

SCIENTIFIC REPORTS

OPEN

Promoting the Adsorption of Metal Ions on Kaolinite by Defect Sites: A Molecular Dynamics Study

Xiong Li, Hang Li & Gang Yang

Received: 16 May 2015

Accepted: 25 August 2015

Published: 25 September 2015

Defect sites exist abundantly in minerals and play a crucial role for a variety of important processes. Here molecular dynamics simulations are used to comprehensively investigate the adsorption behaviors, stabilities and mechanisms of metal ions on defective minerals, considering different ionic concentrations, defect sizes and contents. Outer-sphere adsorbed Pb^{2+} ions predominate for all models (regular and defective), while inner-sphere Na^+ ions, which exist sporadically only at concentrated solutions for regular models, govern the adsorption for all defective models. Adsorption quantities and stabilities of metal ions on kaolinite are fundamentally promoted by defect sites, thus explaining the experimental observations. Defect sites improve the stabilities of both inner- and outer-sphere adsorption, and (quasi) inner-sphere Pb^{2+} ions emerge only at defect sites that reinforce the interactions. Adsorption configurations are greatly altered by defect sites but respond weakly by changing defect sizes or contents. Both adsorption quantities and stabilities are enhanced by increasing defect sizes or contents, while ionic concentrations mainly affect adsorption quantities. We also find that adsorption of metal ions and anions can be promoted by each other and proceeds in a collaborative mechanism. Results thus obtained are beneficial to comprehend related processes for all types of minerals.

Clays are naturally occurring aluminosilicate minerals that have been used in a wide range of catalytic¹, environmental² and pharmaceutical³ processes. Clay minerals have particular ion-exchange capacity and are known as reservoirs for metal ions, while the uptake of heavy metal ions poses serious environmental problems. Currently, the adsorption mechanisms of metal ions on clay minerals remain largely elusive. Two metal ions may correspond to different adsorption behaviors, and the adsorption performances are closely associated with the surface properties of clay minerals that can be greatly altered by the extent of structural defects^{2,4–6}.

Two adsorption modes (inner- and outer-sphere) were proposed for the interaction of metal ions with clay minerals^{7–27}. Inner-sphere metal ions form stable complexes directly with clay minerals, while outer-sphere ones are separated from clay minerals by an intermediate water molecule. The combination of AFM experiments and MD simulations demonstrated that hydrated Na^+ ions are adsorbed close to calcite as Rb^+ ions¹³. With use of MD simulations, Greathouse and Cygan²⁴ found that $[\text{UO}_2(\text{H}_2\text{O})_5]^{2+}$ is the principal species for UO_2^{2+} adsorption on beidellite and montmorillonite, and Yang *et al.*²⁵ further pointed out that $[\text{UO}_2(\text{H}_2\text{O})_5]^{2+}$ can interact with montmorillonite via the U cationic center and O atom of hydration shell corresponding to the inner- and outer-sphere modes, respectively. Sakuma *et al.*²⁶ reported that all Na^+ , K^+ and Cs^+ ions form stable inner-sphere complexes with muscovite, while Vasconcelos *et al.*²⁷ insisted that relatively weak outer-sphere complexes are the main and sole products respectively for Na^+ and Pb^{2+} adsorption on kaolinite. This seems to contradict with the experimental observations where strong adsorption on kaolinite was documented^{2,28–30}.

College of Resources and Environment & Chongqing Key Laboratory of Soil Multi-scale Interfacial Process, Southwest University, Chongqing 400715, China. Correspondence and requests for materials should be addressed to H.L. (email: lihangswu@163.com) or G.Y. (email: theobiochem@sina.com)

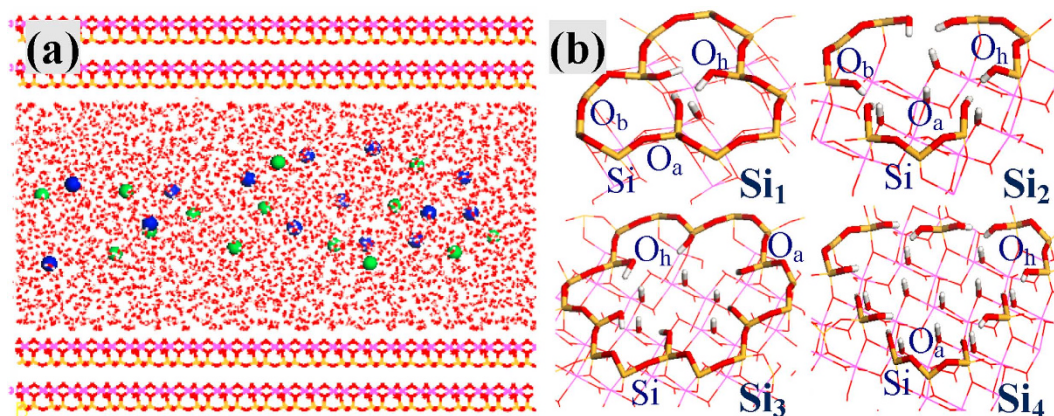


Figure 1. (a) Initial configuration of regular kaolinite in 0.16 mol/L NaCl solutions and (b) local structures of the various defect sites (Si₁, Si₂, Si₃ and Si₄) created at the tetrahedral SiO₄ surface. The tetrahedral SiO₄ and octahedral AlO₆ surfaces of kaolinite that have interfacial interactions with salt solutions are situated at the top and bottom, respectively. O_h is from the hydroxyl of silanol, and O_a (bridging O) is also bonded to the silanol Si atom while O_b (bridging O) has no relation with defect sites. Na⁺ and Cl⁻ ions are represented as blue and green balls, respectively.

Defect sites, which have been considered by not a few researchers as adsorption and catalytic centers in aluminosilicates and analogues^{5,6,31–36}, were assumed to reconcile such contradictions. Defects in silicates and zeolites (analogues of kaolinite and montmorillonite) have been well documented experimentally, observing different sizes of defects with silanol nests^{37–43}. Silanol defects were also detected in clay minerals^{44–47}, and it has recently been found that a significant fraction of carbon resources reside at the defects of minerals in the Earth's mantle⁵. Presently, MD simulations were used to simulate the adsorption of metal ions on regular and defective kaolinite (Fig. 1) with aim to demonstrate the roles of defect sites played during such adsorption processes. The adsorption configurations, quantities and stabilities are all fundamentally altered by defect sites, and the computational results can explain the experimental observations^{2,28–30}. Two metal ions (Na⁺ and Pb²⁺) were chosen that respond distinctly to defect sites. Na⁺ is a common background metal ion while the environmental pollution of Pb²⁺ has emerged as a global concern. The choice of a wide range of ionic concentrations allows us to discern the concentration-dependent adsorption processes and potentially different adsorption behaviors before and after saturation. In order to comprehensively understand defect sites, effects of defect sizes and defect contents were also considered. In the end, the adsorption mechanisms of metal ions and anions on kaolinite were explored.

Results

Regular kaolinite. Discussions of Na⁺ and Pb²⁺ adsorption on regular kaolinite are limited in the report of Vasconcelos *et al.*²⁷ and will be elaborated here, mainly to facilitate the understanding of the adsorption processes on defect sites. As indicated in Fig. 2, S13 and S14, metal ions (Na⁺/Pb²⁺) and anions (Cl⁻) tend to approach the slightly negatively charged tetrahedral SiO₄ surfaces and positively charged octahedral AlO₆ surfaces, respectively; meanwhile, very few Na⁺ ions are driven by Cl⁻ ions towards the octahedral AlO₆ surfaces, which are in line with the previous results²⁷. Ion pairs (Na⁺-Cl⁻ and Pb²⁺-Cl⁻) are detected in bulk solutions that are more apparent with increase of ionic concentrations. Even at low ionic concentrations (0.16 mol/L), a portion of Cl⁻ ions are strongly adsorbed on the octahedral AlO₆ surface as inner-sphere mode. In contrast, at 0.16 mol/L, all Na⁺ and Pb²⁺ ions are relatively far from the tetrahedral SiO₄ surface and for each metal ion, 35.7% are presented as outer-sphere adsorption (Tables 1 and 2), as testified by radial distribution functions (RDF) in Figure S15. Outer-sphere Na⁺ and Pb²⁺ ions departure from the tetrahedral SiO₄ surfaces by approximately 4.9 and 5.1 Å and are coordinated with an average of 4.2 and 5.3 water molecules, respectively. The coordination numbers of water molecules are slightly less than in bulk solutions (e.g., 5.6 for Pb²⁺), suggesting the influence by kaolinite surfaces although gently.

