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3D TiO₂@Ni(OH)₂ Core-shell Arrays with Tunable Nanostructure for Hybrid Supercapacitor Application

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Three dimensional hierarchical nanostructures have attracted great attention for electrochemical energy storage applications. In this work, self-supported TiO₂@Ni(OH)₂ core-shell nanowire arrays are prepared on carbon fiber paper via the combination of hydrothermal synthesis and chemical bath deposition. In this core-shell hybrid, the morphology and wall size of the interconnected nanoflake shell of Ni(OH)₂ can be tuned through adjusting the concentration of ammonia solution. Heterogeneous nucleation and subsequent oriented crystal growth are identified to be the synthesis mechanism affecting the nanostructure of the shell material, which consequently determines the electrochemical performance in both energy storage and charge transfer. Superior capabilities of 264 mAhg⁻¹ at 1A g⁻¹ and 178 mAh g⁻¹ at 10A g⁻¹ are achieved with the core-shell hybrids of the optimized structure. The asymmetric supercapacitor prototype, comprising of TiO₂@Ni(OH)₂ as the anode and mesoporous carbons (MCs) as the cathode, is shown to exhibit superior electrochemical performance with high energy and power densities. The present work provides a clear illustration of the structure-property relationship in nanocrystal synthesis and offers a potential strategy to enhance the battery type Ni(OH)₂ electrode in a hybrid supercapacitor device.

Three dimensional (3D) core-shell nanoarrays grown on substrates have attracted great scientific and technological attentions due to their high versatility and applicability as the essential components in nanoscale electronics, catalysis, chemical sensors, and energy conversion storage devices¹⁻⁴. In comparison to bulk materials, core-shell heterostructures not only provide a large interfacial area between electrode and electrolyte for charge transport and shortened diffusion path for intercalation/de-intercalation of active species, but also give rise to synergetic properties or multifunctionalities provided by individual components. Therefore, a novel constructions of core-shell structures are needed for fabrication of high-performance electrochemical energy storage devices. Currently, considerable efforts have been devoted to explore various core-shell structures, including metal/metal oxide, metal/metal, metal oxide/metal oxide and metal oxide/conductive polymer^{5,6}. Transition metal oxides and hydroxides (e.g., Co₃O₄, Co(OH)₂, MnO₂, Mn(OH)₂, NiO, Ni(OH)₂) acting as branched shell materials were generally integrated with conducting scaffold into core-shell structures (e.g., CoO@Co(OH)₂, CNT@Ni(OH)₂, SnO₂@MnO₂)⁷⁻¹⁰.

TiO₂ is an inexpensive and electrochemically stable semiconductor commonly used in lithium batteries¹¹. Moreover, the electrical conductivity of around 10⁻⁵-10⁻² S cm⁻¹ permits its application as backbone in the core-shell structure (e.g., TiO₂@MnO₂, TiO₂/NiO)^{12,13}. To enhance the storage capability of TiO₂-based hybrids, strategies have been focused on enhancing the conductivity of TiO₂¹⁴. For example, Lu *et al.* managed to increase the electrical conductivity of TiO₂ by three orders of magnitudes through hydrogenation¹¹. Alternatively, Liao *et al.* improved the electrical conductivity through coating carbon across the outer layer of TiO₂¹². These efforts to some extent have boosted the storage capacity,

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rate capability and cyclability in energy storage. However, few studies have been made in optimizing the shell structure, especially in controlling the shell configuration and engineering the interface, which are the key parameters in affecting the electrochemical performance.

There are an increasing number of studies regarding active electrode materials that undergo faradaic reactions but are used for electrochemical capacitor applications. Unfortunately, some of these materials are described as “pseudocapacitive” materials despite the fact that their electrochemical signature (e.g., cyclic voltammogram and charge/discharge curve) is analogous to that of a “battery” material, as commonly observed for Ni(OH)₂ in KOH electrolyte^{15,16}. Ni(OH)₂ has been the focus of research interest, owing to its high stability, low toxicity, low cost and high theoretical capacity, which make it useful for energy storage^{17–21}. For freestanding Ni(OH)₂ particles, the electrochemical performance was reported to highly correlate with several nanostructure parameters, including crystallite size, structural defects and surface area^{21,22}. Ni(OH)₂ particles in the form of hollow microsphere-like, nanoflower-like, and nanoplate-like configuration show strong structure-property relationship^{23–25}. Therefore, successful synthesis of Ni(OH)₂ with desired nanostructures is of paramount importance. In the present work, we have designed a new 3D structure of TiO₂@Ni(OH)₂ core-shell nanoarrays grown on carbon fiber paper (CFP). Interestingly, the nanostructure of Ni(OH)₂ branches can be greatly modified by adjusting the concentration of ammonia added in the solution mixture. Superior electrochemical performances of 264 mAh g⁻¹ at 1 A g⁻¹ and 178 mAh g⁻¹ at 10 A g⁻¹ are achieved with the core-shell hybrid of optimized structure. The structure-property relationships observed for the core-shell nanoarrays in this work offer a new strategy to enhance the performance of supercapacitor electrodes.

