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# Unforeseen high temperature and humidity stability of FeCl<sub>3</sub> intercalated few layer graphene

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We present the first systematic study of the stability of the structure and electrical properties of FeCl<sub>3</sub> intercalated few-layer graphene to high levels of humidity and high temperature. Complementary experimental techniques such as electrical transport, high resolution transmission electron microscopy and Raman spectroscopy conclusively demonstrate the unforeseen stability of this transparent conductor to a relative humidity up to 100% at room temperature for 25 days, to a temperature up to 150 °C in atmosphere and to a temperature as high as 620 °C in vacuum, that is more than twice higher than the temperature at which the intercalation is conducted. The stability of FeCl<sub>3</sub> intercalated few-layer graphene together with its unique values of low square resistance and high optical transparency, makes this material an attractive transparent conductor in future flexible electronic applications.

Transparent conductive electrodes are extensively used for optoelectronic applications such as solar cells<sup>1,2</sup>, light emitting devices (LED) and displays<sup>3-5</sup>. To date, indium tin oxide (ITO) is the industrial standard for transparent electrodes with a high transmittance (T) of 85% in the visible wavelength range and a square resistance ( $R_{sq}$ ) as low as 10  $\Omega$ /sq for a film thickness of 10  $\mu\text{m}$ <sup>6,7</sup>. However this material has a number of limiting properties which are holding back the development of conceptually new types of screen technologies. For example, ITO strongly absorbs the infrared radiation, making it an interesting material largely for applications in the visible range of the light spectrum. Furthermore, ITO is relatively brittle<sup>6-9</sup> accounting for the rigidity of nowadays screens. Hence new materials are needed to let the screen technology emerge from the confines of walls and become foldable and wearable.

The chemical functionalization of graphene with atoms or molecules provides a valuable way to engineer the desired combination of physical properties needed to enable the aforementioned novel technologies<sup>10-13</sup>. For example, FeCl<sub>3</sub> intercalated few-layer graphene has a square resistance lower than ITO (just  $\approx$  8  $\Omega$ /sq) with an optical transmittance of 92% measured in bilayers from ultra-violet up to the near infra-red wavelength range<sup>12</sup>. This functionalization is easily scalable to large area substrates using graphene obtained with different methods such as epitaxial growth on silicon carbide (4H-SiC)<sup>14</sup>. The unique combination of low square resistance, high optical transmittance together with the exceptional mechanical flexibility of graphene makes this hybrid material a valuable substitute of ITO as a transparent electrode. However, it is well known that FeCl<sub>3</sub>, a commonly used etchant for copper, is highly hygroscopic and soluble in water. Intuition might suggest that FeCl<sub>3</sub> functionalized graphene would not be stable in air, with the humidity present in air rapidly diminishing the integrity and properties of the material.

Here we demonstrate that FeCl<sub>3</sub> intercalated few-layer graphene (FeCl<sub>3</sub>-FLG) is an unexpected highly stable form of functionalized graphene. This conclusion is based on a systematic comparative study of the evolution of the Raman spectra, high resolution transmission electron microscopy (HRTEM), and electrical transport characteristics upon exposing FeCl<sub>3</sub>-FLG to high relative humidity ( $H > 90\%$  and up to 100%) at room temperature, to a temperature up to 150 °C in atmosphere and up to a temperature as high as 620 °C in vacuum for HRTEM. After three weeks continual exposure to a relative humidity ranging from 95 to 100% at room temperature, we observe that the square resistance and the Raman spectra are unchanged. HRTEM measurements up to a temperature of 620 °C in vacuum, that is more than twice higher than the temperature at which intercalation

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is conducted, show no measurable change in the structure of the material. The stability of FeCl<sub>3</sub>-FLG to high humidity and temperatures widens considerably the range of potential applications targeted by graphene materials.

