

ORIGINAL ARTICLE

Cocrystallization phenomenon of polyoxymethylene blend samples between the deuterated and hydrogenated species

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The phenomenon of cocrystallization between the deuterated (D) and hydrogenated (H) species of a series of polyoxymethylene blends with various H/D ratios has been established on the basis of the detailed analysis of thermal and infrared (IR) spectral data. The melting and crystallization temperatures were found to shift continuously toward a higher temperature with an increase in the D content. The IR spectra were found to change remarkably depending on the H/D blend content. Both the thermal and the IR spectral observations definitely indicated the cocrystallization between the H and D chains in the common crystallite. These continuous changes could be interpreted by assuming that the arrays of the H and D chains in crystal lattice are random.

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INTRODUCTION

Semicrystalline polymers show complex microstructures or higher-order structures that consist of crystalline and amorphous regions. Control of the higher-order structure may govern the physical properties of the polymers. A better understanding of the formation process is required in order to control the higher-order structure of the polymer and develop new materials with better properties. As the first step, the aggregation state of the molecular chains in the crystallized bulk sample must be determined at the molecular level. As the polymer solid is built up by the complicated aggregation of more or less entangled chains, it is difficult to trace the trajectory of a single particular chain in the bulk sample. The isotopic blend between the deuterated (D) and hydrogenated (H) species is considered one of the most useful methods for studying the behavior of a single chain in the aggregation state on the basis of the difference in the thermal and vibrational characters that are due to the difference in atomic mass.^{1–22} In some cases, the difference in the neutron scattering length is also utilized successfully to distinguish the D and H species.^{14,17,18}

In such cases of isotopic blends, the *cocrystallization* must be confirmed to take place between the H and D species in the crystalline lattice. For example, the blends between the H and D species of the various types of polyethylene have been comprehensively studied for clarification of the cocrystallization phenomenon between the H and D species.^{1–4,8–11,14,16–19} In the case of samples that are a blend of H high-density polyethylene and D high-density polyethylene, slow

cooling from the melting point, in which the D (or H) chains gathered together in the H (or D) matrix, showed the phase segregation phenomenon.² However, the cocrystallization phenomenon was observed for the blend samples between D high-density polyethylene and H linear low-density polyethylene with about 17 ethyl branches/1000 carbon atoms even when cooled slowly from the melting point.^{9–11,16–19} The D/H blend samples that show this perfect cocrystallization phenomenon are useful for the tracing of a single chain during the crystallization process and also for the clarification of the chain folding mode in the crystallization phenomenon. In fact, the systematic interpretation of the various kinds of data obtained from the blend samples of D high-density polyethylene/linear low-density polyethylene (17 ethyl branching) lead to a constructive conclusion concerning the aggregation state of the chains and chain folding in the crystalline lamellae.¹⁸ A similar study has been made for the D/H blend samples of *isotactic* polypropylene, where the helical chains of the D and H species coexist randomly in the common crystal lattice.²⁰

In the present paper, polyoxymethylene (POM, $[-CH_2O-]$) is treated as another typical example of the cocrystallization of the D and H species, when cooled from the melting point. POM is one of the most important engineering thermoplastics: it is commonly used to replace metals and alloys because of its high tensile strength, impact resistance, stiffness, good dimensional stability and corrosion resistance. In the crystal structure of POM, the non-uniform (9/5) helical chains are packed in the trigonal unit cell of $a=4.47 \text{ \AA}$ and

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$c=17.39 \text{ \AA}$.²³ It should be noted that, the crystal structure is more complicated, as described by Tashiro *et al.*²⁴ in their analysis of synchrotron X-ray and neutron diffraction data. The morphological structure of the crystalline phase may be classified into two types as extreme cases, an extended-chain crystal (ECC) and a folded chain crystal (FCC) or lamellar crystal, depending on the crystallization condition as well as the thermal and mechanical histories.^{25–28} The physical properties of POM are significantly influenced by such crystal morphologies. The typical FCC structure is achieved by crystallization from a dilute solution, which generates hexagon-shaped lamellar crystals with a thickness of about 10 nm where the molecular chains are folded back on the lamellar surfaces, with the stems aligning perpendicular to the surface.²⁹ The typical ECC structure is achieved through the heterogeneous cationic polymerization of trioxane, resulting in micron-sized needle-like single crystals, in which the extended POM molecules align parallel to the needle axis.^{30,31} The crystal morphology found in the melt-grown bulk samples lies between the two above mentioned extremes.^{32–34} The ECC type is also developed on the drawing of the sample.²⁸

Among the various techniques available for the investigation of the morphology of a polymer, Fourier transform infrared spectroscopy (FTIR) may be one of the most powerful methods for POM because of its uniqueness and high sensitivity to changes in the morphology, as shown by Shimomura *et al.*²⁶ The ECC and FCC structures give the remarkably different FTIR spectra, although they show essentially the same X-ray diffraction pattern.²⁶ The IR bands with transition dipole moments parallel to the chain axis (parallel bands; A_2) show a high frequency shift (as high as 100 cm^{-1}) for the FCC sample compared with the ECC sample, whereas the IR bands with transition dipole moments perpendicular to the chain axis (perpendicular bands; E_1) are unaltered. The drastic shift of the parallel bands depending on the morphology can be interpreted by assuming that there are long-range interactions among the transition dipoles.^{27,28}