The outer-sphere Na⁺ ions are doubled when ionic concentrations change from 0.16 to 0.32 mol/L while do not expand so obviously with the further increase of ionic concentrations (Table 1). It indicates that the outer-sphere adsorption sites are gradually saturated, and as a result, a larger proportion of Na⁺ ions remain in bulk solutions and inner-sphere mode emerges at concentrated solutions (one and two at 0.80 and 0.96 mol/L, respectively). Inner-sphere Na⁺ ions are not situated at the center of hexagonal cavity but instead interact mainly with one O_b atoms (Fig. 3a), and such adsorption configurations are caused by the involvement of water solvent during the adsorption processes rather than the small ionic size of Na⁺ as proposed before²⁷, which is further testified by density functional calculations in the

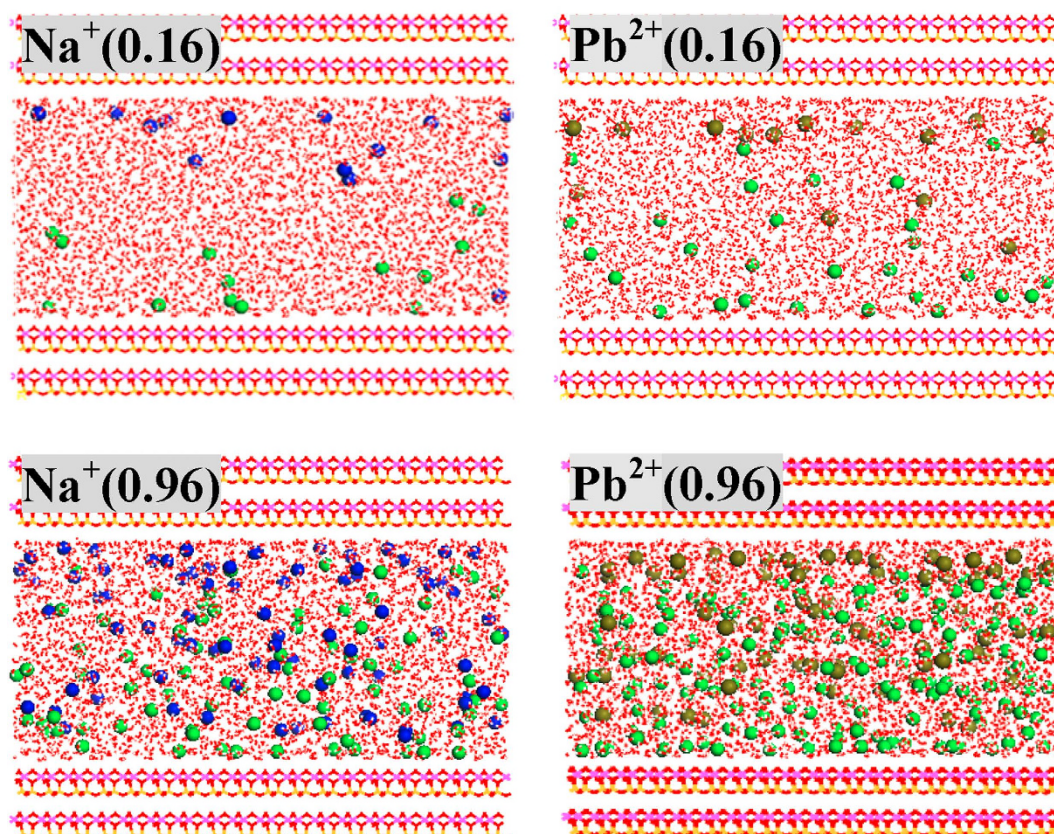


Figure 2. Configurations of regular kaolinite in equilibrium with 0.16 and 0.96 mol/L NaCl/PbCl₂ solutions. Ionic concentrations (mol/L) are indicated in the parentheses of legends. Na⁺, Pb²⁺ and Cl⁻ ions are represented as blue, dark yellow and green balls, respectively.

supplementary information. The gradual increase of ionic concentrations also increase the number of outer-sphere Pb²⁺ ions until saturated (Table 2); however, even if saturated, inner-sphere mode is absent and this differs from the case of Na⁺ adsorption²⁷.

The time-dependent trajectories of both Na⁺ and Pb²⁺ ions are dispersive for the whole ionic concentration range (Figure S16), suggesting that adsorbed metal ions on regular kaolinite are generally highly mobile and interact weakly with the tetrahedral SiO₄ surface. In order to quantify the stabilities of metal ions, root mean square fluctuations (RMSF) are calculated (Figure S17 and S18) and then partitioned into four groups (Table 3): (1) RMSF ≤ 1.2 Å. Strong inner-sphere adsorption; (2) 1.2 Å < RMSF ≤ 1.7 Å. Remaining inner-sphere and strong outer-sphere metal ions; (3) 1.7 Å < RMSF ≤ 2.7 Å. Remaining outer-sphere metal ions as well as those in bulk solutions but with comparable stability; (4) RMSF > 2.7 Å. The others in bulk solutions. At 0.16 mol/L, almost all outer-sphere Pb²⁺ and Na⁺ ions have RMSF > 1.7 Å and have stabilities close to those in bulk solutions. At 0.96 mol/L, a majority of Na⁺ and Pb²⁺ ions in bulk solutions fall within Group 4 instead of Group 3 as in 0.16 mol/L, implying the reduction of stabilities for metal ions in bulk solutions with increase of ionic concentrations. At the same time, stabilities of some outer-sphere metal ions are improved, especially in the case of Na⁺ ions. Seven strong outer-sphere Na⁺ ions are ascribed with RMSF ≤ 1.7 Å. As to the two inner-sphere Na⁺ ions, one binds tightly on the tetrahedral SiO₄ surface (RMSF ≤ 1.2 Å) while the other is relatively more mobile (1.2 Å < RMSF ≤ 1.7 Å).

Defect sites. Configurations of NaCl and PbCl₂ solutions in equilibrium with defective kaolinite (Si₁) are displayed in Fig. 4, S19 and S20. A totally different scenario from the case of regular models (Si₀) has been observed. At 0.16 mol/L, all Na⁺ ions can be considered adsorbed, and 12 out of 14 are inner-sphere (Table 1) with distances of ca. 2.20 Å from the tetrahedral SiO₄ surface as indicated by RDF plots (Figure S21). Accordingly, the distribution of inner- and outer-sphere Na⁺ ions is greatly altered by defect sites and the strong inner-sphere mode is significantly promoted.

With increase of ionic concentrations, the numbers of both inner- and outer-sphere Na⁺ ions show a gradual increase, while their proportions may decline and more Na⁺ ions will remain in bulk solutions. The inner-sphere adsorption sites seem to be filled up since 0.64 mol/L, and a maximum of 25 inner-sphere Na⁺ ions are detected for Si₁ instead of merely 2 for Si₀. Thus, defect sites substantially enhance the inner-sphere adsorption sites that are known to be strong and stable^{9–12,19,26,27}. In fact,

	c (mol/L)	Inner-sphere adsorption			Outer-sphere adsorption		
		X_{ad}	N_{ad}	Γ	X_{ad}	N_{ad}	Γ
Si_0	0.16	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	35.7% \pm 7.1%	5 \pm 1	0.22 \pm 0.04
	0.32	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	35.7% \pm 3.6%	10 \pm 1	0.45 \pm 0.04
	0.48	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	23.8% \pm 4.8%	10 \pm 2	0.45 \pm 0.09
	0.64	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	21.4% \pm 3.6%	12 \pm 2	0.53 \pm 0.09
	0.80	1.4% \pm 0.0%	1 \pm 0	0.04 \pm 0.00	17.1% \pm 4.3%	12 \pm 3	0.53 \pm 0.13
	0.96	2.4% \pm 1.2%	2 \pm 1	0.09 \pm 0.04	16.7% \pm 3.6%	14 \pm 3	0.62 \pm 0.13
$Si_1(9)$	0.16	85.7% \pm 0.0%	12 \pm 0	0.53 \pm 0.00	14.3% \pm 0.0%	2 \pm 0	0.09 \pm 0.00
	0.32	60.7% \pm 3.6%	17 \pm 1	0.76 \pm 0.04	32.1% \pm 7.1%	9 \pm 2	0.40 \pm 0.09
	0.48	45.2% \pm 2.4%	19 \pm 1	0.85 \pm 0.04	31.0% \pm 2.4%	13 \pm 1	0.58 \pm 0.04
	0.64	41.1% \pm 3.6%	23 \pm 2	1.02 \pm 0.09	19.6% \pm 1.8%	11 \pm 1	0.49 \pm 0.04
	0.80	35.7% \pm 4.3%	25 \pm 3	1.11 \pm 0.13	15.7% \pm 2.9%	11 \pm 2	0.49 \pm 0.09
	0.96	28.6% \pm 3.6%	24 \pm 3	1.07 \pm 0.13	17.9% \pm 3.6%	15 \pm 3	0.67 \pm 0.13
$Si_2(9)$	0.16	92.9% \pm 7.1%	13 \pm 1	0.58 \pm 0.04	7.1% \pm 7.1%	1 \pm 1	0.04 \pm 0.04
	0.32	67.9% \pm 3.6%	19 \pm 1	0.85 \pm 0.04	25.0% \pm 7.1%	7 \pm 2	0.31 \pm 0.09
	0.48	47.6% \pm 4.8%	20 \pm 2	0.89 \pm 0.09	21.4% \pm 4.8%	9 \pm 2	0.40 \pm 0.09
	0.64	44.6% \pm 1.8%	25 \pm 1	1.11 \pm 0.04	19.6% \pm 1.8%	11 \pm 1	0.49 \pm 0.04
	0.80	38.6% \pm 2.9%	27 \pm 2	1.20 \pm 0.09	18.6% \pm 4.3%	13 \pm 3	0.58 \pm 0.13
	0.96	31.0% \pm 3.6%	26 \pm 3	1.16 \pm 0.13	17.9% \pm 3.6%	15 \pm 3	0.67 \pm 0.13
$Si_3(9)$	0.16	85.7% \pm 7.1%	12 \pm 1	0.53 \pm 0.04	14.3% \pm 7.1%	2 \pm 1	0.09 \pm 0.04
	0.96	35.7% \pm 2.4%	30 \pm 2	1.33 \pm 0.09	19.0% \pm 2.4%	16 \pm 2	0.71 \pm 0.09
$Si_4(9)$	0.16	92.9% \pm 7.1%	13 \pm 1	0.58 \pm 0.04	7.1% \pm 7.1%	1 \pm 1	0.04 \pm 0.04
	0.96	41.7% \pm 3.6%	35 \pm 3	1.56 \pm 0.13	19.0% \pm 2.4%	16 \pm 2	0.71 \pm 0.09
$Si_1(18)$	0.96	36.9% \pm 1.2%	31 \pm 1	1.38 \pm 0.04	19.0% \pm 1.2%	16 \pm 1	0.71 \pm 0.04
$Si_1(27)$	0.96	47.1% \pm 3.6%	40 \pm 3	1.78 \pm 0.13	21.4% \pm 2.4%	18 \pm 2	0.80 \pm 0.09

Table 1. Percentage (X_{ad}), number (N_{ad}) and adsorption capacity (Γ , $\mu\text{mol}/\text{m}^2$) of Na^+ ions adsorbed on the kaolinite surfaces, where Na^+ ions are from NaCl solutions with a wide range of concentrations^a.