Results

The fabrication procedures of the core-shell nanowire electrodes are illustrated in Figure 1a. In the typical preparation route, the TiO₂ nanowires were prepared on a CFP in the form of branched nanowire arrays by hydrothermal method. In the following step, metal hydroxide nanoflakes of Ni(OH)₂ are grown onto the TiO₂ nanowires by a chemical bath deposition. During this process, different amounts of aqueous ammonia (24% NH₃·H₂O, 3ml, 4ml, 6ml) are added to the solution, and the corresponding samples are labeled as A-3/CFP, A-4/CFP, A-6/CFP, respectively (see the Experimental Section). In the 3D core-shell hierarchical nanostructure, the core material is directly aligned on the current collector. This rationalized configuration is to provide both large active surface areas and good electrical connection for fast redox kinetics, therefore being capable of boosting the efficiency of the energy storage⁷. In the present work, the surface of carbon fibers was largely covered by the TiO₂ nanowires, which have diameters of ~200 nm and grown nearly vertical to the carbon fiber forming nanowires arrays (Fig. 1b). After the chemical bath deposition process, the interconnected Ni(OH)₂ nanoflakes fully decorate the scaffold of TiO₂ nanowires, giving rise to a core-shell configuration with diameters of ~1000 nm (Fig. 1c). A single hybrid nanowire of TiO₂@Ni(OH)₂ was unambiguously examined by the energy dispersive X-ray spectrometry (EDS) mapping analysis (see Fig. 1d), which indicates that the inner nanowire core is TiO₂ and the outer-layer shell is Ni(OH)₂ in the hybrid core-shell structure. To demonstrate the effect of the amount of ammonia on the likely change of chemical states in Ni ion, XPS survey of Ni ion is collected as shown in Fig. 1e. All the Ni XPS surveyed in the three samples show two major peaks centered at 856.1 and 873.6 eV with a spin-energy separation of 17.5 eV, corresponding to Ni 2p_{3/2} and Ni 2p_{1/2}, respectively, together with extra three sets of peaks located at higher energy sides being correlated with the Ni 2p_{3/2} and Ni 2p_{1/2} satellite peaks²⁶. These bands are characteristic peaks of Ni(OH)₂ phase, which is consistent with previous reports²⁷.

The XRD patterns of the A-3/CFP, A-4/CFP, and A-6/CFP nanowire electrodes are shown in Fig. 2, where the diffraction pattern of the TiO₂/CFP is also collected as a reference for comparison. The diffraction peaks positioned at 2θ of 27.9°, 36.6°, 40.01°, 41.7°, 44.4° assigning to (110), (101), (200), (111) and (210) respectively, confirm the rutile phase of the TiO₂ nanowires. In addition to the diffraction peaks from TiO₂, peaks of low intensity arising from Ni(OH)₂ are also collected, as a result of the poor degree of crystallization. Moreover, one notes that compared with the diffraction pattern in A-3/CFP and A-4/CFP, two small humps located at 33.7° and 38.4° were observed for A-6/CFP, which were assigned to (100) and (101) of β-Ni(OH)₂, respectively (JCPDS card no. 1-1047)²⁸. Therefore, the addition of more ammonia tends to increase the crystallinity of Ni(OH)₂ in the chemical bath deposition process, although concentrations of other salts in the solution are also changed.

Table 1 summarizes the BET surface area and BJH desorption cumulative pore volume results for TiO₂/CFP, A-3/CFP, A-4/CFP and A-6/CFP electrodes. The BJH desorption cumulative pore volume is calculated on the basis of pores sizes from 1.7 to 300 nm²⁸. For the TiO₂/CFP, the BET surface area is 2.56 m² g⁻¹. After chemical deposition of Ni(OH)₂ on the TiO₂/CFP, the surface area and pore volume are tremendously increased from 2.56 to 22.68 m² g⁻¹, together with an enhancement in BJH pore volume from 0.000371 to 0.048436 cm³ g⁻¹. This improvement in surface area and pore volume can be ascribed to flaky nanowalls of Ni(OH)₂, which could benefit to the fast charge/discharge kinetic of Ni(OH)₂ electrode^{7,21}. Moreover, we note that the enhanced surface area and pore volume occur in the sample prepared from the smaller amount of ammonia added. For example, the sample of A-3/CFP shows a relatively higher surface area of 22.68 m² g⁻¹ together with a higher BJH deposition cumulative pore volume of 0.048436 cm³ g⁻¹, as compared with those of A-4/CFP and A-6/CFP.