The blending of the H and D species of POM may affect the IR spectral profile sensitively. This will result in more complicated IR spectra as both the effect of the morphology and the isotope species must be taken into account. Kobayashi *et al.*²⁵ reported IR spectral changes for a sample that was a blend between normal and D paraformaldehydes (PFA and PFA- d_2 , linear oligomers of POM). The A_2 parallel bands exhibited remarkable changes in both the peak position and the intensity with variation in the PFA/PFA- d_2 composition. However, cocrystallization between the H and D chains has not yet been experimentally established for these blend samples. To clarify the isotope effect on the FTIR spectral changes and to disclose the aggregation state of the molecular chains in the crystallized bulk sample, the cocrystallization between the H and D chains must be confirmed experimentally. The present work therefore aims to clarify the cocrystallization behavior between the H and D species of POM utilizing thermal and vibrational spectroscopic techniques.

EXPERIMENTAL PROCEDURE

Samples

We used the hydrogenated polyoxymethylene (POM-H) Duracon M90 with M_w ca. 69300 g mol^{-1} and M_w/M_n ca. 2.3, which was supplied by Polyplastics, Shizuoka, Japan. Duracon M90 is a copolymer of trioxane with a small amount of ethylene oxide, about 2.2 wt%, which is added to avoid thermal decomposition. The deuterated polyoxymethylene (POM-D, M_w ca. $140000 \text{ g mol}^{-1}$, M_w/M_n ca. 2.3) was synthesized by a solid-state polymerization reaction of D trioxane single crystal under γ -ray irradiation (^{60}Co , 6 kGy at 60°C).^{35,36} 1,1,1,3,3,3-Hexafluoro-2-propanol (bp 58°C) was used as a solvent. The samples that were cooled from the melt had to be treated carefully. As revealed in previous

papers,³² the POM samples were heated above the melting points of both the H and the D species and maintained for 5 min. A longer duration of the molten state caused thermal degradation, and a shorter duration was insufficient to melt the crystallite completely.

POM-H/D blended samples of 100:0, 90:10, 70:30, 50:50, 30:70, 10:90 and 0:100 molar ratios were prepared by dissolving POM-H and POM-D in 1,1,1,3,3,3-hexafluoro-2-propanol at an ambient temperature, followed by precipitation into acetone with vigorous stirring. The precipitated samples were dried in a vacuum at 80°C . The samples have been denoted here as POM-H**, with ** corresponding to the mol% of the H species used (for example, POM-H50 (50 mol% H species and 50 mol% D species)).

Measurements

The thermograms were measured by a TA instrument differential scanning calorimeter (DSC) Q1000 under nitrogen gas flow at a heating rate of 2°C min^{-1} from 25 to 200°C . The IR spectra were measured using an FT7000 Varian rapid-scan-type FTIR spectrometer at a resolution power of 2 cm^{-1} . The X-ray diffraction profile was measured using a Rigaku/TTR-III X-ray diffractometer with a Cu-K α line as an incident X-ray beam in the 2θ range of $10\text{--}60^\circ$ at a scanning rate 5°C min^{-1} . The crystallite size along the chain axis (l) was estimated using Scherrer's equation, $l=K\lambda/(\beta\cos\theta)$, where K is a constant (0.9), λ is the wavelength of the incident X-ray beam (1.542 \AA), β is a full width at a half maximum peak intensity (FWHM) and 2θ is a Bragg angle. The FWHM was corrected for the peak broadening effect caused by the slit system of the diffractometer, in which $\beta^2=\beta_{\text{obs}}^2-\beta_{\text{m}}^2$, where β_{obs} was the observed FWHM and β_{m} was the FWHM observed for the Si powder sample.

Isothermal crystallization

The samples were heated above the melting temperature and kept at that temperature for 5 min before quenching to the predetermined crystallization temperature (T_c) at a cooling rate of $70^\circ\text{C min}^{-1}$. An optimal melting condition had been reported by Hama *et al.*³² to erase any remaining effect of the residual crystalline lamellae and to avoid sample decomposition. We followed the procedures described by them. The temperature was kept at T_c until crystallization was completed. The crystallized samples were then quenched to the ambient temperature at a rate of $70^\circ\text{C min}^{-1}$ for evaluation of the crystallite thickness (l). The evaluation of the melting temperature $T_m(l)$ of these samples was made by reheating at 2°C min^{-1} .

RESULTS AND DISCUSSION

Cocrystallization phenomena as viewed by thermal analysis

Figure 1 shows the DSC thermograms that were measured for the blend samples POM-H100, POM-H50 and POM-H0. The single melting (T_m) and crystallization (T_c) peaks were observed to occur in the blend sample at a temperature that was between the melting temperatures of the original pure H and D samples. The inset pattern shown in Figure 1b indicates a second T_c , which might be related to the formation of the ECC structure of POM or the straight tie chains passing through the neighboring lamellae.³² It should be noted that the endothermic peak from the melting of ECC was particularly difficult to detect in the heating process of the H/D blend sample. The DSC peak is generally broader during the heating process than it is during the cooling process; this difference can be typically seen in the main melting and crystallization peaks, as shown in this figure. A similar situation may apply to the ECC melting peak as well. The H/D content dependence of the DSC thermograms that were measured during the heating and cooling processes is shown in Figure 2. The T_m and T_c were found to shift continuously toward a higher temperature with an increase in the D content, implying that cocrystallization was occurring between the H and D chains. It should be pointed out that cocrystallization can occur even when the sample is cooled slowly from the melting point (2°C min^{-1}). Figure 3 shows the T_m and T_c plotted against the H/D content, clearly demonstrating the above mentioned situation. Figure 4 shows a continuous shift of the second