^aThe number of defect sites are given in parentheses, and the default is 9.

outer-sphere mode is also affected although relatively slightly. At each concentration, the number of outer-sphere Na^+ ions differs from that of Si_0 ; in addition, the distances to the tetrahedral SiO_4 surface show a decrease suggesting the reinforced interactions; e.g., at 0.16 mol/L about 4.7 and 4.9 Å for Si_1 and Si_0 , respectively (Fig. 5). A straightforward way to show how the distances are altered by defect sites is given in Fig. 5 that is obtained as the average of the last 2.0 ns trajectories. At 0.16 mol/L, only two Na^+ ions are less than 5.0 Å in Si_0 , while all Na^+ ions fall below 5.0 Å in Si_1 and most of them are not more than 2.5 Å (inner-sphere). Three different interactions are observed for Si_1 and inner-sphere Na^+ ions: One resembles that in Si_0 (Fig. 3a) and the other two are closely associated with defect sites, respectively interacting through the O_h atom of silanol (Fig. 3b) and the bridging O_a atom that bonds to the silanol Si atom (Fig. 3c). Thus, it clearly demonstrates the important role of defect sites played during the adsorption processes, consistent with the experimental results^{16,25,35,48,49}.

At 0.16 mol/L, ten Pb^{2+} ions are outer-sphere adsorbed on the tetrahedral SiO_4 surface of Si_1 , twice as that of Si_0 (Fig. 4 and Table 2). This is mainly due to the presence of hydrophilic silanol groups in Si_1 ^{50–52}, which improve the adsorption of water molecules and further outer-sphere Pb^{2+} ions. As a result of enhanced interactions, outer-sphere Pb^{2+} ions in Si_1 are more close to the tetrahedral SiO_4 surface than in Si_0 (4.3 vs. 5.1 Å, see Fig. 5). The outer-sphere adsorption sites have been filled up since 0.48 mol/L, and at most 22 Pb^{2+} ions can be adsorbed in this mode. Figure 5 shows that three Pb^{2+} ions have distances of ca. 2.91 Å from the tetrahedral SiO_4 surface that belong to inner-sphere mode. Unlike the outer-sphere case, the inner-sphere adsorption quantities seem not to be much affected by increasing ionic concentrations, and four are ascribed to this mode at 0.96 mol/L with distances of ca. 2.99 Å from the tetrahedral SiO_4 surface (Fig. 5). More importantly, all inner-sphere Pb^{2+} ions are closely associated with defect sites (Fig. 3f), different from the case of Na^+ adsorption. Accordingly, Pb^{2+} adsorption is fundamentally changed by defect sites, not only enhancing the outer-sphere adsorption quantities but also creating the inner-sphere mode.

	c (mol/L)	Inner-sphere adsorption			Outer-sphere adsorption		
		X_{ad}	N_{ad}	Γ	X_{ad}	N_{ad}	Γ
Si_0	0.16	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	35.7% \pm 7.1%	5 \pm 1	0.22 \pm 0.04
	0.32	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	25.0% \pm 3.6%	7 \pm 1	0.31 \pm 0.04
	0.48	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	19.0% \pm 2.4%	8 \pm 1	0.36 \pm 0.04
	0.64	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	23.2% \pm 1.8%	13 \pm 1	0.58 \pm 0.04
	0.80	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	22.9% \pm 2.9%	16 \pm 2	0.71 \pm 0.09
	0.96	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	19.0% \pm 2.4%	16 \pm 2	0.71 \pm 0.09
$Si_1(9)$	0.16	21.4% \pm 0.0%	3 \pm 0	0.13 \pm 0.00	71.4% \pm 7.1%	10 \pm 1	0.44 \pm 0.04
	0.32	10.7% \pm 0.0%	3 \pm 0	0.13 \pm 0.00	64.3% \pm 7.1%	18 \pm 2	0.80 \pm 0.09
	0.48	7.1% \pm 0.0%	3 \pm 0	0.13 \pm 0.00	47.6% \pm 4.8%	20 \pm 2	0.89 \pm 0.09
	0.64	5.4% \pm 0.0%	3 \pm 0	0.13 \pm 0.00	35.7% \pm 1.8%	20 \pm 1	0.89 \pm 0.04
	0.80	4.3% \pm 0.0%	3 \pm 0	0.13 \pm 0.00	31.4% \pm 4.3%	22 \pm 3	0.98 \pm 0.13
	0.96	4.8% \pm 1.2%	4 \pm 1	0.18 \pm 0.04	25.0% \pm 2.4%	21 \pm 2	0.93 \pm 0.09
$Si_2(9)$	0.16	28.6% \pm 0.0%	4 \pm 0	0.18 \pm 0.00	71.4% \pm 7.1%	10 \pm 1	0.44 \pm 0.04
	0.32	14.3% \pm 3.6%	4 \pm 1	0.18 \pm 0.04	64.3% \pm 7.1%	18 \pm 2	0.80 \pm 0.09
	0.48	9.5% \pm 2.4%	4 \pm 1	0.18 \pm 0.04	47.6% \pm 2.4%	20 \pm 1	0.89 \pm 0.04
	0.64	9.0% \pm 0.0%	5 \pm 0	0.22 \pm 0.00	37.5% \pm 3.6%	21 \pm 2	0.93 \pm 0.09
	0.80	7.1% \pm 0.0%	5 \pm 0	0.22 \pm 0.00	35.7% \pm 4.3%	25 \pm 3	1.11 \pm 0.13
	0.96	4.8% \pm 1.2%	4 \pm 1	0.18 \pm 0.04	29.8% \pm 2.4%	25 \pm 2	1.11 \pm 0.06
$Si_3(9)$	0.16	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	85.7% \pm 7.1%	12 \pm 1	0.53 \pm 0.04
	0.96	0.0% \pm 0.0%	0 \pm 0	0.00 \pm 0.00	34.5% \pm 3.6%	29 \pm 3	1.29 \pm 0.13
$Si_4(9)^b$	0.16	14.3% \pm 0.0%	2 \pm 0	0.09 \pm 0.00	78.6% \pm 7.1%	11 \pm 1	0.49 \pm 0.04
	0.96	2.4% \pm 0.0%	2 \pm 0	0.09 \pm 0.00	38.0% \pm 2.4%	32 \pm 2	1.42 \pm 0.09
$Si_1(18)$	0.96	4.8% \pm 1.2%	4 \pm 1	0.18 \pm 0.04	29.8% \pm 2.4%	25 \pm 2	1.11 \pm 0.09
$Si_1(27)$	0.96	4.8% \pm 1.2%	4 \pm 1	0.18 \pm 0.04	34.5% \pm 2.4%	29 \pm 2	1.29 \pm 0.09

Table 2. Percentage (X_{ad}), number (N_{ad}) and adsorption capacity (Γ , $\mu\text{mol}/\text{m}^2$) of Pb^{2+} ions adsorbed on the kaolinite surfaces, where Pb^{2+} ions are from PbCl_2 solutions with a wide range of concentrations^a. ^aNumbers of defect sites are given in parentheses, and the default is 9. ^bTwo quasi inner-sphere adsorbed Pb^{2+} ions have been taken into account.

Stabilities of metal ions adsorbed on Si_1 are addressed by the time-evolution trajectories (Fig. 6 and S22) and RMSF analyses (Table 3). The trajectories of Si_1 are obviously more concentrated than those of Si_0 . Inner-sphere metal ions are restricted at the adsorption sites; meanwhile, outer-sphere Pb^{2+} ions are also stabilized by the defect sites through the tight anchoring of intermediate water molecules. RMSF data show that a considerable amount of metal ions is transferred to Groups 1 and 2; e.g., in addition to inner-sphere ones, another four and seven Pb^{2+} ions have $\text{RMSF} \leq 1.7 \text{ \AA}$ respectively at 0.16 and 0.96 mol/L that should belong to strong outer-sphere adsorption. Accordingly, defect sites improve the stabilities of both inner- and outer-sphere metal ions and generate long-lived adsorbed species. Both of the adsorption quantities and stabilities are significantly promoted by defect sites, which give a reasonable interpretation for the good adsorption behaviors observed experimentally^{2,28,29}.