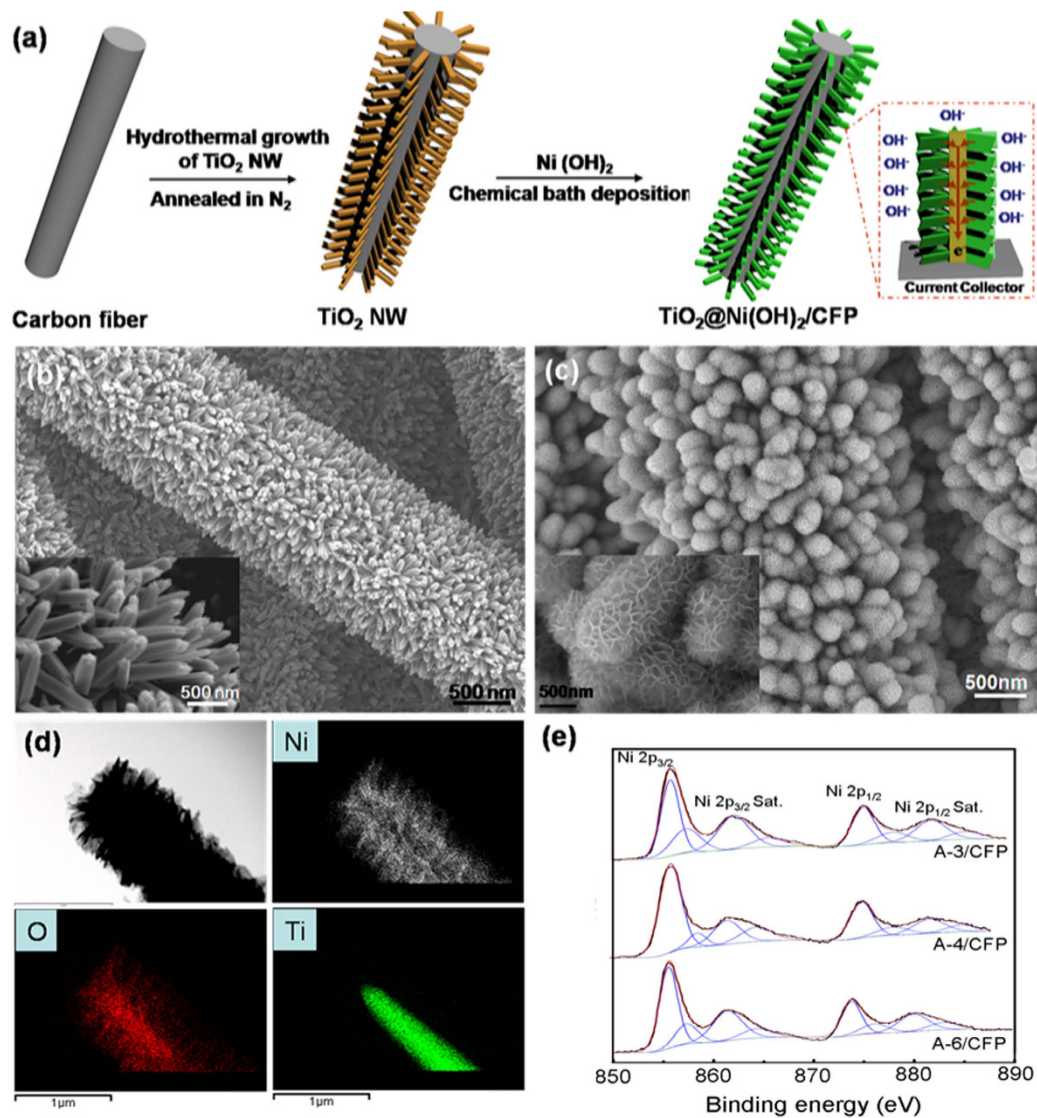


Figure 1. (a) Schematic illustration for the fabrication of $\text{TiO}_2@Ni(OH)_2$ core-shell nanowire arrays on CFP. (b) SEM image of the TiO_2 nanowire arrays, and (c) $\text{TiO}_2@Ni(OH)_2$ core/shell nanowire arrays on the CFP. The fine structures are shown in insets. (d) EDS mapping results from a single hybrid nanowire, demonstrating the TiO_2 core and $Ni(OH)_2$ shell hierarchical structure. (e) XPS spectroscopy of Ni 2p in samples of A-3/CFP, A-4/CFP and A-6/CFP.

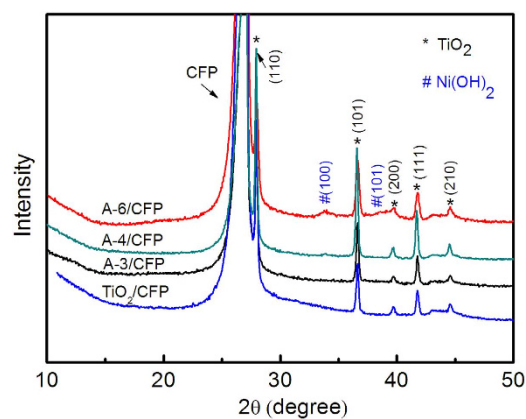


Figure 2. XRD patterns of the A-3/CFP, A-4/CFP, A-6/CFP and TiO_2 /CFP nanowire electrodes.

	BET surface area ($\text{m}^2 \text{g}^{-1}$)	BJH absorption cumulative pore volume ($\text{cm}^3 \text{g}^{-1}$)
TiO ₂ /CFP	2.56	0.000371
A-3/CFP	22.68	0.048436
A-4/CFP	19.58	0.030831
A-6/CFP	15.87	0.028379

Table 1. BET surface area and BJH absorption cumulative pore volume of TiO₂/CFP, A-3 /CFP, A-4 /CFP, and A-6 /CFP.

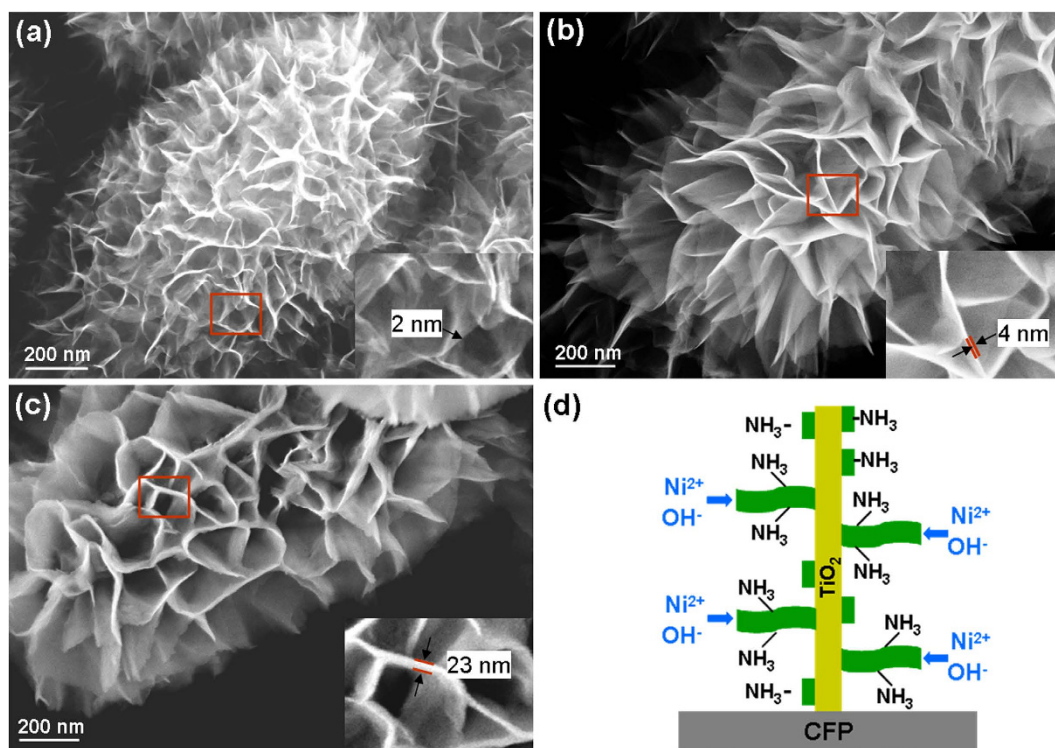


Figure 3. SEM images of (a) A-3/CFP, (b) A-4/CFP, (c) A-6/CFP and the insets show the high-magnified SEM images of Ni(OH)₂. (d) Schematic diagram showing the Ni(OH)₂ nanowalls growth process.