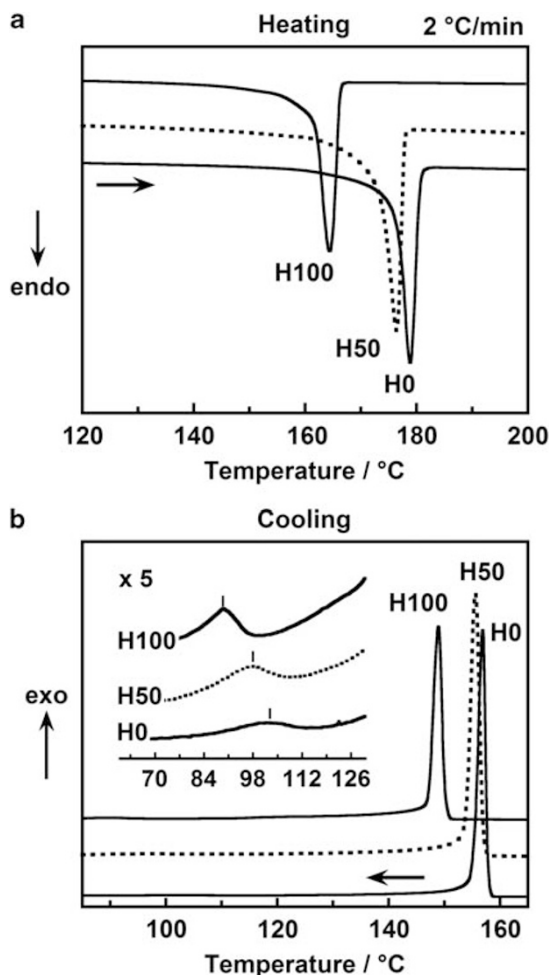


Figure 1 Differential scanning calorimeter thermograms of the polyoxymethylene (POM)-H100, -H50 and -H0 samples measured during (a) the heating and (b) the cooling processes, with an inset indicating the second crystallization temperature T_{c2} . H, hydrogenated.

T_c as a function of the H content. The second T_c observed is found to be 50–60 °C lower than the first T_c for all samples. Figure 5 shows the enthalpy change at T_m plotted against the H molar fraction, which was almost constant independent of the H content.

The DSC data could then be analyzed from a thermodynamic point of view.¹⁸ The probabilities for finding a pair of the H and H chains (H–H), D and D (D–D) chains or H and D (H–D) chains can be expressed approximately as X^2 , $(1-X)^2$ or $2X(1-X)$, respectively, where the X is the molar fraction of the H species. If we assume the coexistence of the H and D chain stems randomly in the common crystallite, then the changes in the enthalpy (ΔH) and entropy (ΔS) at the melting point are expressed by the following equations, respectively:

$$\Delta H = X^2\Delta H_{hh} + (1-X)^2\Delta H_{dd} + 2X(1-X)\Delta H_{hd} \quad (1)$$

$$\Delta S = X^2\Delta S_{hh} + (1-X)^2\Delta S_{dd} + 2X(1-X)\Delta S_{hd} \quad (2)$$

where ΔH_{hh} , ΔH_{dd} and ΔH_{hd} represent the enthalpy changes related to the systems that are constructed by the H–H, D–D and H–D pairs, respectively. At the equilibrium melting point T_m^0 , ΔH and ΔS are

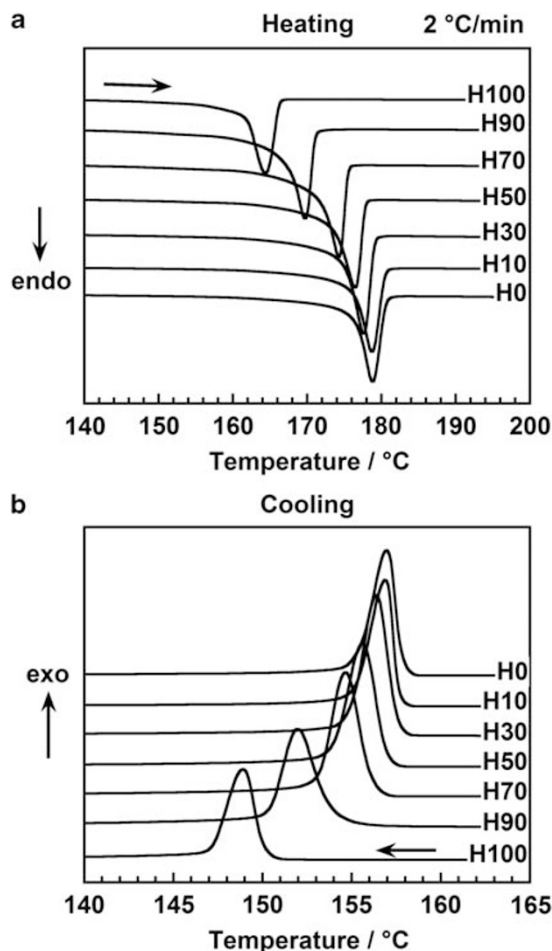


Figure 2 Differential scanning calorimeter thermograms of a series of polyoxymethylene blend samples with various hydrogenated/deuterated contents during the (a) heating and (b) cooling processes.