## Results

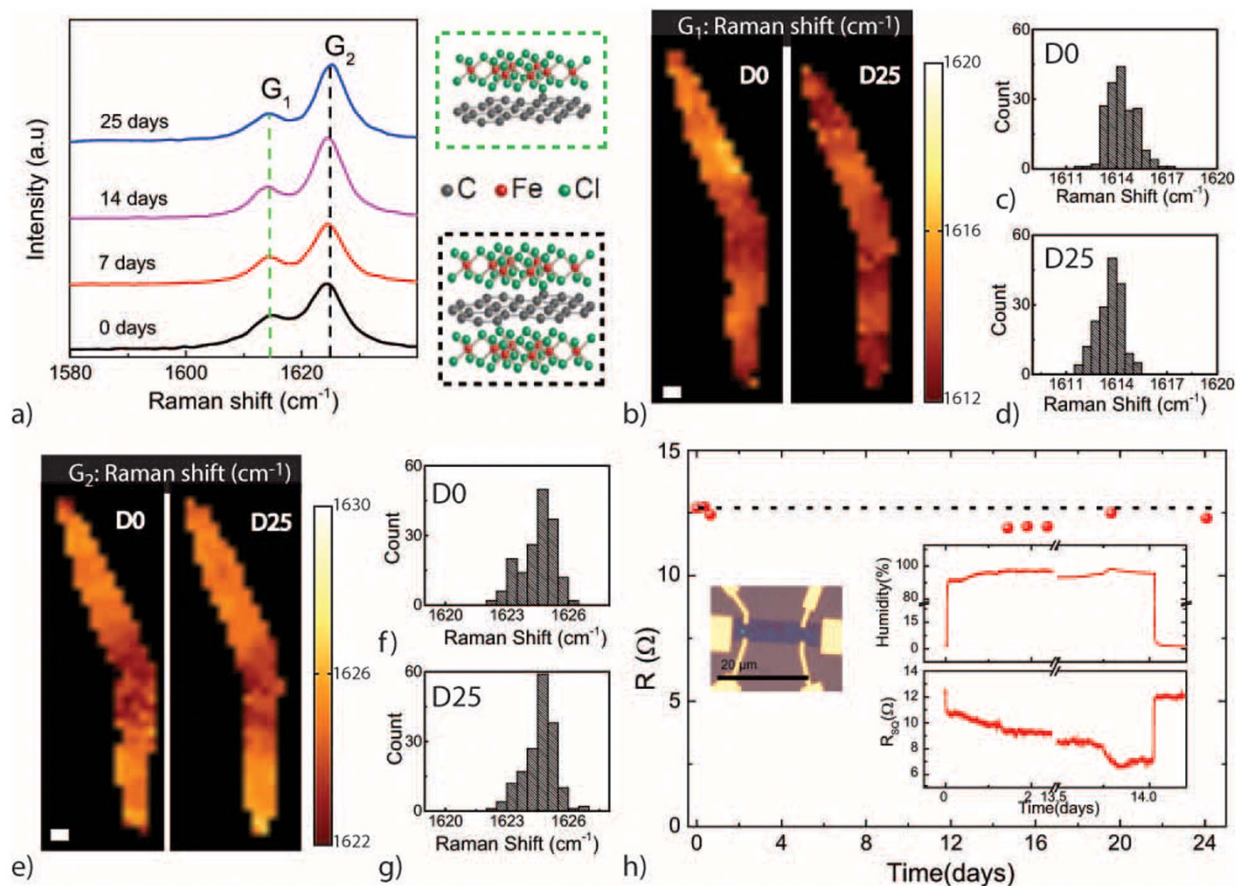
Few layer graphene are prepared by micromechanical cleavage of natural graphite on standard Si/SiO<sub>2</sub> substrates and transmission electron microscopy (TEM) grids. The intercalation with FeCl<sub>3</sub> is conducted using an established vapour transport method in a two-zone furnace<sup>12,15,16</sup>. The stage of intercalation represented by the number of carbon layers separating two subsequent intercalated layers, is determined using Raman spectroscopy<sup>12</sup>. The electrical properties are characterized using multi-terminal devices fabricated with standard electron-beam lithography, deposition of Cr/Au (10/50 nm) followed by lift-off, see inset in Figure 1a. The electrical resistance is measured (1) *in situ* while controlling the humidity in a closed chamber and (2) in air after heating the samples using a hotplate.