Figure 3 shows the SEM images of A-3/CFP, A-4/CFP and A-6/CFP, respectively, where the morphologies of Ni(OH)₂ nanoflakes are different. For the sample prepared from a smaller amount of ammonia added (i.e., A-3/CFP), gauzy Ni(OH)₂ sheets with curved surfaces are observed and the wall thickness of the nanosheet in A-3/CFP is estimated to be ~2 nm. With the increasing amount of ammonia added, the Ni(OH)₂ nanosheets become flat and the wall thickness increases to ~23 nm in A-6/CFP. The core-shell structures and phases of TiO₂@Ni(OH)₂ were further elucidated by the transmission electron microscopy (TEM) and selected area electron diffraction (SAED) analysis. Figure 4a shows a typical TEM image of an individual TiO₂ nanowire and the inset shows the SAED pattern recorded from this particle. The bright and sharp diffraction spots of rutile TiO₂ (JCPDF 65-0192) are clarified, an indication of the single-crystalline nature of the TiO₂ nanowire with (001) orientation²⁹. A detailed morphology and structure of nanoflakes from the shell material of Ni(OH)₂ are shown in Fig. 4b–d. In consistence with the SEM studies, nanosheets in A-3/CFP exhibit a curved thin flaky appearance, while the sample prepared from the higher concentration of ammonia show thick and flat nanosheet layers of Ni(OH)₂. The SAED patterns of the three samples are almost identified, as shown in the insets, where similar ring images are observed. These rings are indicative of a polycrystalline nature, and they correspond to the (111) and (220) planes of Ni(OH)₂, respectively (JCPDS 47-1049).

Discussion

In order to understand the growth mechanism, the chemical reaction process of Ni(OH)₂ in the solution was systemically investigated. The general formation mechanism for sphere-shaped hydroxide particles in the presence of ammonia has been discussed previously. It has been suggested that the metal ions

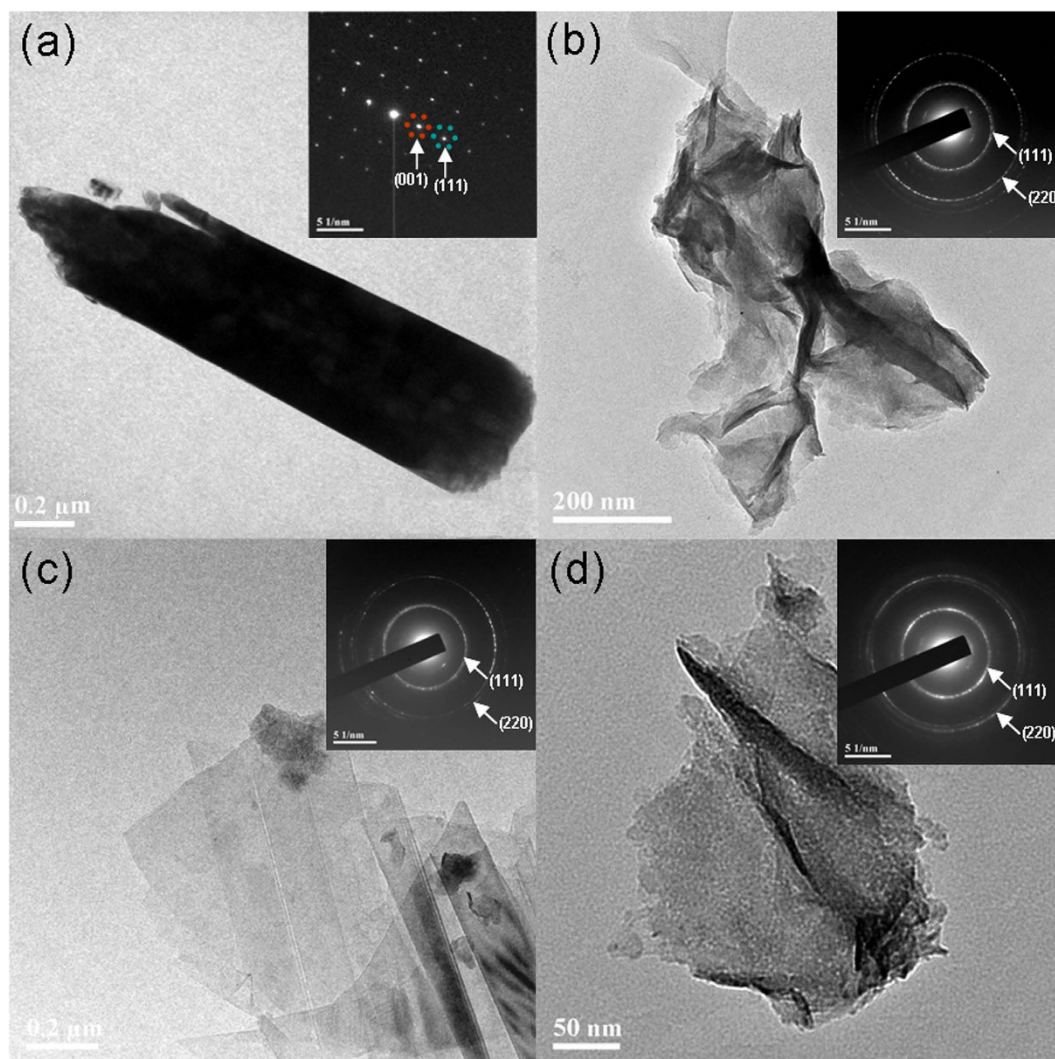
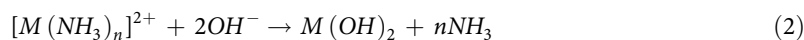
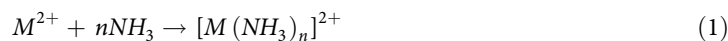
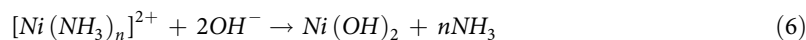


Figure 4. TEM images and SAED patterns of (a) TiO₂ nanowire, (b) Ni(OH)₂ flakes, A-3/CFP, (c) A-4/CFP and (d) A-6/CFP.