related by the equation $\Delta H - T_m^0\Delta S = 0$. Then the following equation is obtained:

$$T_m^0 = \frac{X^2\Delta H_{hh} + (1-X)^2\Delta H_{dd} + 2X(1-X)\Delta H_{hd}}{X^2\Delta S_{hh} + (1-X)^2\Delta S_{dd} + 2X(1-X)\Delta S_{hd}} \quad (3)$$

The interatomic interactions between the H and D species are essentially the same as these interactions originate from the Coulombic interactions between electrons and protons. The assumption of $\Delta H_{hh} = \Delta H_{hd} = \Delta H_{dd}$ in Equation (3) is reasonable. If the entropy terms ΔS_{hh} , ΔS_{hd} and ΔS_{dd} are also assumed to be the same, then the T_m^0 given in Equation (3) no longer depends on the H/D. However, the entropy terms cannot be assumed to be identical, as is clearly shown by the difference in the observed melting points between the blend samples. In general, the entropy term ΔS is influenced by many contributions, such as the conformational difference between the crystalline and melted states (conformational entropy; ΔS_{conf}), the vibrational entropy; ΔS_{vib} and the statistically irregular packing of the different species (packing entropy; ΔS_{pack}). Among these sources, the conformational entropy ΔS_{conf} is reasonably assumed to be nearly identical between the H and D chains as the conformational change is mainly due to a change in the skeletal torsional angles, which are essentially identical between these two species. The CD_2 and CH_2 units

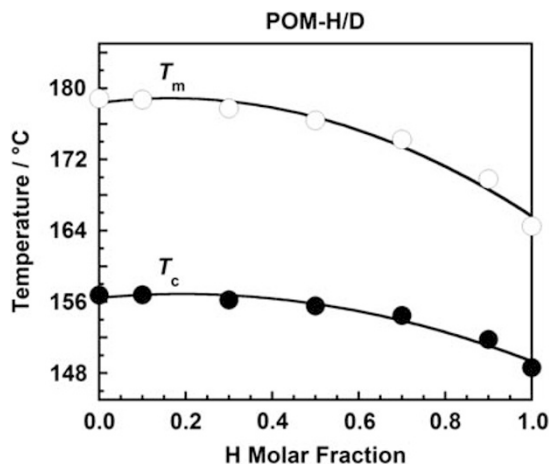


Figure 3 The melting temperature (T_m) and crystallization temperature (T_c) as a function of the hydrogenated (H) molar fraction of the polyoxymethylene-H/deuterated blend samples, observed at heating and cooling rates of $2\text{ }^\circ\text{C min}^{-1}$.

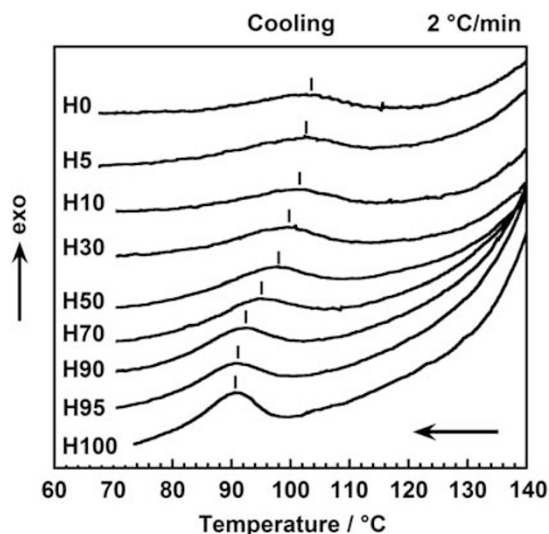


Figure 4 Differential scanning calorimeter thermograms of a series of polyoxymethylene-hydrogenated (H)/deuterated (D) blend samples with various H/D contents during the cooling process, with the second crystallization temperature (T_{c2}) shown.

have different masses, and the vibrational frequencies differ as well. Therefore, the vibrational entropy ΔS_{vib} is appreciably different between these two species. The main reason for the different entropy terms in Equations (2) and (3) ($\Delta S_{\text{hh}} \neq \Delta S_{\text{hd}} \neq \Delta S_{\text{dd}}$) is the difference in the vibrational entropy term. As a result, the entropy change due to the random arrangement of the H and D chain stems in the crystal lattice (ΔS_{pack}), which is expressed in Equation (2), differs depending on the H/D content. Another noteworthy point in the discussion of Equation (3) pertains to the degree of crystallinity of the samples. The crystallinity was estimated using the X-ray diffraction data and was found to be nearly identical for all of the samples used here. Therefore, the effect of crystallinity can conveniently be canceled out in the treatment of Equation (3). Now, utilizing the relation of $\Delta S_{\text{hd}} = \Delta H_{\text{hd}}/T_{\text{mhd}}^{\circ}$ and so on, Equation (3) can be rewritten as below, in which