Extending defect sites. Figure 7, S23 and S24 display the equilibrium configurations of kaolinite with a larger defect site (Si_2) in contact with NaCl and PbCl_2 solutions. The adsorption behaviors are similar for Si_2 and Si_1 , and both inner-sphere Na^+ and Pb^{2+} ions occur at the whole concentration ranges (0.16 ~ 0.96 mol/L). According to the types of bonded $\text{O}(\text{O}_a/\text{O}_b/\text{O}_h)$ atoms, three (Fig. 3a,d,e) and one (Fig. 3g) interactions are respectively assigned for inner-sphere Na^+ and Pb^{2+} ions that resemble those of Si_1 . At 0.16 mol/L, most $\text{Na}^+/\text{Pb}^{2+}$ ions are already inner-sphere/outer-sphere adsorbed in Si_1 and hence the adsorption quantities will remain almost unaltered by enlargement of defect sizes to Si_2 (Tables 1 and 2). With increase of ionic concentrations, the adsorption quantities of Si_2 may show increases as compared to Si_1 , and inner-sphere Na^+ and outer-sphere Pb^{2+} ions of Si_2/Si_1 amount to 26/24 and 25/21 at 0.96 mol/L, respectively. Meanwhile, Si_2 is more structurally flexible and this facilitates the interactions with metal ions. The adsorption energies of inner-sphere $\text{Na}^+/\text{Pb}^{2+}$ ions with defective sites (E_K) are approximately $-323.5 \pm 4.3/-331.1 \pm 4.8$ and $-334.8 \pm 4.0/-342.8 \pm 4.1 \text{ kJ/mol}$ for Si_1 and Si_2 , respectively. As a result of reinforced interactions, distances to the tetrahedral SiO_4 surface

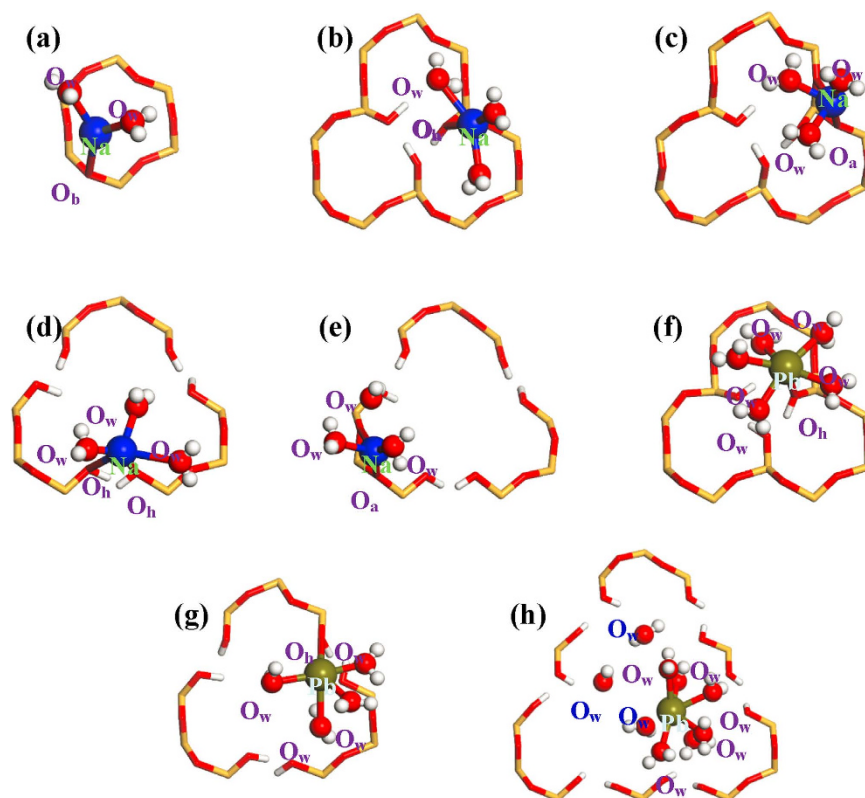


Figure 3. Local structures of inner-sphere Na^+ and Pb^{2+} ions adsorbed on the tetrahedral SiO_4 surfaces of regular (Si_0) and defective kaolinite (Si_1 , Si_2 and Si_4). (a) for $\text{Si}_0(\text{Na}^+)$, (a–c) for $\text{Si}_1(\text{Na}^+)$, (a,d,e) for $\text{Si}_2(\text{Na}^+)$ as well as (f) for $\text{Si}_1(\text{Pb}^{2+})$, (g) for $\text{Si}_2(\text{Pb}^{2+})$ and (h) for $\text{Si}_4(\text{Pb}^{2+})$, respectively. Note that $\text{Si}_4(\text{Pb}^{2+})$ is actually for quasi inner-sphere Pb^{2+} ions adsorbed on the water molecules within the silanol nests of Si_4 .

c (mol/L)	RMSF (Å)	Na^+					Pb^{2+}				
		Si_0	Si_1	Si_2	Si_3	Si_4	Si_0	Si_1	Si_2	Si_3	Si_4
0.16	≤ 1.2	0	6	9	10	10	0	1	4	0	2
	1.2 ~ 1.7	0	6	4	3	4	1	6	9	10	10
	1.7 ~ 2.7	12	2	1	1	0	11	7	1	4	2
	> 2.7	2	0	0	0	0	2	0	0	0	0
0.96	≤ 1.2	1	18	21	26	30	0	3	4	0	2
	1.2 ~ 1.7	8	16	18	18	19	3	8	16	20	24
	1.7 ~ 2.7	28	22	24	21	17	36	43	41	44	40
	> 2.7	47	28	21	19	18	45	30	23	20	18

Table 3. Numbers of Na^+ and Pb^{2+} ions falling with the specified RMSF ranges for NaCl and PbCl_2 solutions in equilibrium with the tetrahedral SiO_4 surfaces of kaolinite.

are slightly shortened with enlargement of defect sites (Fig. 5). There are also inner-sphere Na^+ ions that have no relation with defect sites, and their interactions are obviously weaker as reflected by the E_K values that are equal to -230.9 ± 5.3 kJ/mol for Si_1 and -239.2 ± 4.9 kJ/mol for Si_2 , respectively. As a result, inner-sphere Pb^{2+} ions can be detected only at defect sites that enhance the interactions (Fig. 6).

In order to clarify the effect of defect sizes, two larger defects (Si_3 and Si_4 , see Fig. 1b) are considered, and their equilibrium configurations in 0.16 and 0.96 mol/L NaCl and PbCl_2 solutions are given in Figures S25, S26. At low ionic concentrations (e.g., 0.16 mol/L), the further enlargement of defect sizes from Si_2 to Si_3 and Si_4 causes slight alterations to the adsorption of inner-sphere Na^+ and outer-sphere Pb^{2+} ions, which is in agreement with the results of Si_2 vs. Si_1 discussed earlier; at 0.96 mol/L, however, the numbers of inner-sphere Na^+ ions of Si_3 and Si_4 show observable increases and amount to 30 and 35, respectively, while at any given concentration, the quantities of outer-sphere Na^+ ions remain almost

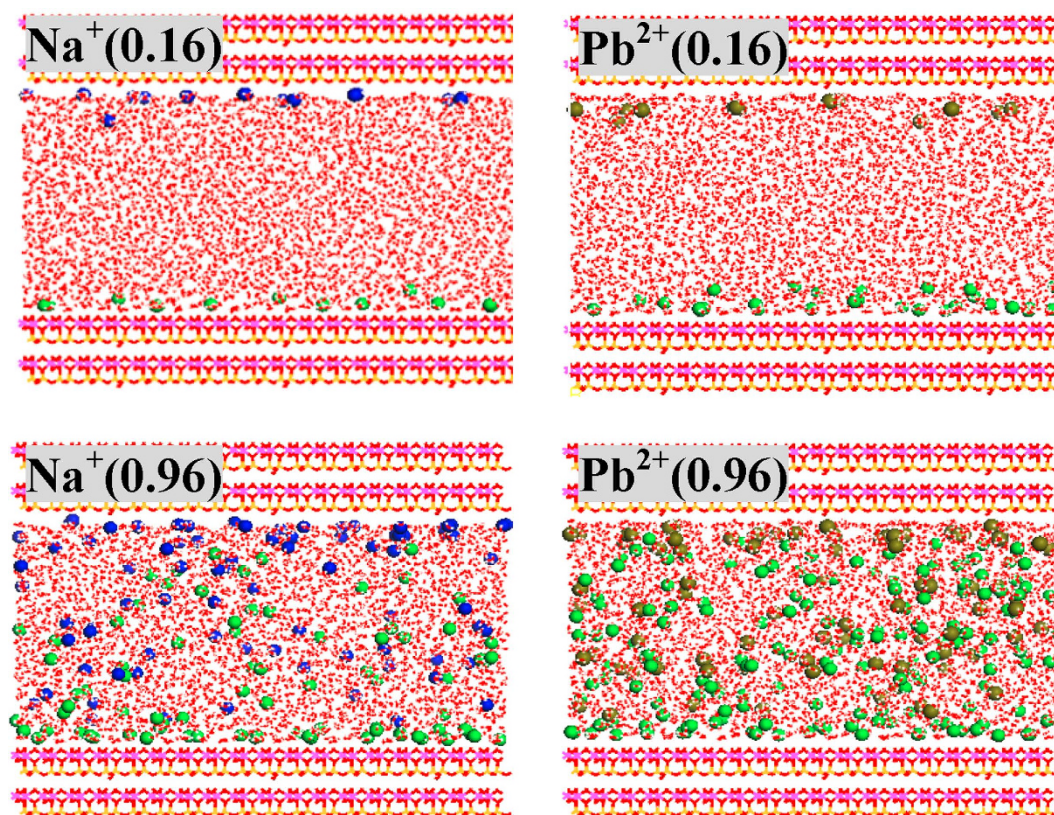


Figure 4. Configurations of defective kaolinite (Si_i) in equilibrium with 0.16 and 0.96 mol/L NaCl/PbCl₂ solutions. Ionic concentrations (mol/L) are indicated in the parentheses of legends. Na^+ , Pb^{2+} and Cl^- ions are represented as blue, dark yellow and green balls, respectively.

