, in which the effect of the second agent is neglected, cannot lead to correct results. Note on the Triassic insects from Fairplay, Colorado, by Samuel N. Scudder. These fossil remains present an assemblage of forms altogether different from anything hitherto found in the Palæozoic series on the one hand, or in the Jurassic beds on the other. They seem to show a commingling of strict Jurassic forms with a larger proportion of types which may be called Upper Carboniferous or Permian, with a distinct Jurassic leaning. Hence the probability that the beds in which they.occur belong to the Triassic or intermediate formation.-On the flexibility of Itacolumite, by Orville A. Derby. From observations made on this extensive series of quartzose rocks occurring in the gold and diamond resions of Minas Gerats, Brazil, the author infers that the peculiar property of flexibility attributed to them is not an original characteristic, but only a surface character, a phase of weathering or decay brought about by percolating waters.On the age of the glazed and contorted slaty rocks in the vicinity of Schodack Landing, Rensselaer County, New York, by S. W. Ford.-On the relations of the mineral belts of the Pacific slope to the great upheavals, by Geo. F. Becker. The views of H. P. Blake and Clarence King regarding the parallelism of the series of mineral belts on the Pacific slope to the great mountain ranges, and attributing the deposits themselves to the solfateric action accompanying the ejection of igneous rocks, have since been mainly confirmed. But, independently of any theory, a conclusion of economical importance cvidently follows from the fresh facts recently brought to light. A great majority of all the rich ores west of the Wahsatch Range occur in belts following the western elges of distinct geological areas-the Cretaceous in Utah, the Palæozoic and Carboniferous in Nevada and Arizona, the Jura-Trias in East California, \&c. Hence analogy points to the neighbourhood of the still unexplored portions of these contacts as the most promising for future discoveries of the precious metals.-Notice of the remarkable marine fauna occupying the outer banks off the southern coast of New England, No. 9, by A. E. Verrill.- Brief contributions to zoology from the Museum of Yale College, No. 1v.--Work of the steamer Allatross in 1883.-Geology of the Blue Ridge, near Balcony Falls, Virginia, by John L.. Campbell.

October.-On the duration of colour-impressions upon the retina, by Edward 1. Nichols. Taking up the subject where it was left fifty years ago by Plateau's researches, the author concludes, from a protracted series of experiments: (I) that the study of the duration of colour-impressions produced by different portions of the spectrum tends to confirm Plateau's results; (2) that the persistence of the image is a function of the wave-length producing it, being greatest at the ends of the spectrum, and least in the ycllow ; (3) that it decreases with the intensity of the ray producing it ; (4) that it is not the same for all cycs; (5) that the duration is in inverse order to the luminosity of the colours producing it ; (6) that each wave-length of the visible spectrum produces three primary impressions, red, green, and violet, of which the green is the most evanescent, violet the most persistent ; (7) that the duration of the retinal image depends upon the length of time during which the eye has been exposed, decreasing as the exposure increases.-Description of a fulgurite from Mount Thielson, Oregon (one illustration), by J. S. Diller.-On the paramorphosis of pyroxene to hornblende in rocks (two illustrations), by Geo. H. Williams.-On the southward ending of a great synclinal in the Taconic Range (with a map and several illustrations), by James D. Dana. The section of the Taconic Range here dealt with extends about 150 miles along the western border of New England, mainly between Middlebury, in Central Vermont, and Salisbury, in North-Western Connecticut. The conclusions arrived at regarding the synclinal character of the system and the Lower Silurian age of the rocks agree with those of Sir William Logan, except that he made the limestone to precede instead of to include the Trenton group.-On supposed glaciation in Pennsylvania, south of the terminal moraine (with a map), by Prof. H. Carville Lewis. The author considers that all the existing surface phenomena may be explained by the action of running waters and other causes independent of glacia-tion.-History and chemical analysis of a mass of meteoric iron