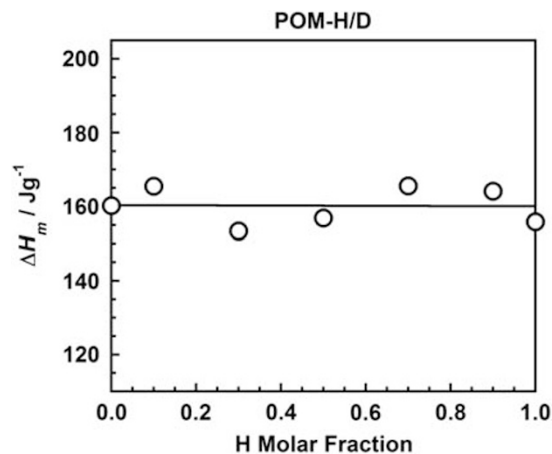


Figure 5 Observed enthalpy of melt for a series of polyoxymethylene-hydrogenated (H)/deuterated (D) blend samples with various H/D contents.

T_{mhh}° , T_{mdd}° and T_{mhd}° are the equilibrium melting temperatures of the pure H, D and H/D blend samples, respectively.

$$T_m^{\circ} = \frac{1}{X^2/T_{\text{mhh}}^{\circ} + (1-X)^2/T_{\text{mdd}}^{\circ} + 2X(1-X)/T_{\text{mhd}}^{\circ}} \quad (4)$$

or

$$1/T_m^{\circ} = X^2/T_{\text{mhh}}^{\circ} + (1-X)^2/T_{\text{mdd}}^{\circ} + 2X(1-X)/T_{\text{mhd}}^{\circ} \quad (5)$$

Equation (5) shows the relationship between the equilibrium melting point T_m° and the molar content X . As can be noticed in the equation, the observed T_m cannot be used, but rather the so-called equilibrium melting temperature T_m° must be applied for all of the samples of the various H/D blend contents.

The equilibrium melting temperature T_m° may be defined as the melting temperature of an ideal crystal of an infinitely large size. The Gibbs–Thomson equation, derived from thermodynamical relations, is given below.³⁷

$$T_m(l) = T_m^{\circ} \left[1 - \frac{2}{\Delta H_f} \left(\frac{\sigma_e}{l} + \frac{\sigma}{b} + \frac{\sigma}{a} \right) \right] \quad (6)$$

where σ and σ_e are the free energies of the lateral ($a \times l$ and $b \times l$) and end surfaces ($a \times b$) of the lamella, with the dimensions $a \times b \times l$, and ΔH_f is the heat of fusion per unit volume of the repeating unit (53 cal cm^{-3}).³⁸ In general, a and b are much larger than l , especially for the FCC structure. Thus, Equation (6) is approximated by

$$T_m(l) \cong T_m^{\circ} - \frac{C}{l} \quad (7)$$

where $C = 2\sigma_e T_m^{\circ} / \Delta H_f$. The plot of $T_m(l)$ vs l^{-1} yields the T_m° from the intercept of the expected straight line. Of course, the σ_e may also be estimated from the slope of this straight line, but this value might be affected by the value ΔH_f . As our purpose is only to estimate, we will skip the discussion of ambiguity in σ_e .

Figure 6 shows the Gibbs–Thomson plot constructed for POM-H100 and POM-H0 as examples. The crystalline thickness l was evaluated by estimating the FWHM (β) for the 009 X-ray reflection, as described in the experimental section. The evaluated values of $\sigma_e(s)$ were $5.04 \pm 0.12 \times 10^{-10}\text{ J m}^{-2}$ for POM-H100 and $8.97 \pm 0.09 \times 10^{-10}\text{ J m}^{-2}$ for POM-H0. Similarly, the $\sigma_e(s)$ were estimated as $5.03 \pm 0.22 \times 10^{-10}$, $5.59 \pm 0.1 \times 10^{-10}$, $6.47 \pm 0.07 \times 10^{-10}$, $8.06 \pm 0.05 \times 10^{-10}$ and $8.98 \pm 0.22 \times 10^{-10}\text{ J m}^{-2}$ for H90, H70, H50, H30 and H10,

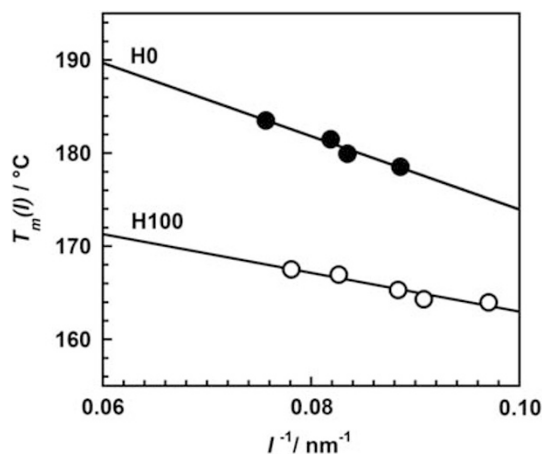


Figure 6 Gibbs–Thomson plots to evaluate the equilibrium melting temperature, where only two cases, polyoxymethylene (POM)-H0 and POM-H100, are shown as examples.