coordinate to the ammonia first to form a starting salt, which subsequently slowly releases to the basic solution to yield hydroxide particles^{30–32}. The detailed process can be briefly described as follows:



In our case of the nanostructured Ni(OH)₂, the hydrolysis of nickel-ammonia complex of [Ni(NH₃)₆]²⁺ has been proposed²³. Together, ammonia and NiSO₄ form [Ni(NH₃)₆]²⁺ and then hydrolyze into Ni(OH)₂ subsequently. These steps are controlled by the ammonia added into the solution and the detailed chemical reactions are given below:

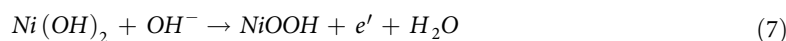


From the reaction processes above, there are dual roles played by ammonia. One is that the OH^- released from the ammonia will react with Ni^{2+} to form $\text{Ni}(\text{OH})_2$. On the other hand, NH_3 in ammonia will consume part of Ni^{2+} to generate $[\text{Ni}(\text{NH}_3)_6]^{2+}$, which will be hydrolyzed into $\text{Ni}(\text{OH})_2$ with the participation of extra OH^- ²³.

Two different formation mechanisms of $\text{Ni}(\text{OH})_2$ sheets have been proposed: “oriented crystal growth” and “self-assembly”. In the latter, it was proposed that the adjacent Ni-based hydroxide particles tend to self-organize at a planar interface to show a common crystallographic orientation³³. In the “oriented crystal growth” theory, it argues that the oriented growth of $\text{Ni}(\text{OH})_2$ originates from the NH_3 molecule selectively adsorbs on certain crystal phases of $\text{Ni}(\text{OH})_2$, which therefore suppress the growth of $\text{Ni}(\text{OH})_2$ along this particular direction. However, we noted that both hypotheses share a common understanding that the formation of the hydroxides starts from heterogeneous nucleation process. In our case, the formation of $\text{Ni}(\text{OH})_2$ can be explained by heterogeneous nucleation and subsequent oriented crystal growth. The TiO_2 nanowires acting as backbones provide sites for the preferential deposition of Ni-based hydroxides, and the growth process of the shell structure is proposed and shown schematically in Fig. 3d. At the beginning of the growth process, Ni^{2+} ions first react with OH^- released from the ammonia to form $\text{Ni}(\text{OH})_2$ nanocrystal seeds, which anchor to the surface of the TiO_2 nanowire to reduce the surface energy. With the increase in reaction time, 2D $\text{Ni}(\text{OH})_2$ is formed due to the NH_3 molecules that selectively adsorb on the surface of $\text{Ni}(\text{OH})_2$ nanosheets by hydrogen bonds, which exist between the nitrogen atom and the hydrogen atom of the surface hydroxyl groups (N-H-O)²³. The NH_3 then suppresses the growth of $\text{Ni}(\text{OH})_2$ along out-of-surface direction, but enables in-plane sheets growth. It is also mentioned that $\text{K}_2\text{S}_2\text{O}_8$ plays a critical role in the formation of nanoflake shell, where it functions as an oxidant that drive the whole reaction and facilitate the heterogeneous nucleation in the chemical solution. Without the participation of $\text{K}_2\text{S}_2\text{O}_8$, the nanoflake shell can not be generated on TiO_2 scaffold³³.

Based on the above understanding of the growth process for $\text{Ni}(\text{OH})_2$ nanosheets, the modification of wall thickness of $\text{Ni}(\text{OH})_2$ by ammonia can be well explained. NH_3 molecules adsorbed on the surface of $\text{Ni}(\text{OH})_2$ nanosheets will react with Ni^{2+} in the solution to form $[\text{Ni}(\text{NH}_3)_6]^{2+}$, which can be further translated into $\text{Ni}(\text{OH})_2$ upon reaction with OH^- as illustrated in Equ. (6). An increase in the concentration of OH^- accompanied with more ammonia added helps the hydrolysis of $[\text{Ni}(\text{NH}_3)_6]^{2+}$ into $\text{Ni}(\text{OH})_2$ effectively, and therefore resulting in an increase in the nanowall thickness. In the present work, when the amount of ammonia is increased from 3 to 6ml, the wall thickness of $\text{Ni}(\text{OH})_2$ is increased from 2 to 23 nm.

The electrochemical behavior of the three $\text{TiO}_2@/\text{Ni}(\text{OH})_2$ core-shell nanostructures are evaluated in a three-electrode measurement system by using the active material as the working electrode, the platinum foil as the counter electrode, and Hg/HgO electrode as the reference electrode. As shown in Fig. 5a, the cyclic voltammogram (CV) curves are recorded at a scan rate of 5 mV s^{-1} with a potential window ranging from -0.2 to 0.6 V. One notes that the area enclosed by the CV curve is decreased for the sample prepared from the larger amount of ammonia added. For instance, The CV curve of A-3/CFP (black line) generates an area larger than that of the A-6/CFP (green line), demonstrating an enhanced charge storage capability. Moreover, one set of sharp redox peaks is observed, which reveals that the capacity characteristics are governed mainly by the faradaic redox reaction, rather than electric double layer capacitors (EDLC) with CV curves in a rectangular shape³⁴. The electrochemical charge transfers in $\text{Ni}(\text{OH})_2$ are along the chain of Ni (II)-Ni(III) in the basic electrolyte and the reaction can be described as follows³⁵



The charging-discharging curves of the three $\text{TiO}_2@/\text{Ni}(\text{OH})_2$ core-shell nanostructures are shown in Fig. 5b, at a current density of 1 A g^{-1} . It indicates that the charging-discharging steps become more obvious for the sample prepared from the smaller amount of ammonia added (e.g. A-3/CFP) and the corresponding discharging time is dramatically extended. This demonstrates an improved electrochemical performance of battery type $\text{Ni}(\text{OH})_2$ electrode in a hybrid supercapacitor device through decreasing the concentration of ammonia.