Figure 1a shows the G-peaks of the Raman spectrum at the same location of a representative FeCl<sub>3</sub>-FLG after exposing the sample to an atmosphere with relative humidity  $H > 95\%$  for various days as indicated in the graph. In contrast to the case of pristine graphene for which a single G-peak<sup>16</sup> is measured at 1585 cm<sup>-1</sup> (G<sub>0</sub>), in the intercalated material this peak shifts to 1615 cm<sup>-1</sup> (G<sub>1</sub>) and 1625 cm<sup>-1</sup>

(G<sub>2</sub>) can be observed. Both G<sub>1</sub> and G<sub>2</sub> are the consequence of stiffening of the E<sub>2g</sub> phonon mode<sup>17,18</sup> caused by charge transfer from FeCl<sub>3</sub> to graphene. More specifically, G<sub>1</sub> is characteristic of a graphene layer being doped by only one adjacent FeCl<sub>3</sub> layer whereas G<sub>2</sub> corresponds to a graphene layer sandwiched between two layers of FeCl<sub>3</sub> (see crystal structure illustrations in Figure 1a). No measurable shift of the G<sub>1</sub>- and G<sub>2</sub>-peaks is observed even after exposing the flake for 25 days to high humidity, suggesting that the intercalated compound is indeed not affected by the humidity.

To demonstrate that the structure of FeCl<sub>3</sub>-FLG is stable against prolonged exposure to extremely high levels of humidity, we conduct a detailed study of the Raman maps of the G<sub>1</sub>- and G<sub>2</sub>-peaks before and after exposing a representative flake to  $H > 95\%$  for 25 days, see Figure 1b–g. We find that overall the position of the aforementioned Raman peaks does not change significantly. Indeed, the distribution of the Raman shifts of G<sub>1</sub> (see Figure 1c–d) and G<sub>2</sub> (see Figure 1f–g) before and after exposure to humidity only exhibit at most a shift of  $\approx 1$  cm<sup>-1</sup> that is within the accuracy of the Raman spectroscopy tool used for this experiment.

Having established that the structure of this intercalated compound does not change upon exposure to humid atmosphere, we proceed to characterize the stability of the electrical properties under the same experimental conditions. Figure 1h shows a plot of the measured R<sub>sq</sub> in dry atmosphere after subsequent exposures to high



**Figure 1** | (a) Shows a plot of Raman spectra measured in a representative FeCl<sub>3</sub>-FLG sample before (0 days) and after exposure to  $H > 95\%$  for 7, 14 and 25 days shifted for clarity along the y-axis. The peaks G<sub>1</sub> and G<sub>2</sub> are highlighted on the graph, and the corresponding crystal structure is shown in the illustrations on the right side of the graph. Panel (b) shows the colour coded Raman maps of the G<sub>1</sub> peak before (D0) and after (D25) exposure to  $H > 95\%$ . The white scale bar corresponds to 2  $\mu\text{m}$ . (c) and (d) are the corresponding histograms of the Raman shift of G<sub>1</sub> for D0 and D25 respectively. Graphs in (e) are colour coded Raman maps of G<sub>2</sub> before (D0) and after (D25) exposure to high levels of humidity. The white scale bar corresponds to 2  $\mu\text{m}$ . (f) and (g) are graphs of the histograms of the Raman shift of G<sub>2</sub> for D0 and D25 respectively. The main graph in (h) is a plot of the value of R<sub>sq</sub> in dry atmosphere after exposing the device shown in the micrograph picture in the inset to high levels of humidity for different intervals of time. The graphs in the inset show the *in situ* R<sub>sq</sub> vs. time and the relative levels of humidity.