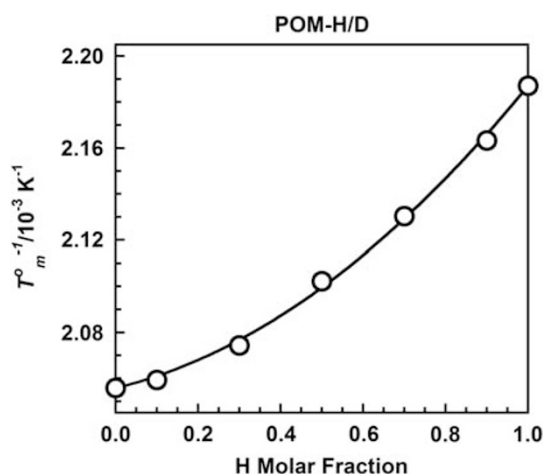


Figure 7 Dependence of an equilibrium melting temperature T_m^0 on the hydrogenated molar fraction. The solid line is the curve that was calculated using Equation (5) in the text, where an adjustable parameter $T_{mhd}^0=208^\circ\text{C}$ was assumed.

respectively. As the value of σ_e is related to the work of forming a chain fold, the lowering of energy caused by the blending of the H and D species makes it possible for the H and D chains to be crystallized in the common crystalline lattice.

The T_m^0 of pure POM-H100 was estimated to be $184.1 \pm 1.0^\circ\text{C}$ and that of pure POM-D (POMH0) was estimated at $213.3 \pm 0.5^\circ\text{C}$. Correspondingly, the T_m^0 (s) were estimated as 189.1 ± 2.1 , 196.3 ± 2.0 , 202.5 ± 0.5 , 208.9 ± 0.3 and $212.4 \pm 1.2^\circ\text{C}$ for H90, H70, H50, H30 and H10, respectively. The T_m^0 of POM-H100 (Delrin 550, homopolymer) was reported to be 210°C by Carter *et al.*,³⁹ 206°C by Majer⁴⁰ and 199.7°C by Wissbrun.⁴¹ The T_m^0 estimated in the present study is lower than that in these reports because a copolymer with ethylene oxide content of 2.2 wt% was used as the pure POM-H sample (Duracon M90). The T_m^0 (s) obtained for all of the samples were used to check the reasonableness of Equation (5). The solid line shown in Figure 7 is the fitted curve, created by assuming that the adjustable parameter $T_{mhd}^0=208^\circ\text{C}$. The T_m^0 changes continuously, indicating the perfect cocrystallization of the D and H types of POM species.

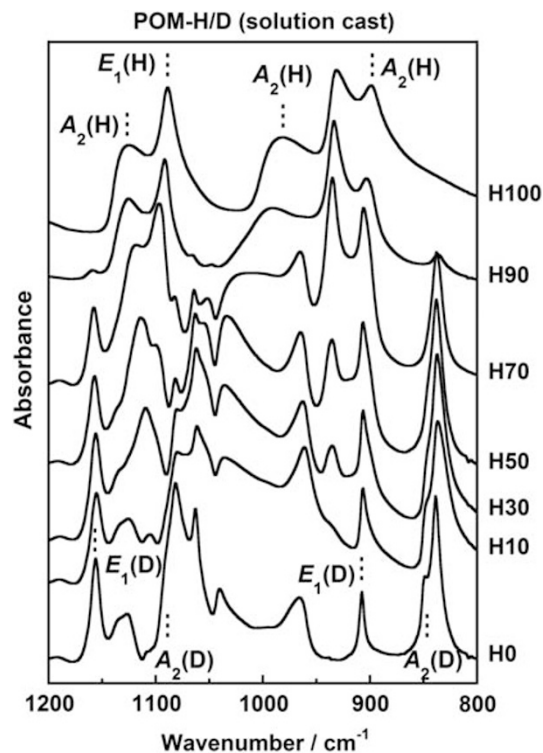


Figure 8 Fourier transform infrared spectra of solution-cast polyoxymethylene-hydrogenated (H)/deuterated (D) blend samples with various H/D compositions, measured at room temperature.

Cocrystallization phenomena as viewed by vibrational spectroscopy

To confirm the cocrystallization phenomena as revealed by the DSC data, the IR spectra were measured for a series of the blend samples. It should again be pointed out that the IR spectra of POM are strongly dependent on the morphology of the sample. At the same time, the coexistence of the D and H chain stems in the crystalline lamellae may also affect sensitively the spectral profiles. As the purpose of the present study is to confirm the coexistence of the H and D chains in the crystalline region, it is important to use a common morphology adjustment for all of the samples. The single crystals grown from the dilute solutions or the FCC were prepared as having the typical morphology. For the IR spectral measurement, the dilute solution of POM in 1,1,1,3,3,3-hexafluoro-2-propanol solution (~ 1 wt%) was cast onto a KBr single-crystal plate. The evaporation rate of the solvent was reduced by reducing the crystallization temperature. Figure 8 shows the series of IR spectra that were observed for the samples with different H/D compositions. Judging from the spectral profile of POM-H0 and POM-H100, the samples commonly have the FCC morphology. As another morphological characteristic, the IR spectra were measured for the melt-cooled samples as shown in Figure 9.