constant for all defective models (Table 1). Figure 3 and S27 indicate that the adsorption structures of inner-sphere Na^+ ions are similar for all defect sites. With increase of defect sizes, however, inner-sphere Na^+ ions have an attendance to be closer to the center of hydrophilic silanol nests, probably due to the attraction by water molecules inside these silanol nests. Similarly, at high concentrations, outer-sphere Pb^{2+} ions are promoted by the further enlargement of defect sites, which, at 0.96 mol/L, are counted to be 25, 29 and 32 for Si_2 , Si_3 and Si_4 , respectively (Table 2). In the case of Si_4 , the inner-sphere Pb^{2+} species vanishes at both 0.16 and 0.96 mol/L, and a quasi inner-sphere mode emerges instead, where Pb^{2+} ions form direct bonds with water molecules that have entered into the silanol nests (Fig. 3h)³¹. Although not associated directly with kaolinite, the vertical distances of quasi inner-sphere Pb^{2+} ions to the tetrahedral SiO_4 surface fall within the scope of inner-sphere mode.

Stabilities of adsorbed metal ions improve gradually with the enlargement of defect sites from Si_1 to Si_2 and then to Si_3 and Si_4 , in line with the adsorption results on zeolites and graphene^{35,53}. As indicated by the time-evolution trajectories (Fig. 6, S22, S28 and S29), adsorbed metal ions become more concentrated on the tetrahedral SiO_4 surface due to the increase of defect sizes; meanwhile, more metal ions are directly associated with defect sites; e.g., the numbers of such inner-sphere Na^+ ions at 0.16/0.96 mol/L are 5/9, 8/13, 8/15, 9/19 for Si_1 , Si_2 , Si_3 and Si_4 , respectively. A more elaborate analysis of stabilities is carried out based on the RMSF data (Table 3). Inner-sphere Na^+ ions at 0.16/0.96 mol/L with $\text{RMSF} \leq 1.2 \text{ \AA}$ are counted to be 0/1, 6/18, 9/21, 10/26 and 10/30 for Si_0 , Si_1 , Si_2 , Si_3 and Si_4 , respectively. It clearly shows that the increase of defect sizes significantly stabilizes the inner-sphere Na^+ ions, especially at higher ionic concentrations. Meanwhile, other Na^+ ions are also stabilized, and so fewer will stay with $\text{RMSF} > 2.7 \text{ \AA}$. With the gradual enlargement of defect sizes, similar changing trends are observed for the stabilities of Pb^{2+} ions. At 0.16/0.96 mol/L, Pb^{2+} ions with strong outer-sphere adsorption ($1.2 \text{ \AA} < \text{RMSF} \leq 1.7 \text{ \AA}$) are counted to be 1/3, 4/7, 9/16, 10/20 and 10/24 for Si_0 , Si_1 , Si_2 , Si_3 and Si_4 , respectively. The quasi inner-sphere Pb^{2+} ions that occur in the case of Si_4 are greatly stabilized by associated water molecules within the hydrophilic silanol nests and show comparable stabilities as inner-sphere species ($\text{RMSF} < 1.2 \text{ \AA}$, see Table 3).

Adsorption capacities and mechanisms. The adsorption capacities (I) of Na^+ and Pb^{2+} ions on the tetrahedral SiO_4 surfaces of kaolinite are calculated for the various defect sites (Si_0 , Si_1 , Si_2 , Si_3 and Si_4) and a wide range of ionic concentrations (0.16 ~ 0.96 mol/L), see Tables 1 and 2 and Figures S30, S31.

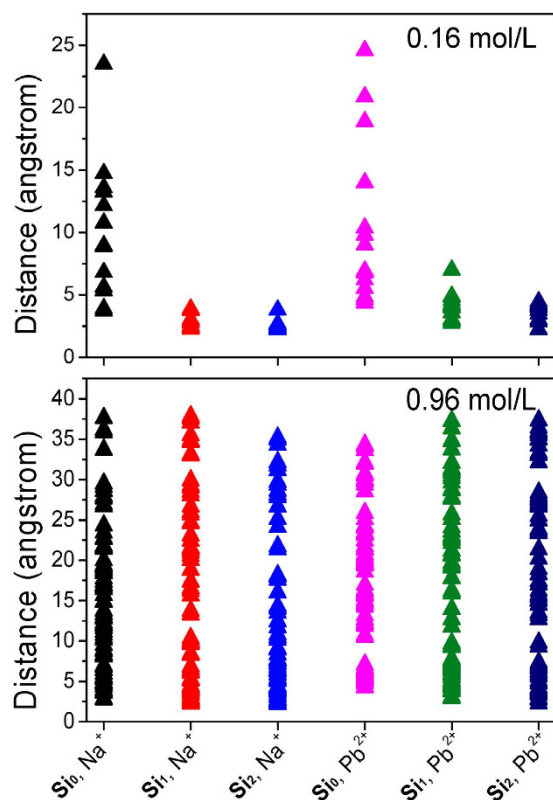


Figure 5. Distances between the tetrahedral SiO_4 surfaces of kaolinite and $\text{Na}^+/\text{Pb}^{2+}$ ions from $\text{NaCl}/\text{PbCl}_2$ solutions.

For a given model (e.g., Si_1), the adsorption quantities of both inner- and outer-sphere metal ions generally increase with ionic concentrations until saturated. Outer- and inner-sphere Na^+ ions are respectively the principal adsorbed species for regular and defective models, and the amounts of the former remain close for all models (regular and defective) while those of the latter increase with enlargement of defect sizes. Outer-sphere Pb^{2+} ions that predominate for all models show an increase with enlargement of defect sizes. At 0.96 mol/L, the adsorption capacities of inner-sphere $\text{Na}^+/\text{outer-sphere } \text{Pb}^{2+}$ ions amount to 0.09/0.71, 1.07/0.93, 1.16/1.11, 1.33/1.29 and 1.56/1.42 $\mu\text{mol}/\text{m}^2$ for Si_0 , Si_1 , Si_2 , Si_3 and Si_4 , accounting for 12.68%/100.00%, 61.49%/83.78%, 63.39%/100.00% and 68.72%/94.04% of the total adsorption capacities, respectively.

Defect contents may also affect the adsorption processes^{49,54}, and so 18 and 27 defect sites of Si_1 type are considered. Figure S32 shows their configurations in equilibrium with 0.96 mol/L NaCl and PbCl_2 solutions. The adsorption behaviors of both Na^+ and Pb^{2+} ions remain the same for all defect contents, and the adsorption sites seem to have a finite rate of increase when the number of defect sites (n) is doubled ($n=18$) and then tripled ($n=27$), see Tables 1 and 2 and Figure S29. Inner-sphere Na^+ ions that predominate for all defective models are the most affected and the adsorption capacities amount to 1.07, 1.38 and 1.78 $\mu\text{mol}/\text{m}^2$ respectively for $n=9$, 18 and 27 (0.96 mol/L). The adsorption capacities of outer-sphere Na^+ and Pb^{2+} ions increase relatively slowly with increase of defect contents, while inner-sphere Pb^{2+} ions that are minor remain constant for all defect contents. Stabilities of adsorbed metal ions improve with increase of defect contents (Table S2): RMSFs of more Na^+ ions and Pb^{2+} ions fall respectively below 1.2 Å (strong inner-sphere) and 1.7 Å (strong outer-sphere), and fewer metal ions remain with $\text{RMSF} > 2.7$ Å.

As discussed above, defect sites promote the adsorption of metal ions, with respect to both adsorption quantities and stabilities; in addition, the adsorption capacities are closely associated with ionic concentrations, defect sizes and contents. Adsorption of metal ions and anions may proceed in a collaborative mechanism, see Fig. 8. The octahedral AlO_6 surface with slightly positive charges attracts Cl^- ions strongly and so an average of 2.45 more Cl^- ions will be adsorbed than M^+ ions. Defect sites are developed on the tetrahedral SiO_4 surface and have no observable impact on Cl^- ions, while the adsorption quantities of Cl^- ions show a linear increase along with those of metal ions (M^+). Accordingly, metal ions and anions can promote each other during the adsorption processes. The slope of the fitted line in Fig. 8 is approximately 1.00 and this demonstrates that on average, the adsorption of one more M^+ ion results in the increase of one adsorbed Cl^- ion; vice versa.