The variation of capacity at different current densities is shown in Fig. 5c, where a nearly stable capacity value at high current densities is clarified. It indicates a desirable retention capability, resulting from the 3D nanostructure supported by the TiO_2 scaffold, which provides plenty of conductive tunnels for an efficient charge transfer³⁶. The electrochemical performance is improved for the $\text{TiO}_2@/\text{Ni}(\text{OH})_2$ prepared from the lower concentration of ammonia. As shown in Fig. 5c, the core-shell nanowire arrays of A-3/CFP exhibit a specific capacity of 264 mAh g^{-1} at 1 A g^{-1} and 178 mAh g^{-1} at 10 A g^{-1} . In contrast, with the increasing amount of ammonia added, the sample shows a lower storage capability (e.g., the A-6/CFP delivers a much lower capacity of 57 mAh g^{-1} at 1 A g^{-1} and 5 mAh g^{-1} at 10 A g^{-1}). Moreover, the electrochemical performance demonstrated for the core-shell arrays of $\text{TiO}_2@/\text{Ni}(\text{OH})_2$ in this work is much better than those of several other nanostructured NiO films³⁷⁻³⁹, TiO_2 -NiO counterpart⁴⁰ and NiO powders⁴¹⁻⁴². The observed structure-property relationship can well be ascribed to the thinner nanowall and higher surface area, which provides more sites for the redox reaction, for the $\text{Ni}(\text{OH})_2$ derived from the lower ammonia concentration^{43,44}.

EIS is measured and the corresponding Nyquist plots are shown in Fig. 5d. The electrochemical impedance measurement was biased to 0 V (vs. the open-circuit voltage) using the three-electrode

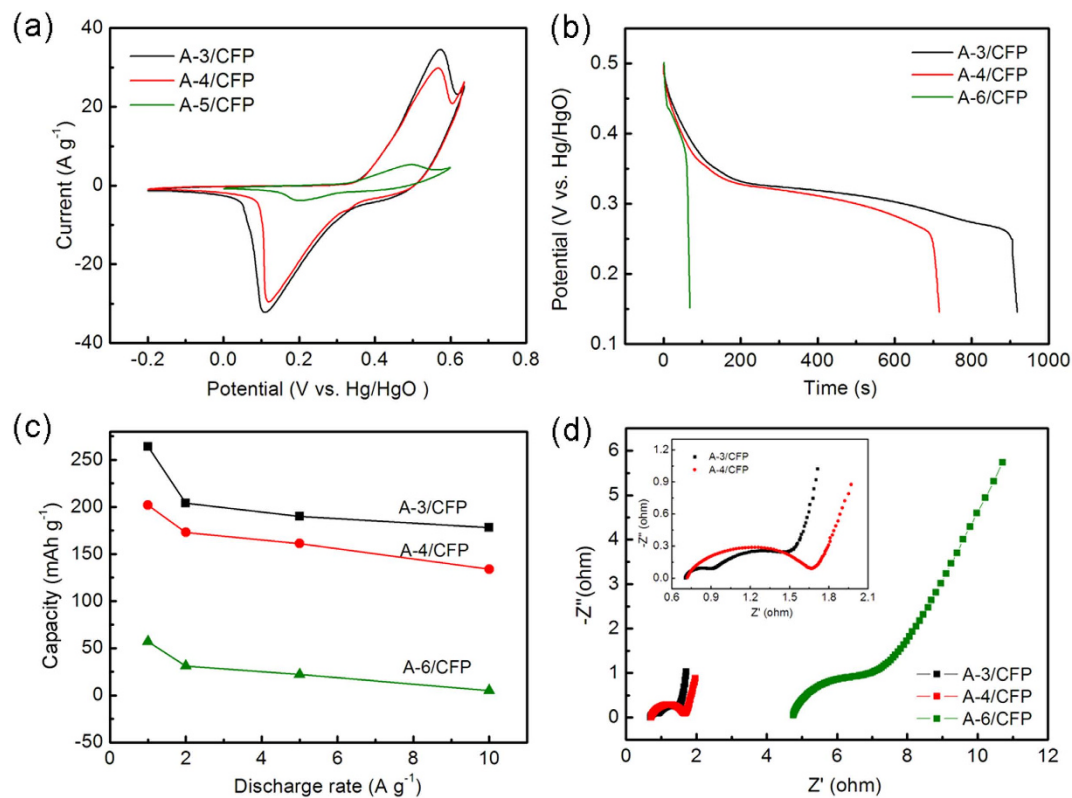


Figure 5. Comparisons of the electrochemical performance of the three samples: (a) CV curves ranging from -0.2 to 0.6 V at a scan rate of 5 mV s⁻¹, (b) charging-discharging curves at 1 A g⁻¹, (c) capacity at different current densities, (d) impedance Nyquist plots between 0.001 and 100 kHz at 5 mV applied voltage.

system, at a frequency ranging from 0.01 to 100 kHz with 5 mV RMS voltage perturbation amplitude. The impedance spectra of the electrodes made of $\text{TiO}_2/\text{Ni}(\text{OH})_2$ nanowire arrays exhibit similar semicircles in the high frequency region and straight line in the low frequency region. It is noted that the spike-like region is more close to the imaginary axis for the sample prepared from the smaller amount of ammonia added (A-3/CFP), demonstrating fast charge-transfer kinetics and electric responses resembling a circuit with low resistance and large capacitance connected in parallel⁴⁵. In contrast, in the impedance curve of A-6/CFP, there is indication of the mass diffusion obstacles for ion or charge transfer in the electrode. In addition, it is also noted that the serial resistance (intercept on real axis) and charge transfer resistance (diameter of the semicircle) of A-3/CFP are lower than those of the A-4/CFP, followed by those of A-6/CFP. The serial resistance is ascribed to the inter-granular electronic resistance between active material particles, and the contact resistance between the active material and current collector, while the charge transfer resistance originates at the electrode/electrolyte interface^{46,47}. The EIS data clearly demonstrates that the $\text{TiO}_2/\text{Ni}(\text{OH})_2$ prepared from the smaller amount of ammonia added to the solution gives rise to improved conductivity in both serial and charge transfer parts. It therefore leads to enhanced electron transport efficiency and contributes greatly to the observed improvement in capacity of A-3/CFP.