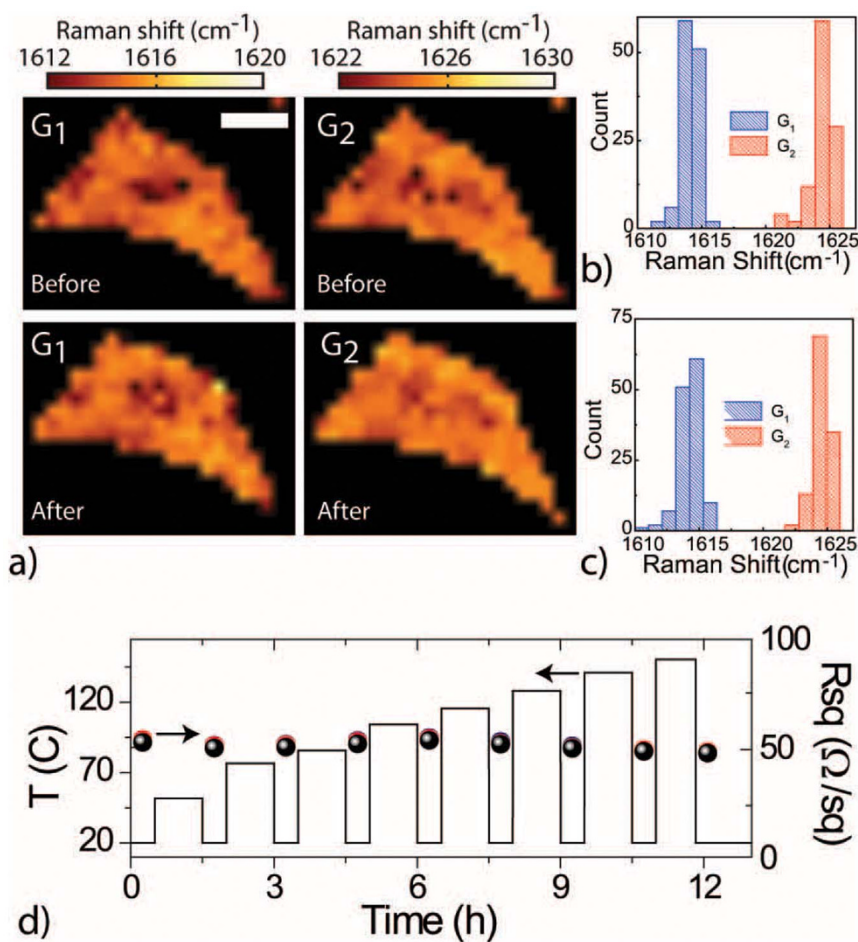
levels of humidity for the sample shown in the micrograph image in the inset. We observe that the initial value of  $R_{sq} \approx 13 \Omega/sq$  is unchanged after exposing the sample to  $H > 95\%$  for 25 days. Furthermore, *in situ* measurements of  $R_{sq}$  while exposing the device to high levels of humidity show that  $R_{sq}$  decreases when the sample is exposed to high levels of humidity (see graph in the inset of Figure 1h). This drop in resistance is a reversible process, since the initial value of  $R_{sq}$  is restored in the sample in dry atmosphere. These observations suggest that water molecules condensed on the surface of FeCl<sub>3</sub>-FLG might contribute to an increase of electrical conductivity without causing irreversible changes to the material. Insight in the microscopic origin behind the stability of this material in atmosphere can be gained when considering that in bulk intercalated graphite the carbon interlayer spacing of non intercalated regions in close proximity to intercalated regions have values similar to the pristine case<sup>22</sup>. Consequently, the large interlayer binding energy characterizing these non-intercalated regions can effectively act as a diffusion barrier for intercalants. In FeCl<sub>3</sub>-FLG a similar mechanism is likely to occur near the edges of the flakes, whereby a narrow de-intercalated edge blocks the diffusion of FeCl<sub>3</sub> molecules out of the structure.

To further evaluate the suitability of FeCl<sub>3</sub>-FLG for future electronic applications, we also need to characterize the stability of this material to high temperatures. Also in this case we conduct a comparative study of Raman spectroscopy and electrical transport

characterization before and after heating FeCl<sub>3</sub>-FLGs in atmosphere on a hot plate. Figure 2a shows the colour coded maps of the Raman shift of G<sub>1</sub> and G<sub>2</sub> before (top graphs) and after (bottom graphs) heating the sample for 1h on a hotplate in atmosphere at 100°C. It is apparent that Raman shifts of just a few cm<sup>-1</sup> are measured, corresponding to the accuracy of the spectrometer. This is more clearly seen when comparing the corresponding histograms of the Raman shifts for G<sub>1</sub> and G<sub>2</sub> peaks before (Figure 2b) and after heating the sample (Figure 2c). The electrical transport measurements also show no significant change of the room temperature square resistance after heating the sample to subsequently higher temperatures from 50°C up to 150°C in multiple steps of 1h duration, see Figure 2d.