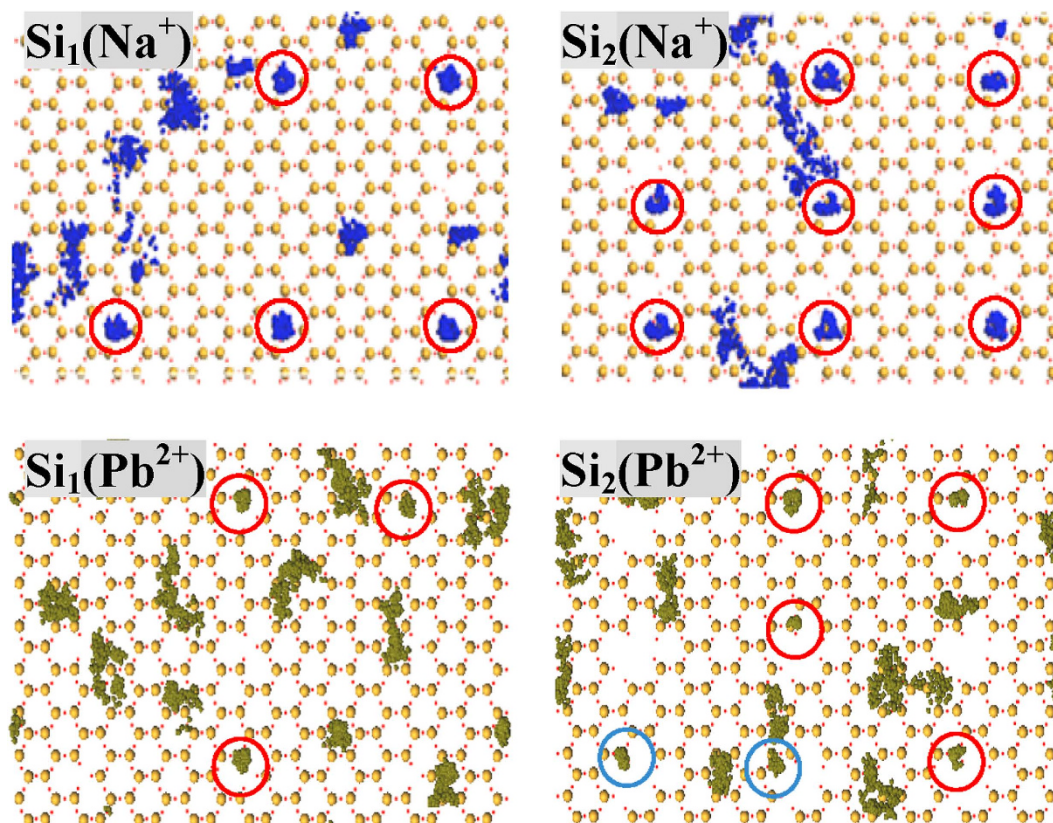


Figure 6. Trajectory maps of tetrahedral SiO_4 surfaces of defective kaolinite (Si_1 and Si_2) in contact with 0.16 mol/L NaCl and PbCl_2 solutions. Only the inner- and outer-sphere metal ions are made visible, and Na^+ and Pb^{2+} ions are represented as blue and dark yellow balls, respectively. Inner- and outer-sphere metal ions that fall around the defective sites are highlighted by red and light blue circles, respectively.

Computational Details

Models. Kaolinite has alternative tetrahedral SiO_4 sheet and octahedral AlO_6 sheet that are bridged by O atoms. The crystal structure of kaolinite was taken from ref. 55 with cell parameters being $a = 5.19 \text{ \AA}$, $b = 8.96 \text{ \AA}$, $c = 7.36 \text{ \AA}$, $\alpha = 90.77^\circ$, $\beta = 104.17^\circ$ and $\gamma = 90.40^\circ$. Models were prepared in line with the report of Vasconcelos *et al.*²⁷. Firstly, the orthogonal transformation was conducted causing the (001) and (00 $\bar{1}$) planes of kaolinite respectively to correspond to the octahedral AlO_6 and tetrahedral SiO_4 surfaces; secondly, models with 324 unit cells ($9 \times 9 \times 4$ along x , y and z directions) were constructed, and the inward octahedral AlO_6 and tetrahedral SiO_4 surfaces were separated by a vacuum layer with thickness of 40 \AA ; thirdly, the aqueous-kaolinite interfaces were achieved by filling 5097 water molecules into the vacuum layer. Because of the choice of a thicker vacuum layer, more water molecules are included in this work than previously²⁷, resulting in a total of 26307 atoms; finally, NaCl and PbCl_2 solutions with a wide range of concentrations (0.16, 0.32, 0.48, 0.64, 0.80 and 0.96 mol/L) were respectively prepared by replacing certain numbers of water molecules with metal cations and anions. These models are referred to “regular” (Si_0), and Fig. 1a illustrates the initial configuration of regular kaolinite (Si_0) in contact with 0.16 mol/L NaCl solutions.

A series of defect sites^{5,6,31–36,56,57} were then constructed on the tetrahedral SiO_4 surface of regular model (Fig. 1b), by removing one lattice Si atom (Si_1) as well as first-shell Si atoms (Si_2) and further second-shell Si atoms (Si_4), respectively. In addition, Si_3 was considered where the six Si atoms of one hexagonal cavity were all leached. The hexagonal cavity was assumed to be the main loci for situating metal ions²⁷. Density functional calculated results in the supplementary information showed that defect sites are facile to form in clay minerals. The defect sizes follow as $\text{Si}_1 < \text{Si}_2 < \text{Si}_3 < \text{Si}_4$, with removal of 1, 4, 6 and 10 lattice Si atoms, respectively. Note that all these defect sites are far smaller than the trenches reported by Croteau³¹. The dangling O atoms of defect sites were saturated by H atoms to ensure electro-neutrality. Unless otherwise noted, the number of defect sites is 9 for defective models. To investigate the effect of defect contents on the adsorption processes, 18 and 27 defect sites of Si_1 type were also taken in account. All concentrations of NaCl and PbCl_2 solutions indicated above were simulated for Si_0 , Si_1 and Si_2 , as well as 0.16 and 0.96 mol/L for Si_3 , Si_4 and 0.96 mol/L for Si_1 with 18 and 27 defect sites, resulting in a total of 48 models.

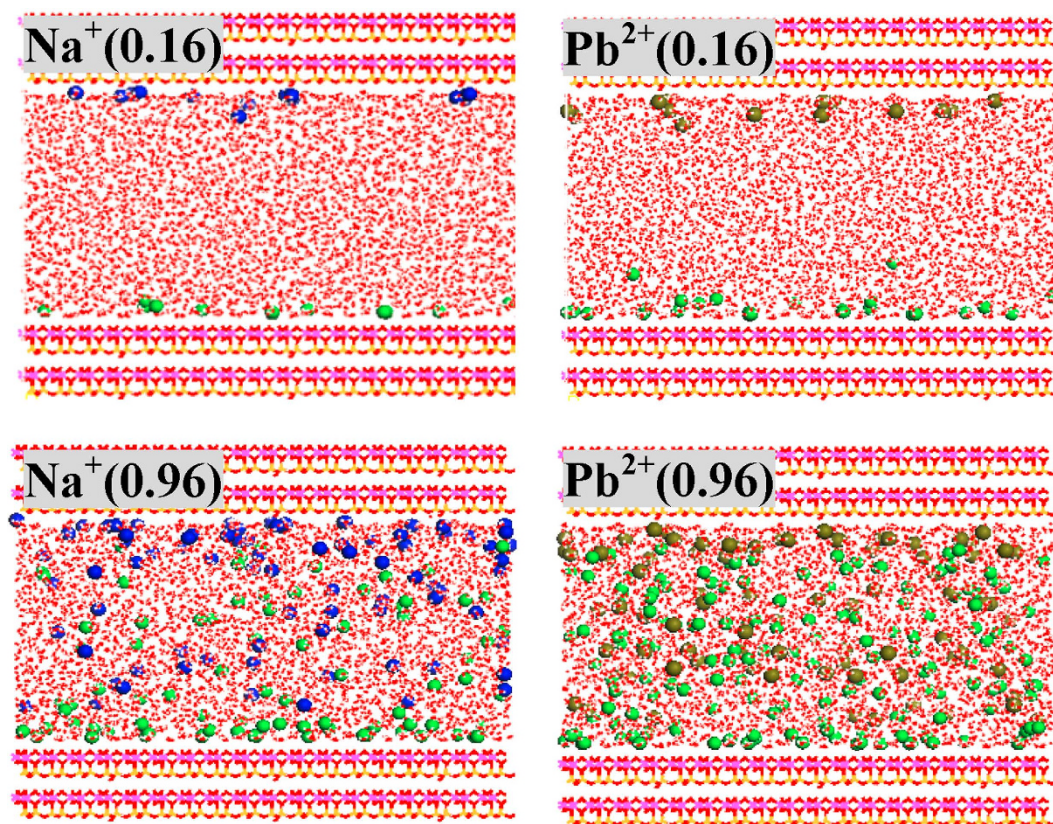


Figure 7. Configurations of defective kaolinite (Si_2) in equilibrium with 0.16 and 0.96 mol/L NaCl/PbCl₂ solutions. Ionic concentrations (mol/L) are indicated in the parentheses of legends. Na^+ , Pb^{2+} and Cl^- ions are represented as blue, dark yellow and green balls, respectively.