To evaluate the capacitive performance for a full-cell device, the A-3/CFP electrode-based asymmetric supercapacitor (ASC) prototype is chosen^{8,48}. As shown in Fig. 6a, $\text{TiO}_2/\text{Ni}(\text{OH})_2/\text{CFP}$ was chosen as the anode electrode and the mesoporous carbon (MCs)-based electrode was chosen as the cathode. For convenience, the asymmetric supercapacitor is labeled as A-3//MCs, in which the mass ratio of the anode to cathode active materials is around 1:5. CV curves were measured at various scan rates with voltage windows ranging from 0.2 to 1.8 V. It shows a rectangular-like shape in CV curves, indicating a similar capacitive behavior as those of EDLCs and RuO_2 . The specific capacitances calculated as a function of scan rates are summarized in Fig. 5c. The maximum specific capacitance is 181 F g⁻¹ at 5 mV s⁻¹, which is substantially higher than those recently reported for the $\text{Co}_3\text{O}_4/\text{Ni}(\text{OH})_2/\text{graphene}$ (110 F g⁻¹)²¹. Moreover, the ASC device demonstrates a good rate retention, with 70% of capacitance retained upon increasing the scan rate from 5 to 100 mV s⁻¹.

To further evaluate the device performance, galvanostatic discharge measurements at various charge-discharge current densities were conducted. As shown in Fig. 6d, there is a nearly linear behavior in the discharge of A-3//MCs, which is an indication of capacitor-like behavior²¹. Both power density and energy density are important parameters for the performance of supercapacitor devices⁴⁹. Ragone plot calculated from the charging-discharging curves is shown in Fig. 6e. At the low current density of 0.1 A

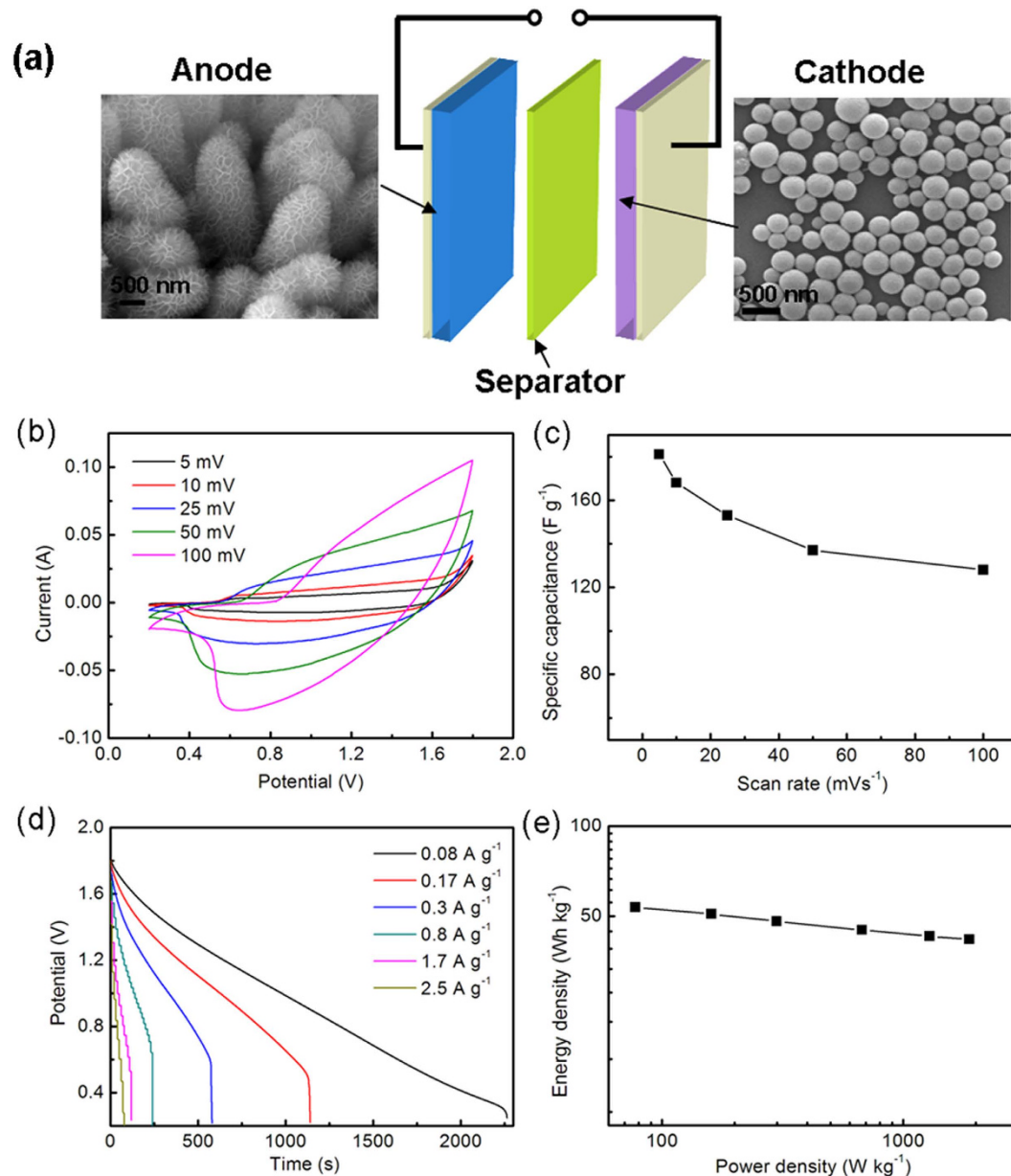


Figure 6. (a) Schematic illustration of the A-3//MCs asymmetric supercapacitor prototype with 6M KOH as the electrolyte. (b) CV curves of A-3//MCs at different scan rates. (c) Specific capacitances at different scan rates. (d) Discharge curves of A-3//MCs asymmetric device at different current densities. (e) Ragone plot of the A-3//MCs asymmetric device.

g⁻¹, the energy density and power density are 53.54 Wh kg⁻¹ and 77.71 W kg⁻¹, respectively, for A-3//MCs. At the high discharge current density of 2.5 A g⁻¹, the energy density remains at 41.78 Wh kg⁻¹ at the power density of 1881.87 W kg⁻¹. The superior energy density measured for the TiO₂@Ni(OH)₂/CFP-based asymmetric device is in part attributed to the high specific capacitance of the electrode material, which can be optimized by modulating nanostructure of Ni(OH)₂ nanoflakes.