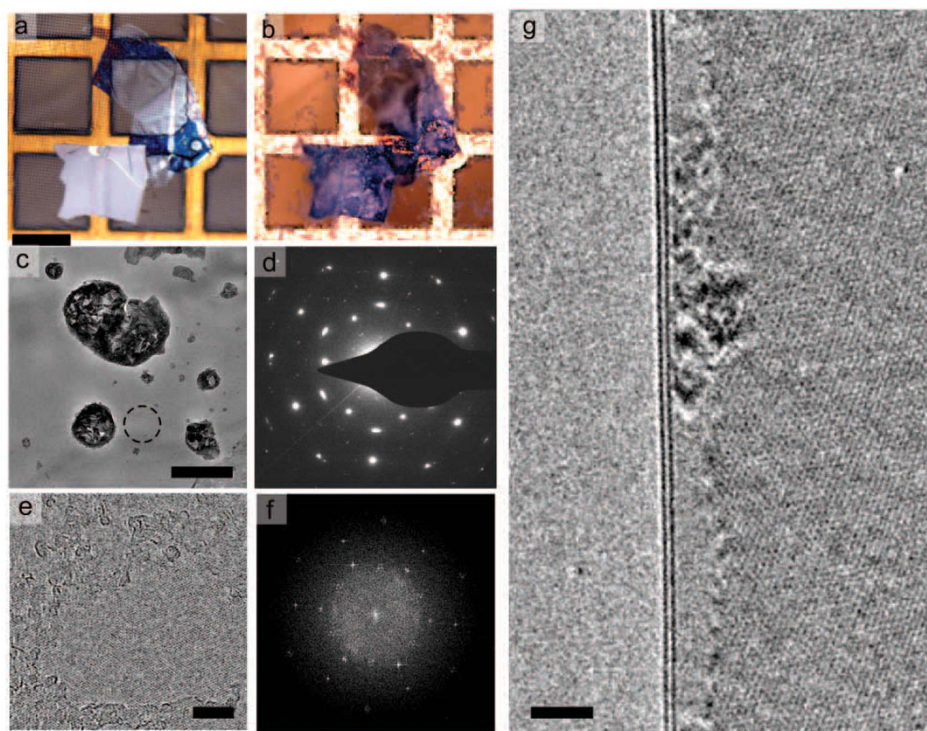
## Discussion

To elucidate the microscopic origin of the thermal stability we have conducted a study of the structure of FeCl<sub>3</sub>-FLG upon heating up to 620°C in vacuum with an FEI Titan E-Cell 80-300ST aberration-corrected transmission electron microscope equipped with an inconel-based heating holder. A few layer graphene flake was transferred to a standard TEM grid (see Fig. 3a) using a published technique<sup>19</sup> and intercalated with FeCl<sub>3</sub> with the two zone method (see Fig. 3b–c). Imaging is performed at 80 keV to reduce the effects of knock-on damage<sup>20</sup>. After intercalation additional lattice periodicities are visible in the multilayer flake and can be observed both in selected area diffraction patterns and high-resolution imaging



**Figure 2** | (a) Shows four colour coded Raman maps of the G<sub>1</sub> and G<sub>2</sub> peaks before (top panels) and after (bottom panels) heating of a representative FeCl<sub>3</sub>-FLG to 100°C for 1 h. The white scale bar corresponds to 5 μm. The graphs in (b) and (c) are the corresponding histograms of the Raman shift of G<sub>1</sub> and G<sub>2</sub> before and after heating the sample. The plots in (d) show the values of  $R_{sq}$  measured at room temperature (top graph) after heating the sample to subsequently higher temperatures for 1 h (bottom graph). The black, red and blue data points refer to three different choices of contact probes on the same flake.