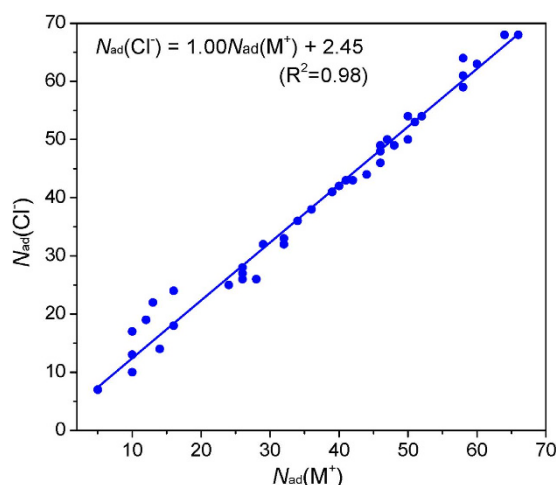


Figure 8. Correlation of the adsorbed quantities between metal cations (M^+) and anions (Cl^-). Both inner- and outer-sphere adsorbed ions are counted. Pb^{2+} is two molar equivalents of M^+ ; i.e., $N_{\text{ad}}(\text{Pb}^{2+}) = 2 \cdot N_{\text{ad}}(\text{M}^+)$.

Methods. MD simulations with periodic boundary conditions (PBC) were performed by using Gromacs suite of programs (Version 4.6.5)⁵⁸. A combination of CLAYFF force field (parameters given in Table S1)⁵⁹ and simple-point-charge (SPC) water model⁶⁰ was used. Recently, the CLAYFF force field has achieved unprecedented success in modeling clay minerals, and the conjunction with SPC water model has been demonstrated to be accurate to simulate hydrated minerals and aqueous-mineral interfaces^{14,23,27,61–65}. Non-bonded (electrostatic and vdW) terms that were developed especially for layered

minerals were parametrized in the CLAYFF force field. The Particle-Mesh-Ewald (PME) method was employed to treat long-range electrostatic interactions, and the cutoff radii for Ewald electrostatic summation and long-range vdW interactions were set to 12.0 Å. The leapfrog algorithm⁶⁶ was used, and 5.0 ns MD simulations were run for each model, at constant temperature and pressure (300 K, 1.0 bar) that were respectively controlled by the V-rescale thermostat⁶⁷ and Parrinello-Rahman barostat⁶⁸. The time step was 2.0 fs, and the atomic coordinates were updated every 1.0 ps. According to the root-mean-square-deviation (RMSD) results (Figures S3, S12), all models have arrived to the equilibrium since 2.0 ns, and hence the last 3.0 ns trajectories were used for analyses. The adsorption energies of metal ions on kaolinite surfaces were calculated on the average structures of 2.0 ~ 5.0 ns MD trajectories, considering the regions within 10.0 Å around the metal ions.

References

1. Bizaia, N. *et al.* Porphyrin-kaolinite as efficient catalyst for oxidation reactions. *ACS Appl. Mater. Interfaces* **1**, 2667–2678 (2009).
2. Gupta, S. S. & Bhattacharyya, K. G. Adsorption of heavy metals on kaolinite and montmorillonite: a review. *Phys. Chem. Chem. Phys.* **14**, 6698–6723 (2012).
3. Batra, M., Gotam, S., Dadarwal, P., Nainwani, R. & Sharma, M. Nano-clay as polymer porosity reducer: a review. *J. Pharm. Sci. Technol.* **3**, 709–716 (2011).
4. Brigatti, M. F., Galan, E. & Theng, B. K. G. Structures and mineralogy of clay minerals, in Bergaya, F. & Lagaly, G. (Eds.), *Developments in clay science* (Vol. 1), Elsevier, Amsterdam, pp. 19–86 (2006).
5. Wu, J. & Buseck, P. R. Carbon storage at defect sites in mantle mineral analogues. *Nature Geosci.* **6**, 875–878 (2013).
6. Kwon, K. Y. *et al.* Defect induced asymmetric pit formation on hydroxyapatite. *Langmuir* **24**, 11063–11066 (2008).
7. Lee, S. S., Fenter, P., Park, C. Y., Sturchio, N. C. & Nagy, K. L. Hydrated cation speciation at the muscovite (001)-water interface. *Langmuir* **26**, 16647–16651 (2010).
8. Catalano, J. G., Park, C. Y., Fenter, P. & Zhang, Z. Simultaneous inner- and outer-sphere arsenate adsorption on corundum and hematite. *Geochim. Cosmochim. Acta* **72**, 1986–2004 (2008).
9. Park, C. Y., Fenter, P. A., Sturchio, N. C. & Nagy, K. L. Thermodynamics, interfacial structure, and pH hysteresis of Rb⁺ and Sr²⁺ adsorption at the muscovite (001)-solution interface. *Langmuir* **24**, 13993–14004 (2008).
10. Loh, S. H. & Jarvis, S. P. Visualization of ion distribution at the mica-electrolyte interface. *Langmuir* **26**, 9176–9178 (2010).
11. Schlegel, M. L. *et al.* Cation sorption on the muscovite (001) surface in chloride solutions using high-resolution X-ray reflectivity. *Geochim. Cosmochim. Acta* **70**, 3549–3565 (2006).
12. Lee, S. S., Fenter, P., Nagy, K. L. & Sturchio, N. C. Changes in adsorption free energy and speciation during competitive adsorption between monovalent cations at the muscovite (001)-water interface. *Geochim. Cosmochim. Acta* **123**, 416–426 (2013).
13. Ricci, M., Spijker, P., Stellacci, F., Molinari, J. F. & Voitchovsky, K. Direct visualization of single ions in the stern layer of calcite. *Langmuir* **29**, 2207–2216 (2013).
14. Ngouana, B. F. W. & Kalinichev, A. G. Structural arrangements of isomorphic substitutions in smectites: molecular simulation of the swelling properties, interlayer structure, and dynamics of hydrated Cs-montmorillonite revisited with new clay models. *J. Phys. Chem. C* **118**, 12758–12773 (2014).
15. Liu, X. D. & Lu, X. C. A Thermodynamic understanding of clay-swelling inhibition by potassium ions. *Angew. Chem. Int. Ed.* **45**, 6300–6303 (2006).
16. Yang, W. & Zaoui, A. Uranyl adsorption on (001) surfaces of kaolinite: a molecular dynamics study. *Appl. Clay Sci.* **80–81**, 98–106 (2013).
17. Churakov, S. V. Mobility of Na and Cs on montmorillonite Surface under partially saturated conditions. *Environ. Sci. Technol.* **47**, 9816–9823 (2013).
18. Suter, J. L., Sprik, M. & Boek, E. S. Free energies of absorption of alkali ions onto beidellite and montmorillonite surfaces from constrained molecular dynamics simulations. *Geochim. Cosmochim. Acta* **91**, 109–119 (2012).
19. Lectez, S., Roques, J., Salanne, M. & Simoni, E. Car-parrinello molecular dynamics study of the uranyl behaviour at the gibbsite/water interface. *J. Chem. Phys.* **137**, 154705 (2012).
20. Greathouse, J. A., O'Brien, R. J., Bemis, G. & Pabalan, R. T. Molecular dynamics study of aqueous uranyl interactions with quartz (010). *J. Phys. Chem. B* **106**, 1646–1655 (2002).
21. Rotenberg, B., Marry, V., Malikova, N. & Turq, P. Molecular simulation of aqueous solutions at clay surfaces. *J. Phys.: Condens. Matter* **22**, 284114 (2010).
22. Ma, Y. M., Zhang, H. & Zhang, B. J. Structure of sodium sulphate aqueous solution/quartz interface: a molecular dynamics simulation. *Mol. Simul.* **40**, 634–639 (2014).
23. Greathouse, J. A. & Cygan, R. T. Molecular dynamics simulation of uranyl (VI) adsorption equilibria onto an external montmorillonite surface. *Phys. Chem. Chem. Phys.* **7**, 3580–3586 (2005).
24. Greathouse, J. A. & Cygan, R. T. Water structure and aqueous uranyl (VI) adsorption equilibria onto external surfaces of beidellite, montmorillonite, and pyrophyllite: results from molecular simulations. *Environ. Sci. Technol.* **40**, 3865–3871 (2006).
25. Yang, W. & Zaoui, A. Behind adhesion of uranyl onto montmorillonite surface: a molecular dynamics study. *J. Hazard. Mater.* **261**, 224–234 (2013).
26. Sakuma, H. & Kawamura, K. Structure and dynamics of water on Li⁺, Na⁺, K⁺, Cs⁺, H₃O⁺-exchanged muscovite surfaces: a molecular dynamics study. *Geochim. Cosmochim. Acta* **75**, 63–81 (2011).
27. Vasconcelos, I. F., Bunker, B. A. & Cygan, R. T. Molecular dynamics modeling of ion adsorption to the basal surfaces of kaolinite. *J. Phys. Chem. C* **111**, 6753–6762 (2007).
28. Bhattacharyya, K. G. & Gupta, S. S. Adsorption of a few heavy metals on natural and modified kaolinite and montmorillonite: a review. *Adv. Colloid Interface Sci.* **140**, 114–131 (2008).
29. Coles, C. A. & Yong, R. N. Aspects of kaolinite characterization and retention of Pb and Cd. *Appl. Clay Sci.* **22**, 39–45 (2002).
30. Miranda-Trevino, J. C. & Coles, C. A. Kaolinite properties, structure and influence of metal retention on Ph. *Appl. Clay Sci.* **310**, 133–139 (2003).
31. Croteau, T., Bertram, A. K. & Patey, G. N. Water adsorption on kaolinite surfaces containing trenches. *J. Phys. Chem. A* **114**, 2171–2178 (2010).
32. Libowitzky, E. & Beran, A. OH defects in forsterite. *Phys. Chem. Minerals* **22**, 387–392 (1995).
33. Etienne, B., Jannick, I., Simon, D., Istvan, K. & Marc, B. Theoretical infrared spectrum of OH-defects in forsterite. *Eur. J. Mineral.* **23**, 285–292 (2011).
34. Nisar, J. C., Arhammar, C., Jamstorp, E. & Ahuja, R. Optical gap and native point defects in kaolinite studied by the GGA-PBE, HSE functional, and GW approaches. *Phys. Rev. B* **84**, 075120 (2011).