Conclusions

3D TiO₂@Ni(OH)₂ core-shell nanostructures are successfully made by the combination of hydrothermal reaction and chemical bath deposition. By tuning the amount of ammonia added into the solution, the density and morphology of the Ni(OH)₂ nanosheets can be well controlled. Flaky Ni(OH)₂ with thin nanowall and high surface area is developed when a smaller amount of ammonia is added. The growth process of Ni(OH)₂ is identified to involve heterogeneous nucleation and subsequent oriented crystal growth. There is a clear relationship between the nanostructural feature and electrochemical performance for the 3D TiO₂@Ni(OH)₂ core-shell arrays. Both the energy storage capability and conductance

are improved in the sample with a larger surface area and smaller nanowall thickness. Superior capacity values of 264 mAh g⁻¹ at 1 A g⁻¹ and 178 mAh g⁻¹ at 10 A g⁻¹ are measured, when the interconnected shell structure of Ni(OH)₂ is optimized, where there are fast ion/electron transfer and sufficient contact between the active material and electrolyte. An asymmetric prototype device is developed by employing TiO₂@Ni(OH)₂ as the anode, MCs as the cathode and 6 M KOH as the electrolyte. The asymmetric device shows superior capacitance behavior, including both high power density and energy density, demonstrating the potential for high performance supercapacitors.

Methods

Synthesis of self-supported TiO₂/Ni(OH)₂ core/shell nanowires. In our work, the self-supported TiO₂ nanowire arrays on carbon fiber paper (CFP) were prepared by modifying a previously reported hydrothermal method¹⁴. In the typical process, the CFP was firstly cleaned with diluted HNO₃ and deionized water for 5 min each, respectively. The Ti precursor was prepared by adding 1 ml of titanium butoxide to a well-mixed solution containing 15 mL of HCl and 15 mL of H₂O. The mixture was then stirred for 24 h to obtain a clear solution. Afterward, a piece of clean CFP (2 × 2 cm²) was immersed into the above aqueous solution. Next, it was sealed into an autoclave and heated at 150 °C for 9 h, and then allowed to cool to room temperature spontaneously. After the reaction, the coated substrate was collected from the solution and then ultrasonicated in deionized water for 1 min to remove free nanoparticles and residual reactant. Finally, the nanowire coated CFP was calcinated at 550 °C in N₂ for 2 h, completing the growth of TiO₂ nanowire arrays on substrate. To grow Ni(OH)₂ shells on the TiO₂ nanowire cores, a solution was prepared by mixing 16 mL of 1M NiSO₄·6H₂O, 20 mL of 0.25 M K₂S₂O₄ and different amounts of aqueous ammonia (24% NH₃·H₂O, 3 ml, 4 ml, 6 ml) in a Pyrex beaker at room temperature¹⁹. The TiO₂-coated CFPs were immersed into the aqueous solution. When the room temperature chemical bath deposition of Ni(OH)₂ was performed for 10 min, the Ni(OH)₂-coated TiO₂/CFP (TiO₂@Ni(OH)₂/CFP) sample was subjected to a high speed rotation at 500 rpm for another 5 min. The resultant samples were dried at 60 °C for 24 h.

Material Characterization. The morphology and structure of each sample at every stages of preparation were characterized using scanning electron microscopy (SEM, Zeiss) and transmission electron microscopy (TEM, JEOL 2010) equipped with an energy-dispersive X-ray (EDS) analyzer. The specific surface area and pore volume were measured by N₂ adsorption-desorption using the Brunauer-Emmett-Teller (BET, Micromeritics ASAP2020) method. X-ray diffraction (XRD) patterns of the samples were obtained with Bruker AXS (D8 Advance, Cu Kα, λ = 0.154060 nm) at 40 kV and 40 mA. The sample composition and structure were studied using a Kratos Axis Ultra X-ray photoelectron spectrograph (XPS, Kratos Analytical) equipped with a monochromatized Al Kα X-ray source, where the chamber pressure was controlled at 10⁹ Torr.

Electrochemical Measurement. Electrochemical measurements of each electrode were conducted in three-electrode configuration with the as-prepared electrode as the working electrode, a platinum plate as the counter electrode, and an Hg/HgO as the reference electrode, in 6 M KOH electrolyte. The weight of each sample was measured using a Mettler Toledo X205DU microbalance (sensitivity, 0.01 mg; repeatability, 0.015 mg). Before weighing, all samples were dried in an oven at 120 °C for at least 1 day. The mass loading of the active material was calculated by m₂-m₁, where m₁ is the mass of TiO₂/CFP and m₂ is the mass of the TiO₂@Ni(OH)₂/CFP. The loading of Ni(OH)₂ is kept at around 5.2 mg cm⁻². Electrochemical performance was evaluated by cyclic voltammetry (CV), galvanostatic charge-discharge and impedance spectroscopy by using Solartron System 1470E and 1400A, respectively. The electrochemical impedance measurement was biased to 0 V (vs. the open-circuit voltage) using the three-electrode system, at a frequency ranging from 0.001 to 100 kHz with a 5 mV RMS voltage perturbation amplitude. The device capacitance was calculated from the charging-discharging curve by the following equation: $C = \frac{I\Delta t}{m\Delta V}$, where C (F g⁻¹) is the capacity, I (A) is the current, ΔV (V) is the potential window, m (g) is the mass of the active material and the Δt (s) is the discharging time.

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Author Contributions

Q.K. and J.W. conceived and design the experiment. Q.K. and M.Z. fabricated the samples. H.L. contributed XRD scanning and analysis. Q.K. conducted all characterizations. C.G. and L.M. conducted TEM studies. Q.K. and J.W. drafted the manuscript. All authors provided revision and contribution to the manuscript.

Additional Information

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