As mentioned previously, the drastic shift in the parallel bands depending on the morphology can be interpreted by assuming that there are long-range interactions among the transition dipoles.^{27,28} Vibrations of the CO bonds induce the transition dipoles, and the interactions between these dipoles affect the vibrational frequency in a feedback manner. The vibrational frequency shift is expressed by the following equation:

$$\Delta\nu = (1/\nu)(\partial\mu/\partial Q_k)^2 \sum_j (1/R_j^3) \quad (8)$$

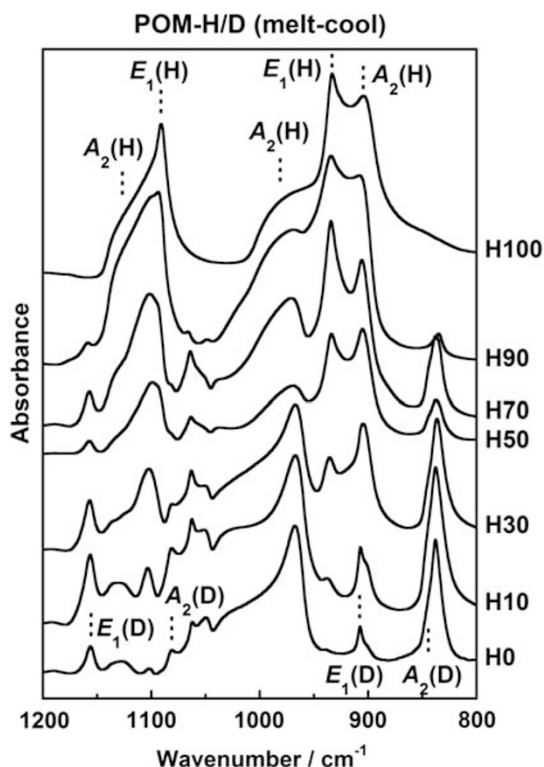


Figure 9 Fourier transform infrared spectra of melt-cooled polyoxymethylene-hydrogenated (H)/deuterated (D) blend samples with various H/D compositions, measured at room temperature.

where ν is the vibrational frequency, $(\partial\mu/\partial Q_k)$ is the transition dipole moment and R_j is the lateral distance of the j th chain from the central chain.²⁸ The total sum of dipole–dipole interactions over the crystalline region is determined by the geometry or the morphology of the crystal phase. Vibrational coupling between the transition dipoles can, in principle, occur for all of the substances, but the magnitude of the frequency shift is dependent on the strength of the dipole–dipole interactions.²⁸ The C–O bonds of POM give very large dipoles, which result in the remarkably large frequency shifts observed.

Now, let us analyze the blend sample of POM-D and -H species. We assume that the normal POM chains or the hydrogenous POM chains are homogeneously blended with the fully D POM chains in the common crystal lattice. The H and D chain species show different vibrational frequencies because of differences in the atomic mass. By changing the H/D blend ratio, the IR spectral profile will be changed, as illustrated in Figure 10a. At the same time, however, we must consider the effect of vibrational couplings between the transition dipoles. As the CD₂O units vibrate at much lower frequencies than the CH₂O units, the vibrational couplings between the CH₂O units (CD₂O units) are blocked when the CD₂O (CH₂O) groups penetrate into the aggregated space of the CH₂O (CD₂O) groups, where the couplings between the CH₂O and CD₂O units are almost negligible. As a result, the total sum of the dipole–dipole interactions expressed in Equation (8) is modified, and the parallel bands of the CH₂O (CD₂O) units or the A₂ bands will show remarkable changes in the vibrational frequencies compared with the original values, as shown in Figure 10b.²⁸ The E₁ modes with transition dipoles that are perpendicular to the chain axis are doubly degenerated and are not affected by the dipole–dipole interactions.²⁸ In this manner, the blending of the H- and D-POM chains in the crystal lattice affects the IR spectra of POM

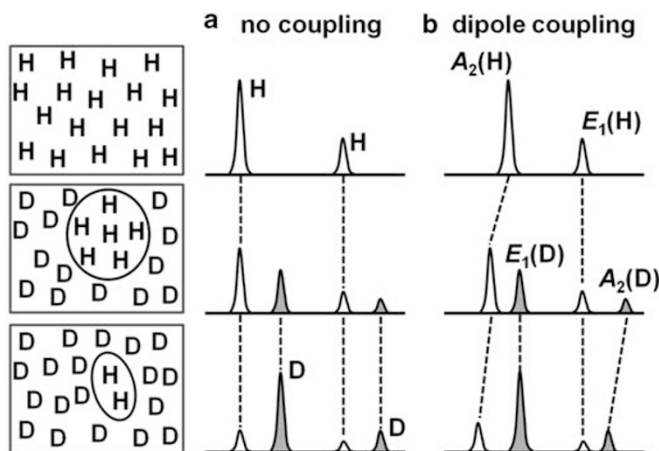


Figure 10 Schematic illustration of the relationship between the molecular aggregation state and the infrared spectral profile. (a) No dipole–dipole interactions between the same types of chain stems exist, and the relative band intensities change almost in proportion to the hydrogenated (H)/deuterated (D) content. (b) The position and relative intensity of the A₂ bands, with transition dipoles parallel to the chain axis, are shifted depending on the spatial distribution of the H and D chain stems in the crystal lattice. The doubly degenerated E₁ mode bands do not show any frequency shift as the transition dipoles are perpendicular to the chain axis.²⁸ Some of the bands detected in Figures 8 and 9 are assigned to the A₂ and E₁ bands of the H and D species.

more markedly than is expected from the simple change in the relative intensity of the vibrational bands, as illustrated in Figure 10a. As indicated in Figures 8 and 9, the complicated spectral changes that are observed for the blend samples may also be interpreted in this manner as long as the D and H chain stems are mixed together randomly in the common lattice. As the domain sizes of the H (and D) species surrounded by the D (and H) species in the crystalline region may distribute statistically randomly, the summation of the dipole–dipole interactions changes in a complicated manner, making it difficult to reasonably simulate the observed spectral changes. The interpretation of these remarkable changes in the spectral profiles is now being processed on the basis of the equation of dipole–dipole coupling.