35. Yang, G., Pidko, E. A. & Hensen, E. J. M. The mechanism of glucose isomerization to fructose over Sn-BEA zeolite: A periodic density functional theory study. *Chem Sus Chem* **6**, 1688–1696 (2013).
36. Yang, G. & Zhou, L. J. Zwitterionic versus canonical amino acids over the various defects in zeolites: a two-layer ONIOM calculation. *Sci. Rep.* **4**, 6594 (2014).
37. Bayahia, H., Kozhevnikova, E. & Kozhevnikov, I. High catalytic activity of silicalite in gas-phase ketonisation of propionic acid. *Chem. Commun.* **49**, 3842–3844 (2013).
38. Fickel, D. W., Shough, A. M., Doren, D. J. & Lobo, R. F. High-temperature dehydrogenation of defective silicalites. *Micropor. Mesopor. Mater.* **129**, 156–163 (2010).
39. Zhuravlev, L. T. The surface chemistry of amorphous silica. Zhuravlev model. *Colloid Surface A* **173**, 1–38 (2000).
40. Kawai, T. & Tsutsumi, K. Reactivity of silanol groups on zeolite surfaces. *Colloid Polym. Sci.* **276**, 992–998 (1998).
41. Yamagishi, K., Namba, S. & Yashima, T. Defect sites in highly siliceous HZSM-5 zeolites: a study performed by alumination and IR spectroscopy. *J. Phys. Chem.* **95**, 872–877 (1991).
42. Chezeau, J. M., Delmotte, L., Guth, J. L. & Gabelica, Z. Influence of synthesis conditions and postsynthesis treatments on the nature and quantity of structural defects in highly siliceous MFI zeolites: A high-resolution solid-state ^{29}Si n.m.r. study. *Zeolites*. **11**, 598–606 (1991).
43. Forni, L. *et al.* Vapor phase Beckmann rearrangement using high silica zeolite catalyst. *Phys. Chem. Chem. Phys.* **6**, 1842–1847 (2004).
44. Bobon, M., Christy, A. A., Kluvanec, D. & Illasova, L. State of water molecules and silanol groups in opal minerals: a near infrared spectroscopic study of opals from Slovakia. *Phys Chem Minerals* **38**, 809–818 (2011).
45. Khraisheh, M. A. M., Ghouti, M. A. A., Allen, S. J. & Ahmad, M. N. Effect of OH and silanol groups in the removal of dyes from aqueous solution using diatomite. *Water Res.* **39**, 922–932 (2005).
46. Frost, R. L. & Vassallo, A. M. The dihydroxylation of the kaolinite clay minerals using infrared emission spectroscopy. *Clays Clay Miner.* **44**, 635–651 (1996).
47. Libowitzky, E. & Beran, A. O. H. Defects in Forsterite. *Phys. Chem. Minerals* **22**, 387–392 (1995).
48. Bargara, J. R., Towle, S. N., Brown, G. E., Jr. & Parks, G. A. XAFS and bond-valence determination of the structures and compositions of surface functional groups and Pb(II) and Co(II) sorption products on single-crystal $\alpha\text{-Al}_2\text{O}_3$. *J. Colloid Interface Sci.* **185**, 473–492 (1997).
49. Catalano, J. G., Trainor, T. P., Eno, P. J., Waychunas, G. W. & Brown, G. E., Jr. CTR diffraction and grazing-incidence EXAFS study of U (VI) adsorption onto $\alpha\text{-Fe}_2\text{O}_3$ (1102) surfaces. *Geochim. Cosmochim. Acta* **69**, 3555–3572 (2005).
50. Morrow, B. A. *et al.* The Silica Surfaces. in Papirer, E. (Ed.). *Adsorption on Silica Surfaces*, CRC Press, Santa Barbara, pp. 9–276 (2000).
51. Rimola, A., Costa, D., Sodupe, M., Lambert, J. F. & Ugliengo, P. Silica surface features and their role in the adsorption of biomolecules: computational modeling and experiments. *Chem. Rev.* **113**, 4216–4313 (2013).
52. Ju, H. *et al.* Application of hydrophilic silanol-based chemical grout for strengthening damaged reinforced concrete flexural members. *Materials* **7**, 4823–4844 (2014).
53. Yildirim, H., Kinaci, A., Zhao, Z. J., Chan, M. K. Y. & Greeley, J. P. First-principles analysis of defect-mediated Li adsorption on graphene. *ACS Appl. Mater. Interfaces* **6**, 21141–21150 (2014).
54. Yazaydin, A. O. & Thompson, R. W. Molecular simulation of Water adsorption in silicalite: effect of silanol groups and different cations. *Microporous Mesoporous Mater.* **123**, 169–176 (2009).
55. Bish, D. L. Rietveld refinement of the kaolinite surface at 1.5 K. *Clays Clay Miner.* **41**, 738–744 (1993).
56. Aquino, A. J. A., Tunega, D., Haberhauer, G., Gerzabek, M. H. & Lischka, H. Adsorption of organic substances on broken clay surfaces: a quantum chemical study. *J. Comput. Chem.* **24**, 1853–1863 (2003).
57. Delorme, M., Labbez, C., Caillet, Celine. & Thomas, F. Acid-base properties of 2:1 clays. I. modeling the role of electrostatics. *Langmuir* **26**, 9240–9249 (2010).
58. van der Spoel, D. *et al.* Gromacs: fast, flexible, and free. *J. Comput. Chem.* **26**, 1701–1718 (2005).
59. Cygan, R. T., Liang, J. J. & Kalinichev, A. G. Molecular models of hydroxide, oxyhydroxide, and clay phases and the development of a general force field. *J. Phys. Chem. B* **108**, 1255–1266 (2004).
60. Mizan, T. I., Savage, P. E. & Ziff, R. M. Comparison of rigid and flexible simple point charge water models at supercritical conditions. *J. Comput. Chem.* **17**, 1757–1770 (1996).
61. Suter, J. L., Coveney, P. V., Greenwell, H. C. & Thyveetil, M. A. Large-scale molecular dynamics study of montmorillonite clay: emergence of undulatory fluctuations and determination of material properties. *J. Phys. Chem. C* **111**, 8248–8259 (2007).
62. Du, H. & Miller, J. D. Adsorption states of amphipathic solutes at the surface of naturally hydrophobic minerals: a molecular dynamics simulation study. *Langmuir* **23**, 11587–11596 (2007).
63. Mazo, M. A. *et al.* Molecular dynamics simulation of thermomechanical properties of montmorillonite crystal. 1. isolated clay nanoplate. *J. Phys. Chem. B* **112**, 2964–2969 (2008).
64. Cheng, T. & Sun, H. Adsorption of ethanol vapor on mica surface under different relative humidities: A molecular simulation study. *J. Phys. Chem. C* **116**, 16436–16446 (2012).
65. Shapley, T. V., Molinari, M., Zhu, R. & Parker, S. C. Atomistic modeling of the sorption free energy of dioxins at clay-water interfaces. *J. Phys. Chem. C* **117**, 24975–24984 (2013).
66. Bussi, G., Donadio, D. & Parrinello, M. Canonical sampling through velocity rescaling. *J Chem Phys.* **126**, 014101 (2007).
67. Hockney, R. W., Goel, S. P. & Eastwood, J. W. Quiet high-resolution computer models of a plasma. *J. Comput. Phys.* **14**, 148–158 (1974).
68. Nose, S. & Klein, M. L. Constant pressure molecular dynamics for molecular systems. *Mol. Phys.* **50**, 1055–1076 (1983).

Acknowledgement

We gratefully acknowledged the financial supports from the National Natural Science Foundation of China (41371249 and 21473137) and the Fourth Excellent Talents Program of Higher Education in Chongqing (2014–03).

Author Contributions

Conceived and designed the experiments: X.L., G.Y. and H.L. Performed the experiments: X.L. Wrote the paper: X.L. and G.Y.

Additional Information

Supplementary information accompanies this paper at <http://www.nature.com/srep>

Competing financial interests: The authors declare no competing financial interests.

How to cite this article: Li, X. *et al.* Promoting the Adsorption of Metal Ions on Kaolinite by Defect Sites: A Molecular Dynamics Study. *Sci. Rep.* **5**, 14377; doi: 10.1038/srep14377 (2015).



This work is licensed under a Creative Commons Attribution 4.0 International License. The images or other third party material in this article are included in the article's Creative Commons license, unless indicated otherwise in the credit line; if the material is not included under the Creative Commons license, users will need to obtain permission from the license holder to reproduce the material. To view a copy of this license, visit <http://creativecommons.org/licenses/by/4.0/>