**Figure 3** | (a), (b) Optical images of a few-layer graphene flake before and after intercalation. Scale bar 100  $\mu\text{m}$ . (c) Low magnification image of intercalated multilayer graphene. Nanoparticles of  $\text{FeCl}_x$  can be seen on the surface. Scale bar 250 nm. (d) Selected area diffraction pattern of region indicated in (c). (e) High resolution image of  $\text{FeCl}_3$ -intercalated few layer graphene after heating to 850 K. Scale bar 5 nm. (f) Fourier transform of region in (e). (g) Edge of folded bilayer after heating to 850 K -  $\text{FeCl}_3$  periodicity is visible up to 1–2 nm from the edge of the folded bilayer. Scale bar 2 nm.

(Fig. 3d–f). These periodicities are also visible in a folded bilayer region of the intercalated graphene flake (Fig. 3g), proving that this is a stage I intercalated compound. More specifically, the intercalated material is visible throughout the bilayer region except for some 1–2 nm away from the edge of a folded bilayer, which is not intercalated and that is measured *prior* to any heating process. It is likely that a small radius of curvature in the bilayer graphene results in the exclusion of intercalated material here - the width of this observed exclusion region corresponds well to the typically observed minimum diameters of double-walled carbon nanotubes<sup>21</sup>.

The TEM studies also show on the surface of the suspended graphene flake the presence of nanoparticles, probably consisting of  $\text{FeCl}_x$ , of less than 250 nm diameter, resulting from an excess of intercalant species condensing on the surface of graphene (Fig. 3c). During heating of the sample up to 620°C over 1500 sec at a pressure of  $10^{-5}$ – $10^{-6}$  mbar and concurrent imaging no change is observed in the structure of the sample -  $\text{FeCl}_3$  intercalated bilayer graphene retains additional periodicities due to the graphene and intercalant (Supplementary Video S1). We note that the nanoparticles visible on the surface of intercalated graphene also do not change structure during this ramped heating.  $\text{FeCl}_3$  has a melting point of 315°C, as compared to the melting point of  $\text{FeCl}_2$  of 677°C, indicating that the particles on the surface are likely formed by  $\text{FeCl}_2$ . These studies show that the intercalation of  $\text{FeCl}_3$  is stable at least up to 620°C in vacuum.

In conclusion, we demonstrate that  $\text{FeCl}_3$  intercalated few-layer graphene is highly stable to high levels of humidity and to high temperature. This is supported by a systematic comparative study of the measurements obtained from three complementary techniques: (1) Raman spectroscopy, (2) high resolution transmission electron microscopy, and (3) electrical transport. The Raman spectra and the square resistance of this material are unchanged upon exposing continually  $\text{FeCl}_3$ -FLG to an atmosphere with relative humidity

up to 100% at room temperature for at least 25 days, and to a temperature up to 150°C in atmosphere. At the same time high resolution transmission microscopy confirms that the structure of the material is unaffected by heating  $\text{FeCl}_3$ -FLG up to 620°C in vacuum. The surprising stability of the structure and electrical properties of  $\text{FeCl}_3$ -FLG together with its unique values of low square resistance and high optical transparency, makes this material an attractive replacement for ITO in future transparent and flexible electronic applications.

## Methods

**Measurement techniques.** The Raman spectra were measured in air and at room temperature with a Renishaw spectrometer using a 532 nm laser wavelength with a 1.5  $\mu\text{m}$  spot size and 1 mW of incident power.

The electrical measurements were conducted in a four terminal geometry using an AC current bias (1  $\mu\text{A}$  and 11 Hz frequency) while the voltage was recorded using phase sensitive lock-in detection.

The relative humidity was detected using a HIH-4000 Humidity sensor from Honeywell, and controlled by passing dry nitrogen or ambient air through DI water in a gas bubbler.

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## Author contributions

M.F.C. and S.R. conceived and directed the experiment. D.J.W. conducted the fabrication, Raman and electrical measurements. T.H.B. produced intercalated few-layer graphene and conducted the electrical measurements at high temperature in atmosphere. T.B. conducted the HRTEM measurements. T.B. and P.B. interpreted the HRTEM measurements.

## Additional information

**Supplementary information** accompanies this paper at <http://www.nature.com/scientificreports>

**Competing financial interests:** The authors declare no competing financial interests.

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