The cocrystallization can also be confirmed by observing the simultaneous appearance of the crystalline bands of the H and D chains in the crystallization process from the melt. For example, Figure 11 shows the temperature dependence of the IP spectra that were measured for the POM-H50 blend sample during the cooling process from the melting point at a cooling rate of 2 °C min⁻¹. As plotted in Figure 11b, the H band (1237 cm⁻¹) and the D band (1158 cm⁻¹) were found to begin simultaneously appearing at around 159 °C. This clearly indicates that the D and H chains approach the frontier plane of the growing lamella almost simultaneously and are trapped together as crystalline stems, as illustrated in Figure 12.

CONCLUSION

The continuous changes observed in the DSC thermogram and IR spectra for a series of POM-H/D blend samples have been reasonably interpreted by assuming the random distribution of the H and D chain stems in the common crystalline lattice or with a concept of cocrystallization between the H and D chains. The continuous shift of the equilibrium melting temperature was reasonably analyzed by the random distribution of the H and D chain stems. The simultaneous appearance of the IR crystalline bands of the H and D species also

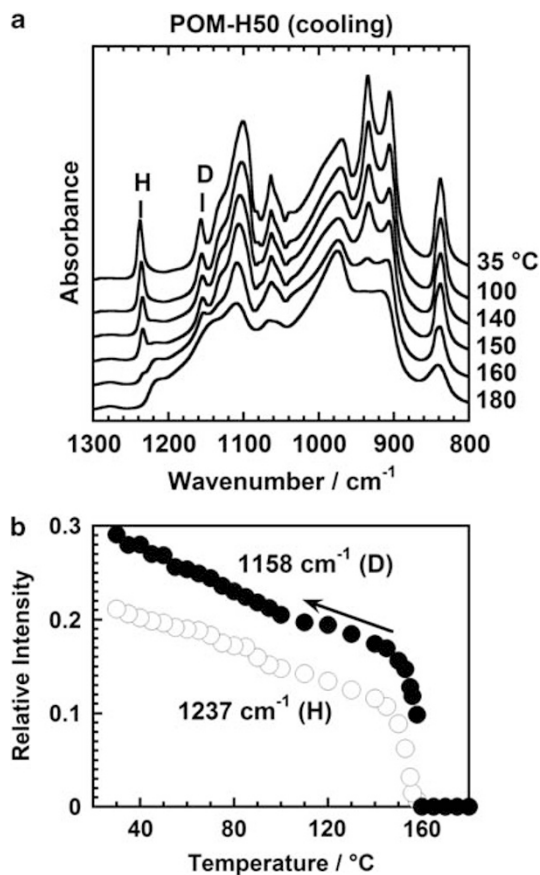


Figure 11 (a) Temperature dependent Fourier transform infrared spectra and (b) the integrated intensity of the hydrogenated (H) bands (1238 cm^{-1}) and deuterated bands (1158 cm^{-1}) observed for the polyoxymethylene-H50 blend sample during the cooling process from the melting point.

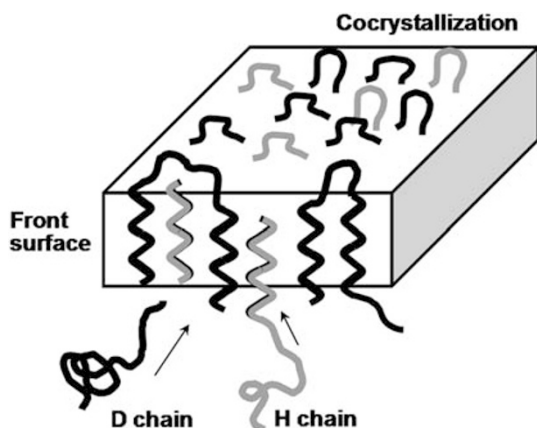


Figure 12 Schematic illustration of the cocrystallization process of the polyoxymethylene (POM)-deuterated (D) and POM-hydrogenated (H) chains, where the D and H chains simultaneously approach the front surface of the lamella.

supports this concept. It should be noticed, however, that the thermal and FTIR data reported in the present paper are only a portion of the experimental evidence that supports the concept of cocrystallization of the POM-D and -H species. The wide-angle X-ray diffraction profiles,

which are not given in the present paper, have been found to be essentially the same among all of the D/H samples used here, telling us that the crystal structures are the same for all of the samples independent of the D/H content. From all of these experimental data, it is not unreasonable to conclude that cocrystallization occurs between the D and H species of POM when they are cooled from the melting point. As for the IR spectral data, it may in principle be possible to deduce the spatial distribution of the H and D chains in the crystallization process on the basis of the evolution of the IR spectral profile. The details of this analysis will be reported elsewhere.

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